

Selective Extraction of Copper, Mercury, Silver and Palladium Ions from Water Using Hydrophobic Ionic Liquids

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Nicolas Papaiconomou^{1, 2}, Jong-Min Lee^{1, 2}, Justin Salminen^{1, 3}, Moritz Von Stosch^{1, 4} and John M. Prausnitz^{1, 2}*

1 Department of Chemical Engineering, University of California, Berkeley, CA 94720-1462

2 Chemical Sciences Division, Lawrence Berkeley National Laboratory, Berkeley, CA 94720

3 Environmental Energy Technology Division, Lawrence Berkeley National Laboratory, Berkeley, CA 94720

4 Institut fuer Verfahrenstechnik, Technische Universitaet, Aachen, Germany

* Corresponding author. Email: prausnit@cchem.berkeley.edu

Abstract

Extraction of dilute metal ions from water was performed near room temperature with a variety of ionic liquids. Distribution coefficients are reported for fourteen metal ions extracted with ionic liquids containing cations 1-octyl-4-methylpyridinium [4MOPYR]⁺, 1-methyl-1-octylpyrrolidinium [MOPYRRO]⁺ or 1-methyl-1-octylpiperidinium [MOPIP]⁺, and anions tetrafluoroborate [BF₄]⁻, trifluoromethyl sulfonate [TfO]⁻ or nonafluorobutyl sulfonate [NfO]⁻. Ionic liquids containing octylpyridinium cations are very good for extracting mercury

ions. However, other metal ions were not significantly extracted by any of these ionic liquids. Extractions were also performed with four new task-specific ionic liquids. Such liquids containing a disulfide functional group are efficient and selective for mercury and copper, whereas those containing a nitrile functional group are efficient and selective for silver and palladium.

Introduction

Wastewaters often contain offensive cations.¹⁻⁸ Because of their high affinity for water, it is difficult to remove them using conventional solvents for liquid – liquid extraction. Hydrophobic ionic liquids may provide a useful extraction process.⁹⁻²⁸

Because the properties of ionic liquids are tunable, it may be possible to identify some ionic liquids that have low viscosity, very low solubility in water, and high affinity for selected metal ions.²⁹⁻⁴¹

Previous reports show that ionic liquids containing imidazolium cations with alkyl chains and $[\text{PF}_6]^-$ anion are not useful for extraction of cations from water.⁹⁻¹⁵ Previous studies using ionic liquids containing $[\text{PF}_6]^-$ anion indicate that this anion decomposes to produce HF when contacted with water.⁴² A few recent reports have investigated the toxicity of ILs in water, suggesting that ionic liquids may be toxic for the environment.⁴³⁻⁴⁶ When co-extractants, such as crown-ethers, calixarene or other organic ligands, are added to ionic liquids, it is possible to attain high distribution coefficients for cations such as mercury, sodium or strontium.¹⁵⁻²¹ However, because ion-exchange may occur between the metal ion and the ionic-liquid's cation, the ionic liquid may decompose in water.²²⁻²⁵

Task-specific ionic liquids (TSIL), can extract cations without ion-exchange.²⁶⁻²⁸ TSILs are ionic liquids with functional groups appended to the cation. Some TSILs appear to be efficient

for extraction with high selectivity for some metal ions such as mercury, cadmium,^{27,28} or actinides.²⁶

In this work, we investigate the ability of selected ionic liquids to extract offensive or valuable cations from water. We have studied the extraction properties of ten ionic liquids: 1-octyl-3-methylimidazolium tetrafluoroborate [OMIM]⁺[BF₄]⁻, 1-octyl-3-methylpyridinium tetrafluoroborate [3MOPYR]⁺[BF₄]⁻, 1-octyl-4-methylpyridinium tetrafluoroborate [4MOPYR]⁺[BF₄]⁻, 1-octyl-4-methylpyridinium trifluoromethylsulfonate [4MOPYR]⁺[TfO]⁻, 1-octyl-4-methylpyridinium nonafluorobutylsulfonate [4MOPYR]⁺[NfO]⁻ and 1-octyl-4-methylpyridinium bis(trifluoromethyl)sulfonylimide [4MOPYR]⁺[Tf₂N]⁻, 1-butyronitrile-4-methylpyridinium bis(trifluoromethyl)sulfonylimide [4MPYRCN]⁺[Tf₂N]⁻, 1-methyl-1-butyronitrilepiperidinium bis(trifluoromethyl)sulfonylimide [MIPICN]⁺[Tf₂N]⁻, 1-methyl-1-[4,5-bis(methylsulfide)]pentylpiperidinium bis(trifluoromethyl)sulfonylimide [4MPIPS₂]⁺[Tf₂N]⁻, and 1-methyl-1-[4,5-bis(methylsulfide)]pentyrrolidinium bis(trifluoromethyl)sulfonylimide [4MPYRROS₂]⁺[Tf₂N]⁻. We have synthesized new TSILs: [4MPYRCN]⁺[Tf₂N]⁻, [MIPICN]⁺[Tf₂N]⁻, [4MPIPS₂]⁺[Tf₂N]⁻ and [4MPYRROS₂]⁺[Tf₂N]⁻. Figure 1 shows structures and abbreviations of ionic liquids used here.

Except for [4MOPYR]⁺[NfO]⁻, all ionic liquids studied here are liquid at room temperature. Some physical properties including solubility in water are given elsewhere.⁴⁰ [4MOPYR]⁺[NfO]⁻ is solid at room temperature with melting point 78 °C. However, when mixed with water, [4MOPYR]⁺[NfO]⁻ near 25 °C, we observe two liquid phases. The water content in [4MOPYR]⁺[NfO]⁻ decreases its melting point below room temperature.

Experimental:

Synthesis

1-octyl-3-methylpyridinium tetrafluoroborate and 1-methyl-3-octylimidazolium tetrafluoroborate were used as received from Solvent Innovation. Synthesis of [3MOPYR]⁺[BF₄]⁻, [4MOPYR]⁺[BF₄]⁻, [4MOPYR]⁺[Tf₂N]⁻, [4MOPYR]⁺[TfO]⁻, and [4MOPYR]⁺[TfO]⁻ was reported previously.⁴⁰ Syntheses of our new TSILs are reported below.

The purity of each synthesized ionic liquids was checked with NMR and elemental analysis.



8.8 g ($8.53 \cdot 10^{-2}$ mol) of 4-chlorobutyronitrile, used as received from Aldrich (99 % purity), was added drop-wise under nitrogen atmosphere, to a solution of 5 g ($5.37 \cdot 10^{-1}$ mol) of 4-methylpyridine, used as received from Sigma (98 % purity), and 5 cm³ acetonitrile. After cooling to room temperature, the solvent was removed by rotary evaporation and the solid product was washed with ethyl acetate (3×100 cm³) and filtered. Ethyl acetate was removed by rotary evaporation, and the resulting solid was set under vacuum for 12 h. 8.62 g ($4.39 \cdot 10^{-2}$ mol) of 1-butyronitrile-4-methylpyridinium chloride, [4MPYRCN]⁺[Cl]⁻, was obtained as a gray solid (yield: 82 %). 4 g ($2.04 \cdot 10^{-2}$ mol) of [4MPYRCN]⁺[Cl]⁻ and 5.95 g ($1.86 \cdot 10^{-2}$ mol) of potassium bis(trifluoromethylsulfonyl)imide, used as received from Wako Chemicals (99 % purity), were mixed in a flask containing 100 cm³ water for 24 h. After filtering, removal of acetonitrile by rotary evaporation, washing with methylene chloride (20 cm³) and water (3×20 cm³), the product was set under high vacuum for 24 h. The final product was 8.1 g of light-red liquid (yield: 90 %).

$^1\text{H NMR}$ (400 MHz, CDCl_3): δ 2.163-2.367 (2H, m, CH_2), 2.407-2.542 (2H, m, CH_2), 2.571-2.650 (3H, d, CH_3), 4.438-4.693 (2H, m, N-CH_2), 7.657-7.890 (2H, t, $\text{CH}_{\text{pyr-m}}$), 8.486-8.919 (2H, t, $\text{CH}_{\text{pyr-o}}$).

(Found: C, 33.52; H, 3.13; N, 9.47. Calc. for $\text{C}_{12}\text{H}_{13}\text{N}_3\text{O}_4\text{S}_2\text{F}_6$: C, 32.65; H, 2.97; N, 9.52).

$[\text{MIPCN}]^+[\text{TF}_2\text{N}]^-$

6.5 g ($6.3 \cdot 10^{-2}$ mol) of 4-chlorobutyronitrile, used as received from Aldrich (99 % purity), was added drop wise, under nitrogen atmosphere, to a solution of 5 g ($5.04 \cdot 10^{-2}$ mol) of 1-methylpiperidine, used as received from Sigma (98 % purity), and 5 cm^3 acetonitrile. After cooling to room temperature, the solvent was removed by rotary evaporation and the solid product was washed with ethyl acetate ($3 \times 100 \text{ cm}^3$) and filtered. Ethyl acetate was removed by rotary evaporation, and the resulting solid was set under vacuum for 12 h. 8.62 g ($4.25 \cdot 10^{-2}$ mol) of 1-methyl-1-butyronitrilepiperidinium chloride, $[\text{MIPCN}]^+[\text{Cl}]^-$, was obtained as a gray solid (yield: 84 %). 4 g ($1.97 \cdot 10^{-2}$ mol) of $[\text{MIPCN}]^+[\text{Cl}]^-$ and 7.87 g ($2.47 \cdot 10^{-2}$ mol) of potassium bis(trifluoromethylsulfonyl)imide, used as received from Wako Chemicals (99 % purity), were mixed in a flask containing 100 cm^3 water for 24 h. After filtering, removal of acetonitrile by rotary evaporation, washing with methylene chloride (20 cm^3) and water ($3 \times 20 \text{ cm}^3$), the product was set under high vacuum for 24 h. The final product was 7.09 g of light-red liquid (yield: 80.4 %).

$^1\text{H NMR}$ (400 MHz, acetone-d_6): δ 1.736-1.816 (2H, m, CH_2), 2.000-2.011 (4H, m, CH_2), 2.284-2.363 (2H, m, CH_2), 2.655-2.691 (2H, t, CH_2), 3.294-3.322 (3H, d, N-CH_3), 3.558-3.605 (2H, m, N-CH_2), 3.615-6.805 (4H, m, N-CH_2).

(Found: C, 32.07; H, 4.30; N, 9.27. Calc. for $\text{C}_{12}\text{H}_{19}\text{N}_3\text{O}_4\text{S}_2\text{F}_6$: C, 32.21; H, 4.28; N, 9.39).



4.29 g ($2.9 \cdot 10^{-2}$ mol) of 5-bromopentene, used as received from Aldrich (99 % purity), was added drop wise, under nitrogen atmosphere, to a solution of 2 g ($2.35 \cdot 10^{-2}$ mol) of 1-methylpyrrolidine, used as received from Sigma (98 % purity), and 10 cm³ acetonitrile. After the same procedure described above, 5.08 g (yield: 68.5 %) of $[\text{MPPYRRO}]^+[\text{Tf}_2\text{N}]^-$ were obtained with a light-yellow color. K-10 montmorillonite, used as received from Fluka Chemical Company, was modified by adding ZnCl₂ (2 mmol per gram of clay) to water-impregnated clay.⁴⁷ The catalyst was activated by heating at 150 °C for 24 h in a vacuum oven. Then, 5 g of the catalyst was mixed with 30 ml of 1,2-dichloroethane and 15 mmol of dimethyl disulfide. A solution of 10.4 mmol of $[\text{MPPYRRO}]^+[\text{Tf}_2\text{N}]^-$ in 5 cm³ of 1,2-dichloroethane was added drop-wise to the mixture. After filtering, removal of the solvents by rotary evaporation, the product was set under high vacuum for 24 h. The final product was 3.79 g of a light-yellow liquid (yield: 69 %).

¹H NMR (400 MHz, acetone-d₆): δ 2.018-2.045 (2H, m, CH₂), 2.064-2.085 (3H, s, S-CH₃), 2.092-2.112 (3H, s, S-CH₃), 2.113-2.200 (2H, m, CH₂), 2.257-2.356 (4H, s, CH₂), 2.634-2.706 (H, m, S-CH), 2.769-2.908 (2H, m, S-CH₂), 3.222-3.513 (3H, m, N-CH₃), 3.527-3.620 (2H, m, N-CH₂), 3.836-3.857 (4H, s, CH₂).

(Found: C, 32.92; H, 4.82; N, 5.89. Calc. for C₁₄ H₂₆ N₂ O₄ S₄F₆: C, 31.81; H, 4.96; N, 5.30).



7.51 g ($5.04 \cdot 10^{-2}$ mol) of 5-bromopentene, used as received from Aldrich (99 % purity), was added drop wise, under nitrogen atmosphere, to a solution of 4 g ($4.03 \cdot 10^{-2}$ mol) of 1-methylpiperidine, used as received from Sigma (98 % purity), and 10 cm³ acetonitrile. After the same procedure described above, 8.11 g of $[\text{MPPIP}]^+[\text{Tf}_2\text{N}]^-$ were obtained with a light-

yellow color (yield: 76.8%). K-10 montmorillonite, used as received from Fluka Chemical Company, was modified by adding ZnCl_2 (2 mmol per gram of clay) using a wet-impregnation procedure⁴⁷. The catalyst was activated by heating at 150 °C in a vacuum oven for 24 hours. 5 g of the catalyst was mixed with 30 ml of 1,2-dichloroethane and 15 mmol of dimethyl disulfide. A solution of 6.69 mmol of $[\text{MPPIP}]^+[\text{Tf}_2\text{N}]^-$ in 5ml of 1,2-dichloroethane was added into the mixture. After filtering, removal of the solvents by rotary evaporation, the product was set under high vacuum for 24 h. 3.25 g (yield: 89.5 %) were obtained as a light yellow liquid.

¹H NMR (400 MHz, acetone-*d*₆): δ 1.661-1.741 (2H, m, CH₂), 1.889-2.041 (6H, m, CH₂), 2.062-2.079 (3H, s, S-CH₃), 2.110-2.206 (2H, m, CH₂), 2.622-2.694 (H, m, S-CH), 2.769-2.906 (2H, m, S-CH₂), 3.226-3.280 (3H, t, N-CH₃), 3.487-3.535 (2H, m, N-CH₂), 3.540-3.616 (4H, m, N-CH₂).

(Found: C, 33.59; H, 5.07; N, 5.52. Calc. for C₁₅ H₂₈ N₂ O₄ S₄ F₆: C, 33.2; H, 5.2; N, 5.16).

Extraction

Extraction of metal ions in water was performed as follows:

First, a stock solution was prepared. Chloride metal salts were added to 500 cm³ of pure water. Concentration of metal ions in water ranged from 50 to 500 mg / kg.

A sample of 20 g of stock solution, was then taken and saturated with an ionic liquid. When the solubility in water of an ionic liquid is known, an amount of ionic liquid, slightly higher than the one corresponding to the solubility in water of the IL, was added to the solution. When the solubility in water of the ionic liquid was unknown, 1 % by weight of

ionic liquid was added to 20 g of stock solution. The resulting stock solution saturated in ionic liquid is referred to as the initial solution.

When the water content of an ionic liquid is high, an ionic liquid phase saturated in water is prepared. Water is added to an ionic liquid until two phases appear.

The extraction experiments were carried out by mixing a sample of initial solution (typically 5 g) with an ionic liquid. After the two phases are placed on a shaker for 48 h, they were centrifuged at 2500 RPM for 20 minutes. The two phases remain in contact for another 24 h to reach equilibrium.

3 g of the upper aqueous phase, referred to as final solution, was then removed for analysis.

Initial and final solutions are analyzed by inductively-coupled-plasma-optical-emission spectroscopy (ICP - OES). Distribution coefficients are calculated from:

$$D = \frac{C_{in} - C_{fin}}{C_{fin}} \frac{m_w}{m_{IL}}$$

where C_{in} is the concentration in mg / kg of metal M in the initial aqueous solution and C_{fin} is the concentration in mg / kg of metal M in the final aqueous solution; m_w and m_{IL} correspond to the mass of water and ionic liquid, respectively.

It was not possible to measure distribution coefficients below 0.05. For very high distribution coefficients, we report $D = 1000$.

Using [OMIM]⁺[BF₄]⁻, [3MOPYR]⁺[BF₄]⁻ or [4MOPYR]⁺[BF₄]⁻, extraction was studied for seven metal ions. For all other ionic liquids, fifteen metal ions were studied. Table 1 shows the ILs and metal ions used.

Experimental results, done in duplicates, agree within 5 %.

Results and discussion

Distribution coefficients of metal ions obtained with classical ionic liquids are discussed first followed by those obtained with task-specific ionic liquids.

All ionic liquids form two liquid phases when contacted with water. [4MOPYR]⁺[NfO]⁻ and [4MOPYR]⁺[Tf₂N]⁻ have solubilities in water of 0.065 and 0.022 wt%, respectively.²⁰ All other generic ionic liquids (not task-specific ionic liquids) exhibit relatively high solubilities in water, typically above 1 wt-% in water. These solubilities are too high for industrial application.

Tables 2 to 6 show distribution coefficients for metal ions in water.

Extraction with ionic liquids containing [BF₄]⁻

Table 2 shows distribution coefficients for metal ions extracted from water with [OMIM]⁺[BF₄]⁻, [3MOPYR]⁺[BF₄]⁻ and [4MOPYR]⁺[BF₄]⁻.

Except for mercury, distribution coefficients are low. D is 25 for mercury with [OMIM]⁺[BF₄]⁻. With [3MOPYR]⁺[BF₄]⁻ or [4MOPYR]⁺[BF₄]⁻; D for mercury is at least two orders of magnitude higher than those for other cations. The cation ring of the ionic liquid containing [3MOPYR]⁺ or [4MOPYR]⁺ has a significant influence on the extraction of Hg²⁺. That is known to form stable complexes HgCl₂, HgCl₃⁻ or HgCl₄²⁻, which may interact with [3MOPYR]⁺ or [4MOPYR]⁺ cations.

To investigate if ion-exchange occurs during extraction, single extraction was studied for mercury chloride in water with [3MOPYR]⁺[BF₄]⁻ or [4MOPYR]⁺[BF₄]⁻.

When a cation exchange occurs, the metal ion in the aqueous phase is exchanged with a cation of the ionic liquid. The metal ion is then extracted in the ionic liquid, while the IL

cation moves to the aqueous phase. (When anion-exchange occurs, the IL anion is exchanged with the metal counter ion.)

If cation exchange occurs when $[3\text{MOPYR}]^+[\text{BF}_4]^-$ or $[\text{OMIM}]^+[\text{BF}_4]^-$ is used for mercury extraction, an increased amount of $[3\text{MOPYR}]^+$ or $[\text{OMIM}]^+$ is expected to be found in the aqueous phase. If anion exchange occurs, we expect to find more $[\text{BF}_4]^-$ anion in the aqueous phase after extraction. ICP-OES data give the emission intensities of carbon and boron, in the aqueous phase before and after extraction experiment gives insight into the amount of IL cation (containing carbon atoms) and anion (containing boron atom), respectively, present into the aqueous phase.

To make sure the increase in carbon or boron is significant enough to be observed, we prepared highly concentrated solutions of mercury in water. The weight percent mercury in water was approximately 0.05, one third of the water solubility of ionic liquids containing $[\text{BF}_4]^-$ anion: (1.59 and 1.65) wt-% for $[\text{OMIM}]^+[\text{BF}_4]^-$ and $[3\text{MOPYR}]^+[\text{BF}_4]^-$, respectively.^{39,40}

Following single extraction of mercury in water with either $[3\text{MOPYR}]^+[\text{BF}_4]^-$ or $[\text{OMIM}]^+[\text{BF}_4]^-$, there was no increase of boron or carbon in the aqueous phase. Therefore, no ion exchange occurs during the extraction of mercury from water with $[3\text{MOPYR}]^+[\text{BF}_4]^-$ or $[4\text{MOPYR}]^+[\text{BF}_4]^-$.

Extraction with ionic liquids containing $[\text{Tf}_2\text{N}]^-$, $[\text{TfO}]^-$ or $[\text{NfO}]^-$

Table 3 shows distribution coefficients for 4 cations (i.e., Na^+ , Cu^{2+} , Tl^+ and Hg^{2+}) in water extracted with $[4\text{MOPYR}]^+[\text{Tf}_2\text{N}]^-$, $[4\text{MOPYR}]^+[\text{TfO}]^-$ or $[4\text{MOPYR}]^+[\text{NfO}]^-$. For all other metal ions (see Table 1), no extraction was observed with $[4\text{MOPYR}]^+[\text{Tf}_2\text{N}]^-$, $[4\text{MOPYR}]^+[\text{TfO}]^-$ or $[4\text{MOPYR}]^+[\text{NfO}]^-$.

Extracting properties of all ionic liquids are low. Ionic liquids containing $[\text{Tf}_2\text{N}]^-$ anions do not extract any cation.

For $[\text{4MOPYR}]^+[\text{TfO}]^-$ and $[\text{4MOPYR}]^+[\text{NfO}]^-$, extraction results are similar to those for $[\text{4MOPYR}]^+[\text{BF}_4]^-$. Only mercury is well extracted with these pyridinium ionic liquids.

Following extraction, we observed an increase in sulfur (from the anion of the ionic liquid) or carbon (from cation and anion) in the aqueous phase.

Considering the high solubility in water for $[\text{4MOPYR}]^+[\text{TfO}]^-$ (1.59 wt-%), nearly as large as that for $[\text{4MOPYR}]^+[\text{BF}_4]^-$ (1.65 wt-%), these two ionic liquids are not suitable for extraction. However, because $[\text{4MOPYR}]^+[\text{NfO}]^-$ exhibits a solubility of 0.065 wt-%, it may be a good extracting agent for mercury.

Extraction with task-specific ionic liquids

Following the work of Rodgers et al. and others,^{27,28,48} we have synthesized some task-specific ionic liquids for extraction of specific metal ions. To study the effect of both the cation ring and the functional group appended to that ring, we synthesized two types of task-specific ionic liquids.

First, we studied ionic liquids containing a functional nitrile group appended to either a pyridinium or a piperidinium cation. Second, we studied ionic liquids containing a functional disulfide group appended to a piperidinium or a pyrrolidinium cation. Figure 1 shows structures and abbreviations.

TSILs containing a nitrile group

Table 4 shows results for competitive extraction of 12 metal ions with [4MPYRCN]⁺[Tf₂N]⁻ or [MIPICN]⁺[Tf₂N]⁻. Table 5 shows results for the single extractions of aqueous silver nitrate or palladium chloride.

Synthesis and application of a similar TSIL containing a pyridinium ring and a nitrile group are described elsewhere.⁴⁸ In water, this TSIL was reported to form a stable complex with palladium cation.

In our work, TSILs with a nitrile functional group are expected to extract ions like nickel or zinc. Following a previous report, they are also expected to extract silver or palladium.⁴⁸ Regrettably, nickel and zinc ions are not efficiently extracted by the two TSILs used here. However, silver and palladium are completely removed from water.

For ionic liquids containing a nitrile group, results depend significantly on the cation. Except for silver and palladium, distribution coefficients for all metal ions obtained with [4MPYRCN]⁺[Tf₂N]⁻ are one order or magnitude higher than those obtained with [MIPICN]⁺[Tf₂N]⁻. Compared to distribution coefficients obtained with [4MOPYR]⁺[Tf₂N]⁻, the presence of a nitrile functional group in [4MPYRCN]⁺[Tf₂N]⁻ significantly improves extraction. For example, the distribution coefficient for mercury is only 0.22 in [4MOPYR]⁺[Tf₂N]⁻ but 1.35 in [4MPYRCN]⁺[Tf₂N]⁻.

All other metal ions are modestly extracted with [4MPYRCN]⁺[Tf₂N]⁻. Three metal ions have a distribution coefficient above unity: Tl⁺ (1.69), Na⁺ (1.26) and Hg²⁺ (1.35). All others have distribution coefficients between 0.30 and 1.00.

With [MIPICN]⁺[Tf₂N]⁻, results are quite different. Despite the presence of a nitrile group, most distribution coefficients are below unity.

Surprisingly, changing the cation ring appears to have an important effect on the extraction properties of TSILs containing a nitrile group on the IL's cation.

Both [4MPYRCN]⁺[Tf₂N]⁻ and [MIPCN]⁺[Tf₂N]⁻ appear to be very good and selective extracting agents for palladium and silver. Whereas all distribution coefficients shown in Table 4 are near or below 1, we observe high distribution coefficients for silver: 41 and 8.1 with [4MPYRCN]⁺[Tf₂N]⁻ and [MIPCN]⁺[Tf₂N]⁻, respectively. Here again, the presence of a piperidinium cation reduces the extracting properties of the TSIL. D for palladium shows a maximum at 1000.

When the initial solution of 200 ppm of palladium chloride was presaturated with [4MPYRCN]⁺[Tf₂N]⁻ or [MIPCN]⁺[Tf₂N]⁻, analysis of the initial solution by ICP-OES shows that almost all palladium was removed from the aqueous phase. It appears, therefore, that palladium ion forms a complex with the TSIL.⁴⁸

TSILs containing a disulfide group

To study an ionic liquid with specific attraction for mercury, we synthesized a new TSIL containing a disulfide group.

Because mercury and sulfur are known to interact strongly, we expect mercury ion to bind strongly with this functional group. Considering the structure of the functional group, it is also expected to be selective for ions of a certain size.

To investigate the effect of the cation ring on the distribution coefficients of metal ions, two cations, piperidinium and pyrrolidinium, were used here. Figure 1 shows structures and abbreviations.

Table 6 shows results for competitive extraction with [MPYRROS₂]⁺[Tf₂N]⁻ and [MPIP₂]⁺[Tf₂N]⁻.

As expected, mercury is completely extracted with both TSILs synthesized here. The selectivity towards mercury is very high; the distribution coefficient for cadmium is 0.17 and lead is not extracted at all.

Cu²⁺ is also completely removed from water. The selectivity for Cu²⁺ is also high because neither Ni²⁺ nor Zn²⁺ is extracted.

As for nitrile-functionalized ionic liquids, the cation ring has a significant influence on the extraction of metal ions. While mercury and copper are equally well extracted with [MPYRROS₂]⁺[Tf₂N]⁻ or [MPIPS₂]⁺[Tf₂N]⁻, other metal ions are somewhat extracted with [MPYRROS₂]⁺[Tf₂N]⁻ but not at all with [MPIPS₂]⁺[Tf₂N]⁻.

As shown in Table 6, [MPIPS₂]⁺[Tf₂N]⁻ extracts only mercury, copper, and tin ($D = 70.5$). All other metals are not extracted at all. For [MPYRROS₂]⁺[Tf₂N]⁻, however, low distribution coefficients (about 0.5) are obtained for most metals. Iron, zinc, cadmium and tin are more extracted with distribution coefficients 2.89, 0.89, 0.97 and 4.35, respectively.

Once again, the ionic liquid containing piperidinium is less efficient than that containing pyrrolidinium.

Conclusion

Ten ionic liquids were studied for metal-ion extraction from water. All ionic liquids (except [OMIM]⁺[BF₄]⁻) were newly synthesized for this work. Except for mercury, ionic liquids containing pyridinium cations without any functional group exhibit poor extracting properties; similar results were reported for imidazolium homologues¹². However, extraction of mercury is very high with ionic liquids containing pyridinium cations.

Task-specific ionic liquids containing a nitrile functional group appear to be a good and selective extracting agent for silver or palladium ions. Low distribution coefficients are obtained for all other ions studied here. The influence of the cation ring of the ionic liquid seems to have a significant effect on the extraction properties of TSILs. Compared to pyridinium ions, piperidinium ions reduce the extracting properties of a TSIL containing a nitrile functional group.

With a TSIL containing a disulfide group, mercury and copper were totally extracted from water. Once again, extraction of other metal ions is reduced when a piperidinium cation is used instead of a pyrrolidinium ion.

It appears that an ionic liquid containing piperidinium cations exhibits reduced extracting properties relative to other ionic liquids containing pyrrolidinium or pyridinium cations.

As observed by others, extraction of metal ions can be achieved by using TSILs with specific functional groups. However, in addition to the functional groups, the cation ring as well as the anion of an ionic liquid also has an influence on distribution coefficients. Pyridinium or pyrrolidinium cations enhance low distribution coefficients as compared to imidazolium or piperidinium cations. Using [TfO]⁻ or [NfO]⁻ instead of [Tf₂N]⁻ anions also leads to enhanced extraction for mercury.

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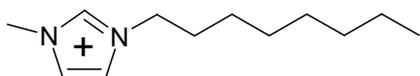
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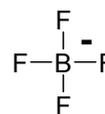
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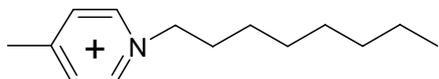
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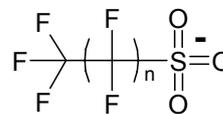
1-octyl-3-methylimidazolium cation [OMIM]⁺



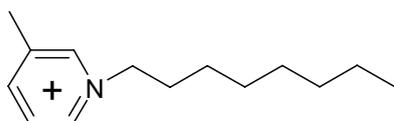
tetrafluoroborate anion [BF₄]⁻



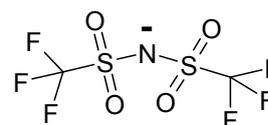
1-octyl-4-methylpyridinium cation [4MOPYR]⁺



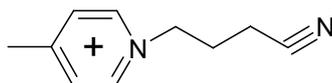
n=0 triflate anion [TfO]⁻
n=3 nonaflate anion [NfO]⁻



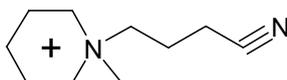
1-octyl-3-methylpyridinium cation [3MOPYR]⁺



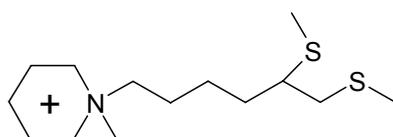
bis(trifluoromethylsulfonyl)imide anion [Tf₂N]⁻



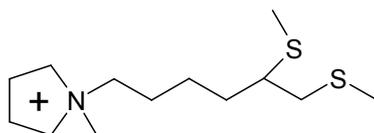
1-butyronitrile-4-methylpyridinium cation [4MPYRCN]⁺



1-butyronitrile-1-methylpiperidinium cation [MPIPCN]⁺



1-methyl-1-[4,5-bis(methylsulfide)]pentylpiperidinium cation [MPIPS₂]⁺



1-methyl-1-[4,5-bis(methylsulfide)]pentylpyrrolidinium cation [MPYRROS₂]⁺

Figure 1- Structures of cations and anions used in this work.

Table 1- Ionic liquids and cations in water studied here at pH 7 and 25°C

Extracting phase	Metal ions dissolved in water
[OMIM] ⁺ [BF ₄] ⁻ , [3MOPYR] ⁺ [BF ₄] ⁻ , [4MOPYR] ⁺ [BF ₄] ⁻	Na ⁺ , K ⁺ , Cs ⁺ , Zn ²⁺ , Cd ²⁺ , Hg ²⁺ , Pb ²⁺
[4MOPYR] ⁺ [Tf ₂ N] ⁻ , [4MOPYR] ⁺ [TfO] ⁻ , [4MOPYR] ⁺ [NfO] ⁻ ,	Na ⁺ , Mg ²⁺ , Fe ³⁺ , Co ²⁺ , Ni ²⁺ , Cu ²⁺ , Zn ²⁺ , Cd ²⁺ , Hg ²⁺ , Tl ⁺ , Pb ²⁺ , Sn ⁴⁺ , Ba ²⁺
[4MPYRCN] ⁺ [Tf ₂ N] ⁻ , [MPIPNCN] ⁺ [Tf ₂ N] ⁻ ,	Na ⁺ , Mg ²⁺ , Fe ³⁺ , Co ²⁺ , Ni ²⁺ , Cu ²⁺ , Zn ²⁺ , Cd ²⁺ , Hg ²⁺ , Tl ⁺ , Pb ²⁺ , Sn ⁴⁺ , Ag ⁺ ^a , Pd ²⁺ ^b
[MPYRROS ₂] ⁺ [Tf ₂ N] ⁻ , [MPIPS ₂] ⁺ [Tf ₂ N] ⁻	Na ⁺ , K ⁺ , Mg ²⁺ , Fe ³⁺ , Co ²⁺ , Ni ²⁺ , Cu ²⁺ , Zn ²⁺ , Cd ²⁺ , Hg ²⁺ , Tl ⁺ , Pb ²⁺ , Sn ⁴⁺ , Ba ²⁺

^a: Silver nitrate was extracted separately. ^b: Palladium chloride was extracted separately.

Table 2- Distribution coefficients *D* for cations in water extracted with [OMIM]⁺[BF₄]⁻, [3MOPYR]⁺[BF₄]⁻ and [4MOPYR]⁺[BF₄]⁻ at pH 7 and 25°C.

	[OMIM] ⁺ [BF ₄] ⁻	[3MOPYR] ⁺ [BF ₄] ⁻	[4MOPYR] ⁺ [BF ₄] ⁻
Na ⁺	0.03	a	a
K ⁺	0.06	0.05	a
Cs ⁺	0.15	a	a
Zn ²⁺	0.03	0.17	0.23
Cd ²⁺	0.04	0.31	1.71
Hg ²⁺	24.9	1000	1000
Pb ²⁺	11.9	11.0	8.66

^a: no extraction observed, *D* is below 0.05.

Table 3- Distribution coefficients D for cations in water extracted with $[4MOPYR]^+[Tf_2N]^-$, $[4MOPYR]^+[TfO]^-$, or $[4MOPYR]^+[NfO]^-$ at pH 7 and 25°C.

	$[4MOPYR]^+[Tf_2N]^-$	$[4MOPYR]^+[TfO]^-$	$[4MOPYR]^+[NfO]^-$
Na^+	a	a	0.54
Cu^{2+}	a	0.13	a
Tl^+	a	0.50	a
Hg^{2+}	0.22	1000	1000

^a: no extraction observed. D is below 0.05.

Table 4- Distribution coefficients D for cations in water extracted with $[4MPYRCN]^+[Tf_2N]^-$ or $[MIPCN]^+[Tf_2N]^-$ at pH 7 and 25°C.

	$[4MPYRCN]^+[Tf_2N]^-$	$[MIPCN]^+[Tf_2N]^-$
Na^+	1.26	0.05
Mg^{2+}	0.68	a
Fe^{3+}	0.98	0.06
Co^{2+}	0.54	a
Ni^{2+}	0.53	a
Cu^{2+}	0.83	a
Zn^{2+}	0.33	a
Cd^{2+}	0.39	a
Hg^{2+}	0.47	0.11
Tl^+	1.7	0.23
Pb^{2+}	0.64	a
Sn^{4+}	1.36	0.11

^a: no extraction observed. D is below 0.05.

Table 5- Distribution coefficients D for silver and palladium in water extracted with $[4MPYRCN]^+[Tf_2N]^-$ or $[MIPICN]^+[Tf_2N]^-$ at pH 7 and 25°C. Each metal ion was extracted separately. Silver nitrate salt and potassium palladium chloride were used.

	$[4MPYRCN]^+[Tf_2N]^-$	$[MIPICN]^+[Tf_2N]^-$
Ag^+	41	8.1
Pd^{2+}	1000	1000

Table 6- Distribution coefficients D for metal ions in water extracted with $[MPYRROS_2]^+[Tf_2N]^-$ or $[MIPIS_2]^+[Tf_2N]^-$ at pH 7 and 25°C.

	$[MPYRROS_2]^+[Tf_2N]^-$	$[MIPIS_2]^+[Tf_2N]^-$
Na^+	0.41	a
K^+	0.46	a
Mg^{2+}	0.37	a
Fe^{3+}	2.89	a
Co^{2+}	0.20	0.11
Ni^{2+}	0.19	a
Cu^{2+}	1000	1000
Zn^{2+}	0.85	a
Cd^{2+}	0.97	0.17
Hg^{2+}	1000	1000
Tl^+	0.68	a
Pb^{2+}	0.23	a
Sn^{4+}	4.35	70.5
Ba^{2+}	0.14	a

^a: no extraction observed. D is below 0.05.