Experimental and Computational Studies of Ruthenium(II)-Catalyzed Addition of Arene C–H Bonds to Olefins

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Hydroarylation reactions of olefins are catalyzed by the octahedral Ru(II) complex TpRu-CNMe(Ph)(CO) (1) (Tp = hydrideotris(pyrazolyl)borate). Experimental studies and density functional theory calculations support a reaction pathway that involves initial acetonitrile/olefin ligand exchange and subsequent olefin insertion into the ruthenium-phenyl bond. Metal-mediated C–H activation of arene to form a Ru–aryll bond with release of alkyl arene completes the proposed catalytic cycle. The cyclopentadienyl complex CpRu(PPh3)2(Ph) produces ethylbenzene and styrene from a benzene/ethylene solution at 90 °C; however, the transformation is not catalytic. A benzene solution of (PCP)Ru(CO)(Ph) (PCP = 2,6-(CH2PH2)2C6H3) and ethylene at 90 °C produces styrene in 12% yield without observation of ethylbenzene. Computational studies (DFT) suggest that the C–H activation step does not proceed through the formation of a Ru(IV) oxidative addition intermediate but rather occurs by a concerted pathway.

Introduction

The activation of C–H bonds using transition metal complexes has received attention due to the chemically inert nature of the C–H bond.1–4 The availability of aromatic compounds renders the development of methods for catalytic aromatic C–H functionalization an area of significance. For example, the production of alkyl arenes is performed on a large scale.5 Simple σ-Lewis acid catalysts can be used for the synthesis of alkyl arenes by activating olefins toward electrophilic addition to arenes;6,7 however, due to the formation of species with carboxylonic character, products with linear alkyl chains cannot be accessed with these traditional catalytic methods. Linear alkyl arenes can be produced in low yields by Friedel–Crafts alkylation of arenes using 1-haloalkanes or by acylation of arenes followed by Clemmensen reduction of the carbonyl fragment. Other methods for C–C bond-forming reactions involving aromatic molecules include transformations that typically incorporate aryl halides and the preparation of aryl halides often requires multistep syntheses.8–14 Whereas these methodologies have been widely incorporated for the synthesis of small organic molecules, C–C bond-forming reactions of aromatic molecules that proceed through metal-mediated C–H rather than C–X (X = halide) bond activation would expand the scope of available transformations and reduce the need for aryl halide syntheses.

Significant advances in catalytic metal-mediated C–H activation have been achieved for aromatic systems;2,15 however, such reactions are primarily limited to compounds that possess heteroatomic functionality tethered to the aromatic ring or carbonylation reactions of heteroaromatic substrates.16–24 Metal-catalyzed conver-
Most relevant to the present research are Ir-catalyzed termolecular olefin/alkyne hydroarylation.\textsuperscript{33,34} Catalytic Ru(III) and Au(III) complexes catalyze intra- and intermolecular olefin/alkyne hydroarylation of aromatic compounds across unsaturated C–C bonds.\textsuperscript{25} Recent computational study suggested that at high ethylene pressures multiple insertions of ethylene will occur,\textsuperscript{10} and that high ethylene pressures multiple insertions of ethylene will occur,\textsuperscript{10} which is consistent with the combined experimental and computational study of a Ru(III) complex.\textsuperscript{43} At elevated temperatures, TpRu(CO)(NCMe)(Ph) (Tp = hydridotris(pyrazolyl)borate) serves as a catalyst for the addition of arenes to ethylene. The results of several reactions are shown in Table 1. For example, under 25 psi of ethylene pressure, 0.1 mol % of TpRu(CO)(NCMe)(Ph) (Tp = hydridotris(pyrazolyl)borate) serves as a catalyst for the addition of arenes to ethylene. The results of several reactions are shown in Table 1. For example, under 25 psi of ethylene pressure, 0.1 mol % of \(1\) in benzene (90 °C) catalyzes the formation of ethylbenzene with 51 turnovers after 4 h. Using 0.1 mol % of \(1\), trace quantities of 1,3- and 1,4-diethylbenzene are also produced. After 4 h, the rate of ethylene pressure begins to slow; however, 77 turnovers are achieved after 24 h at 90 °C. Beyond this time, the production of ethylbenzene significantly decreases. Evidence of styrene formation has not been obtained for catalytic reactions of ethylene and benzene using between 0.1 and 1 mol % of complex \(1\) when the ethylene pressure is \(<1\) atm. A recent computational study suggested that at high ethylene pressures multiple insertions of ethylene will be observed.\textsuperscript{44} In accord with this prediction, reaction at 25 psi of ethylene pressure produces a small amount of butylbenzene (<1 turnover).

### Results and Discussion

**Catalytic Hydroarylation of Olefins.** At elevated temperatures, TpRu(CO)(NCMe)(Ph) (Tp = hydridotris(pyrazolyl)borate) serves as a catalyst for the addition of arenes to ethylene. The results of several reactions are shown in Table 1. For example, under 25 psi of ethylene pressure, 0.1 mol % of \(1\) in benzene (90 °C) catalyzes the formation of ethylbenzene with 51 turnovers after 4 h. Using 0.1 mol % of \(1\), trace quantities of 1,3- and 1,4-diethylbenzene are also produced. After 4 h, the rate of ethylene pressure begins to slow; however, 77 turnovers are achieved after 24 h at 90 °C. Beyond this time, the production of ethylbenzene significantly decreases. Evidence of styrene formation has not been obtained for catalytic reactions of ethylene and benzene using between 0.1 and 1 mol % of complex \(1\) when the ethylene pressure is \(<1\) atm. A recent computational study suggested that at high ethylene pressures multiple insertions of ethylene will be observed.\textsuperscript{44} In accord with this prediction, reaction at 25 psi of ethylene pressure produces a small amount of butylbenzene (<1 turnover).

### Table 1. Catalytic Addition of Arenes to Unsaturated Substrates (unless otherwise noted, reaction conditions are 90 °C, 25 psi of gas, 0.1 mol % of \(1\), 4 h)

<table>
<thead>
<tr>
<th>Arene</th>
<th>Unsaturated Substrate</th>
<th>TON</th>
<th>TOF a</th>
<th>Products</th>
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<tr>
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<td>ethylene</td>
<td>51 (77) b</td>
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<td>acetylene\textsuperscript{d}</td>
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</table>

\(\text{a mol } 1^{-1} \cdot \text{s}^{-1}\). \(\text{b Turnovers observed after } 24 \text{ h are given in parentheses.}\text{ c Trace quantities of 1,3- and 1,4-diethylbenzene are also produced.}\text{ d At 250 psi of ethylene.}\text{ e }50 \text{ equiv based on } 1, \text{ after } 6 \text{ h.}\text{ f }24 \text{ h.}\text{ g }10 \text{ psi.}

\(25^\circ \text{C, } 25 \text{ psi of gas, } 0.1 \text{ mol % of } 1, \text{ 4 h).}\)

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\(\text{(33) Yoon, S. W.; Pastine, S. J.; Sames, D. Org. Lett. 2004, 6, 581–584.}\)
In contrast to Lewis acid-catalyzed Friedel–Crafts reactions, complex 1 is mildly selective for linear over branched alkyl benzene product. For example, the reaction of propane and benzene at 25 psi (90 °C) yields a 1.6:1 ratio of linear propyl benzene to cumene. In addition, the formation of trans-β-methylstyrene is observed in nearly quantitative yield based on complex 1. After 4 h at 90 °C, catalyst activity ceases. The reaction of 1-hexene and benzene (90 °C, 1 mol % of 1) results in the production of linear 1-phenylhexane and branched 2-phenylhexane in a 1.7:1.0 ratio along with formation of trans-1-phenylhex-1-ene. In contrast to an Ir(III) system that catalyzes similar reactions, 42 evidence for the isomerization of 1-hexene to internal olefins has not been obtained for catalysis with complex 1. Accordingly, detectable quantities of 3-phenylhexane are not produced in the catalytic reactions of benzene with 1-hexene. The combination of 0.1 mol % of 1 in benzene with isobutylene does not yield new organic products after 24 h at 90 °C. The reaction of benzene and acetylene with 1 mol % of 1 yields only trace quantities of styrene after 24 h at 90 °C.

Complex 1 can also catalyze reactions of monoalkylated arenes. For example, the combination of 0.1 mol % of 1 with ethylbenzene and ethylene results in the formation of 1,3-diethylbenzene and 1,4-diethylbenzene. No evidence for the formation of the ortho disubstituted product 1,2-diethylbenzene is observed. The rate of conversion of ethylbenzene to diethylbenzenes is approximately 7 times slower than the rate of conversion of benzene to ethylbenzene. For example, after 2 h at 90 °C with 0.1 mol % of 1, 43 turnovers of benzene to ethylbenzene are observed. For the identical reaction in ethylbenzene, after 2 h a total of approximately 6 turnovers to 1,3- and 1,4-diethylbenzene are observed.

We have previously reported that TpRu(II) complexes catalyze radical polymerization reactions of electron-deficient olefins. 46 Thus, efficient catalytic hydroarylation reactions of olefins such as methyl methacrylate, styrene, or acrylonitrile are not possible due to radical polymerization side reactions. Attempted hydrophenylation of methylvinyl ketone resulted in a previously reported thermal cyclodimerization reaction. 46

Simple Lewis acids are known to catalyze the net addition of arene C–H bonds across olefin C=C bonds by activating the olefin toward carbanion reactivity (i.e., electrophilic addition to the arene). 45 However, the Ru-catalyzed reactions exhibit characteristics that are inconsistent with classic Friedel–Crafts chemistry. The primary evidence against the TpRu(II) complex acting as a Friedel–Crafts catalyst is the mild selectivity of linear over branched products in the reactions of benzene with the α-olefins propane and 1-hexene. In addition, Friedel–Crafts reactions typically produce 1,2-dialkyl products in addition to 1,3- and 1,4-dialkyl benzenes. TpRu(CO)(NCO)(Me)(Ph) (1) does not produce 1,2-diethylbenzene in detectable quantity during the reaction of ethylene with ethylbenzene. Presumably, the lack of formation of 1,2-diethylbenzene using complex 1 as catalyst is due to steric selectivity in the metal-mediated C–H activation step (see below) similar to that directly observed in stoichiometric aromatic C–H oxidative addition reactions. 47 Heating (90 °C) a 1:1 molar mixture of 1,4-diethylbenzene and benzene in the presence of 1 does not result in the formation of ethylbenzene (after 24 h) as is often observed for Friedel–Crafts catalysts. The relative rate of reaction of benzene versus monoalkylated benzene is contrary to observations made for most Lewis acid-catalyzed Friedel–Crafts reactions. For example, catalytic addition of C–H bonds to ethylene using 1 as catalyst is approximately 7 times more rapid for benzene than ethylbenzene (based on analysis after 24 h of reaction). Finally, the lack of catalytic activity for TpRu(CO)(Me), a precursor to complex 1, indicates that minor amounts of impurities from commercial starting reagents are not responsible for the catalysis.

### Proposed Catalytic Cycle

A proposed catalytic cycle is shown in Scheme 1 using ethylene as the olefinic substrate. The first step in the proposed cycle involves dissociation of acetonitrile. Heating a CDCl3 solution of 1 in the presence of CD3CN results in a decrease in the intensity of the resonance (1H NMR) of the coordinated NCCCH3 and the appearance of a singlet consistent with free NCCCH3. In CD2CN at 90 °C the exchange reaction has a half-life of approximately 15 min. The ability of the acetonitrile ligand to undergo exchange suggests that proposed ethylene/acetonitrile ligand exchange is feasible under catalytic conditions (Scheme 1). The lack of catalysis with disubstituted olefins (e.g., isobutylene) is likely explained by steric inhibition of olefin coordination in this first step.

Insertion of ethylene into the Ru–phenyl bond (second step in the proposed catalytic cycle) provides a pathway for C–C bond formation. Evidence for olefin insertion into the Ru–phenyl bond (second step in the proposed catalytic cycle) provides a pathway for C–C bond formation. Evidence for olefin

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insertion has been obtained from the isolation of TpRu(CO)(NCMe)(CH2CH2Ph) (2) in approximately 50% yield upon reaction of TpRu(CO)(NCMe)(Ph) with ethylene (250 psi) in acetonitrile (eq 1). TpRu(CO)(NCMe)(CH2CH2Ph) (2) has been independently prepared and characterized (see below).

For the putative complex TpRu(CO)(CO)(Ph)(η²-olefin) that precedes the insertion step, the presence of the strong π-acid CO likely results in a preferred olefin orientation in which the C=C bond is parallel to the Ru–phenyl bond. In the proposed olefin orientation, which is supported by DFT calculations (see below), π-back-bonding interactions are optimized (Scheme 2), and the C=C bond is aligned for the subsequent insertion step. Insertions of olefins into late transition metal–aryl bonds have precedent. For example, the Heck reaction involves alkene insertion into Pd–aryl bonds as a key reaction step.8,9 In addition, stoichiometric ethylene insertion into Ru(II) phenyl bonds and Ru(II)-catalyzed Heck reactions have been reported.48–52

For α-olefins, the regioselectivity of olefin insertion would determine the selectivity of branched versus linear alkyl arene products, and the observed selectivity for linear over branched alkyl arenes suggests that 2,1-insertion (product I in Scheme 3) is slightly favored.

A computational study by Oxgaard and Goddard suggests that the regioselectivity of α-olefin insertion for TpRu(CO)(CO)(Ph)(η²-H2C=C(CH3)Me) is controlled by steric interactions between the alkyl group of the olefin and pyrazolyl rings.44 The linear-to-branched ratios for the hydrophenylation of propene and 1-hexene catalyzed by Ir(III) are nearly identical to that using complex 1 as catalyst.40 However, the regioselectivity of propene hydrophenylation mediated by CpRu(PPh3)2(Ph) of 5:1 (linear:branched ratio, see below) indicates that the similarity between the Ir(III) catalyst and TpRu(CO)(NCMe)(Ph) (1) is likely coincidental rather than implicating that electronic factors during insertion dictate the selectivity. Heck reactions involving the insertion of α-olefins with alkyl substituents into Pd–Ph bonds, as well as Ru-catalyzed Heck reactions that incorporate acrylates, are selective for 2,1-insertion.51–53

Ethylene insertion into the Ru–phenyl bond of TpRu(CO)(CO)(η²-H2C=C(CH3)2Ph) creates an open coordination site that can bind ethylene to yield the proposed catalyst resting state TpRu(CO)(η²-H2C=C(CH2)2CH2Ph). The formation of butylbenzene at high ethylene pressure (250 psi, see above) indicates that a second insertion of ethylene is competitive with arene C–H activation under conditions that increase the concentration of the putative catalyst resting state TpRu(CO)(η²-H2C=C(CH2)2CH2PH). Between 15 and 60 psi of ethylene, the rate of catalytic addition of benzene to ethylene decreases with increasing ethylene pressure (Figure 1), and the inhibition of catalysis upon increasing ethylene pressure is consistent with TpRu(CO)(η²-H2C=C(CH2)2CH2Ph) as the catalyst resting state. Deriving a rate law for the proposed steps in catalytic C–H activation (Scheme 4, L = ethylene) yields eq 2. If $k_1[C_6H_6] / k_2$ is then the reaction rate should exhibit an inverse first-order dependence on ethylene concentration. Similar to the hydrophenylation of ethylene using 1 as catalyst, an inverse dependence on ethylene pressure

$$\text{Rate} = \frac{k_1k_2[\text{C}_6\text{H}_6]}{k_1[\text{C}_6\text{H}_6] + k_2[\text{C}_6\text{H}_6]}$$  (2)

[eq 2]
Ru(II)-Catalyzed Addition of C–H Bonds to Olefins

Figure 2. Dependence of the rate of benzene addition to 1-hexene on concentration of 1-hexene (0.1 mol % of 1, 70 °C).

Scheme 4. Proposed Pathway for Benzene C–H Activation by TpRu(CO)(R)(L) (R = Me or CH₂CH₂Ph and L = NCMe or ethylene)

has been reported for Ir-catalyzed formation of ethylbenzene at high ethylene/benzene ratios. Unfortunately, attempts to isolate the proposed catalyst resting state have failed. Monitoring a catalytic reaction at 20 psi of ethylene in an NMR tube reaction allows observation of a new TpRu system that exhibits resonances consistent with the proposed resting state; however, the presence of the catalytic precursor, ethylbenzene, and catalyst decomposition make definitive assignment of this complex difficult.

For catalytic reactions of 1-hexene/benzene using complex 1, at low 1-hexene concentrations the rate of catalysis increases with increasing concentration of 1-hexene; however, the reaction rate reaches an apex at approximately 80 equiv (based on 1) of 1-hexene. Beyond this point, increasing the concentration of 1-hexene results in a decrease in the rate of catalysis (Figure 2). These results suggest that at low concentrations of 1-hexene the catalyst resting state is TpRu(CO)(NCMe)(Ph) (1), and IR spectroscopy of a catalytic reaction in progress reveals a CO absorption identical to that of complex 1. At higher concentrations of 1-hexene, the catalyst resting state likely shifts to TpRu(CO)(NCMe)(H)(R) (R = hexylphenyl isomers).

The final step of the proposed catalytic cycle is Ru-mediated C–H activation with release of alkyl benzene, and the proposed metal-mediated C–H activation is consistent with the assertion that the reaction proceeds via a non-Friedel Crafts mechanism. Performing the catalytic addition of C₆D₆ to ethylene (C₂H₄) using complex 1 results in the formation of d₅-PhCH₂CH₂CD₃ as indicated by ¹H NMR spectroscopy and mass spectrometry (eq 3). For example, the ¹H NMR spectrum of the catalyst solution (C₆D₆) after 12 h at 70 °C reveals a multiplet at 1.1 ppm due to the monodeuterated methyl group of d₅-ethylbenzene as well as a triplet of triplets (JHH ≈ 7 Hz and JHD ≈ 1 Hz) at 2.4 ppm due to the methylene group. In addition, approximately 1 equiv (based on concentration of 1) of PhCH₂CH₂D is formed due to the protio-phenyl group of the catalyst 1.

Additional evidence of the ability of {TpRu(CO)(R)} (R = alkyl) fragments to initiate arene C–H activation comes from the stoichiometric reaction of TpRu(CO) (Me)(NCMe) with benzene. In the presence of acetonitrile, the reaction of TpRu(CO)(Me)(NCMe) with benzene produces TpRu(CO)(NCMe)(Ph) (1) and methane. The formation of methane via C–H activation of benzene is confirmed by the observation of CH₂D in reactions performed in sealed NMR tubes using C₆D₆ (eq 4). The complex TpRu(PPh₃)(NCMe)(H) has been reported to initiate H/D exchange between methane and deuterated solvents. The rate of conversion of TpRu(CO)(Me)(NCMe) and benzene to complex 1 and methane is suppressed by increasing the concentration of acetonitrile. Thus, access to the coordinatively unsaturated fragment {TpRu(CO)(R)} (R = alkyl) is apparently necessary for benzene C–H activation (Scheme 4).

A kinetic isotope effect (KIE) was determined by analysis of reaction products from a catalytic reaction of ethylene in a 1:1 ratio of C₆H₆ and C₆D₆. The possible isotopic distributions of products and their molecular weights are shown in Scheme 5. Comparison of the ratios of peaks due to the stoichiometric reaction of TpRu(CO)(Me)(NCMe) with benzene produces TpRu(CO)(NCMe)(Ph) (1) and methane. The formation of methane via C–H activation of benzene is confirmed by the observation of CH₂D in reactions performed in sealed NMR tubes using C₆D₆ (eq 4). The complex TpRu(PPh₃)(NCMe)(H) has been reported to initiate H/D exchange between methane and deuterated solvents. The rate of conversion of TpRu(CO)(Me)(NCMe) and benzene to complex 1 and methane is suppressed by increasing the concentration of acetonitrile. Thus, access to the coordinatively unsaturated fragment {TpRu(CO)(R)} (R = alkyl) is apparently necessary for benzene C–H activation (Scheme 4).

A kinetic isotope effect (KIE) was determined by analysis of reaction products from a catalytic reaction of ethylene in a 1:1 ratio of C₆H₆ and C₆D₆. The possible isotopic distributions of products and their molecular weights are shown in Scheme 5. Comparison of the ratios of peaks due to methane gives M₁/M₂ = 112/111 allowing a determination of the kinetic isotope effect for the carbon–hydrogen bond-breaking step. The ratio reveals a kinetic selectivity for C–H activation over C–D activation by 2.1(1).

A control experiment using a 1:1 mixture of C₆H₆/C₆D₆ in the presence of 1 mol % of complex 1 (in


the absence of ethylene) at 90 °C reveals isotopic scrambling in the benzene. For example, after 2 h the ratio of molecular weight 78/79 (C6H6/C6D6) decreased by 30%. During the same time period, the ratio of C6D6/C6H6H undergoes an analogous change. These results indicate that the Ru(II) phenyl complex 1 catalyzes isotopic scrambling between C6H6 and C6D6 (the abundance of other mixed isotopomers also increases). However, monitoring the ratios of peaks due to molecular weights 78/79 (C6H6/C6D6) and 84/85 (C6D5H/C6D6) during the catalytic reaction reveals that isotope scrambling between C6H6 and C6D6 under ethylene pressure (25 psi) is negligible. The significant reduction in the extent of isotopic scrambling between C6H6 and C6D6 in the presence of ethylene indicates that monitoring the molecular weights of ethylbenzene products under catalytic conditions provides a reliable method to approximate the KIE for the catalytic reaction.

It is anticipated that the kinetic isotope effect for the catalytic reaction should be similar to the kinetic isotope effect for benzene C–H activation by TpRu(CO)(Me)-(NCMe) to produce methane and TpRu(CO)(Ph)(NCMe). The reaction of TpRu(CO)(Me)(NCMe) in a 1:1 molar mixture of C6H6 and C6D6 in a gastight NMR tube at 90 °C for 30 min produces resonances for both CH4 and CH3D (eq 5). Integration of these two resonances reveals a kinetic isotope effect of 2.5(5). The large deviation is due, in part, to the poor signal-to-noise that results from the necessity of analyzing the reaction before significant isotopic scrambling. GC–MS of the reaction solution after 30 min of reaction reveals that approximately 10% of the C6H6 and the C6D6 have undergone isotopic scrambling. Although analysis at longer time periods would improve signal-to-noise of the resonances due to methane, isotopic scrambling would complicate analysis of the KIE. Even though the isotopic scrambling after 30 min of reaction does not allow a quantitative determination of the KIE, the KIEs for the stoichiometric benzene C–H activation by TpRu(CO)(NCMe)(Me) and the catalytic hydrophenylation of ethylene are identical within deviation.

**Study of β-Hydride Elimination.** The previously reported dicarbonyl complex [TpRu(CO)3(THF)][PF6] reacts with PhCH2CH2MgCl to yield TpRu(CO)2(Ph)(CH2CH2Ph) (3) (eq 6). Complex 3 has been characterized by IR and 3H and 13C NMR spectroscopy with νCO = 2025 and 1954 cm⁻¹ (IR spectrum), methylene resonances at 3.38 and 1.93 ppm (1H NMR), and 1H/13C NMR spectra consistent with C13 molecular symmetry.

Oxidative removal of a carbonyl ligand from complex 3 upon addition of Me3NO in refluxing acetonitrile yields the monocarbonyl product TpRu(CO)(NCMe)(CH2CH2Ph) (2) (eq 7). The conversion of 3 to 2 is characterized by the disappearance of the CO stretching absorptions due to 3 and the appearance of a single CO absorption at 1917 cm⁻¹. Complex 2 has been characterized by a single-crystal X-ray diffraction study. The ORTEP of 2 is displayed in Figure 3, and data collection parameters are listed in Table 2. The structure reveals an asymmetric complex with an approximately octahedral coordination sphere. The Ru–alkyl carbon bond distance is 2.115(4) Å.

**Table 2. Selected Crystallographic Data and Collection Parameters for TpRu(CO)(NCMe)(CH2CH2Ph) (2)**

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<thead>
<tr>
<th>Parameter</th>
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**Figure 3. ORTEP of TpRu(CO)(NCMe)(CH2CH2Ph) (2)** (30% probability with hydrogen atoms omitted). Selected bond lengths (Å): Ru–N7, 2.020(3); Ru–C13, 2.115(4); Ru–N1, 2.062(3); Ru–N3, 2.138(3); Ru–N5, 2.191(3); Ru–C12, 1.819(4); N7–C10, 1.128(5); C13–C14, 1.474(6); C14–C15, 1.521(7). Selected bond angles (deg): Ru–C13–C14, 119.0(3); Ru–N7–C10, 170.6(3); Ru–C12–O1, 179.2(4); C12–Ru–C13, 90.85(16); N1–Ru–N3, 84.95(12); N1–Ru–N5, 86.71(12); N3–Ru–N5, 83.90(11); C13–Ru–N7, 88.94(14).

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<table>
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</tbody>
</table>

Consistent with the proposed catalytic pathway (Scheme 1), the reaction of complex 2 with C6D6 at 90 °C results in the production of PhCH2CH2D and TpRu(CO)(NCMe)(Ph-d5)(1-d5) as determined by 1H NMR spectroscopy (eq 8). In addition, the combination of 1 mol % of 2 in benzene under 25 psi (90 °C) results in the catalytic production of ethylbenzene. The proposed reaction mechanism for the addition of benzene to ethylene involves the formation of the coordinatively unsaturated \{TpRu(CO)(CH2CH2Ph)\} system that would likely be susceptible to a \(\beta\)-hydride elimination reaction to produce TpRu(CO)(H)(2-styrene); however, styrene formation is not observed during catalytic hydrophenylation reactions of ethylene at ethylene pressures up to 60 psi. The lack of observation of styrene indicates that either \(\beta\)-hydride elimination is not kinetically competitive or that dissociation of styrene from TpRu(CO)(H)(2-styrene) has a substantial activation barrier (i.e., \(\beta\)-hydride elimination is reversible). We sought to explore \(\beta\)-hydride elimination from the putative intermediate \{TpRu(CO)(CH2CH2Ph)\} in more detail using TpRu(CO)(NCMe)(CH2CH2Ph) (2). Heating (70 °C) a CDCl3 solution of TpRu(CO)(NCMe)(CH2CH2Ph) results in the quantitative production (by 1H NMR spectroscopy) of styrene, CHDCl2, and TpRu(CO)(NCMe)(Cl) (4) after 30 min. Complex 4 has been independently prepared upon reaction of HCl with TpRu(CO)(NCMe)(Ph) (Scheme 6). The formation of 4 likely occurs via \(\beta\)-hydride elimination to produce (TpRu(CO)(H)(2-styrene)) followed by dissociation of styrene and conversion of the Ru–H bond to a Ru–Cl bond upon reaction with CDCl3 (regardless of reaction order; Scheme 6). The rapid conversion of complex 2 to complex 4 indicates that \(\beta\)-hydride elimination is kinetically facile, and we attribute the lack of observation of styrene in catalytic ethylene hydrophenylation reactions to reversible \(\beta\)-hydride elimination. In CDCl3, TpRu(CO)(H)(2-styrene) is likely trapped by the chlorinated solvent; however, attempts to prepare and isolate TpRu(CO)(H)(2-styrene) have failed. The observations of trans-1-phenylhex-1-ene and trans-1-methylstyrene during the hydrophenylation of 1-hexene and propene, respectively, are likely attributable to more facile dissociation of the disubstituted olefins that are formed from \(\beta\)-hydride elimination. In CDCl3, TpRu(CO)(H)(2-styrene) is likely trapped by the chlorinated solvent, however, attempts to prepare and isolate TpRu(CO)(H)(2-styrene) have failed. The observations of trans-1-phenylhex-1-ene and trans-1-methylstyrene during the hydrophenylation of 1-hexene and propene, respectively, are likely attributable to more facile dissociation of the disubstituted olefins that are formed from \(\beta\)-hydride elimination (Scheme 7).

Reactions of CpRu(PPh3)2(Ph) and (PCP)Ru(CO)(Ph).

To study the generality of the catalytic reactions, we have begun to synthesize and explore the catalytic properties of closely related Ru(II) systems. It has previously been reported that CpRu(PPh3)2(Me) (Cp = cyclopentadienyl) undergoes an intramolecular C–H activation at 90 °C to yield the orthometalated product CpRu(PPh3)(CH2PC6H4Me) and methane.56,57 CpRu(PPh3)2(Me) is isoelectronic with TpRu(CO)(NCMe)(Me), and loss of phosphine from CpRu(PPh3)2(Ph) complex provides the coordinatively unsaturated complex \{CpRu(PPh3)(Ph)\}, which is isoelectronic to the proposed TpRu(II) species that initiates arene C–H activation. At 70 °C, CpRu(PPh3)2(Ph) reacts with benzene to yield the previously reported phenyl complex CpRu(PPh3)2(Ph) and methane (eq 9).58 The formation of methane is confirmed by resonances consistent with CH2D upon reaction of CpRu(PPh3)2(Ph) with C6D6 in a sealed NMR tube. Dissolution of CpRu(PPh3)2(Ph) in benzene under 25 psi of ethylene at 90 °C results in the production of ethylbenzene; however, the reaction is not catalytic, as only 0.12 equiv of ethylbenzene is produced per equivalent of ruthenium. The formation of styrene in 31% yield is also observed for the CpRu(II) system (eq 10). The more facile production of styrene by the CpRu system compared with the TpRu system might be explained by...
the ability of Cp to undergo a ring slip to provide a pathway for associative ligand exchange after \(\beta\)-hydride elimination. Thus, an \(\eta^2\) to \(\eta^3\) ring-slip would provide an open coordination site for binding of ethylene (or phosphine), and subsequent dissociation of styrene is likely to be facile.\(^{59}\) Although the production of olefin accounts for a portion of the decomposition of the CpRu complex,\(^{31}\) P NMR spectroscopy indicates that the fate of the remaining Ru (\(\sim 70\%\)) is dispersed between multiple uncharacterized complexes. Thus, the CpRu-(PPh\(_3\))\(_2\)(Ph) system apparently decomposes under catalytic conditions. A benzene solution of CpRu(PPh\(_3\))\(_2\)(Ph) under 25 psi of propene at 70 °C produces less than stoichiometric quantities of cumene and \(n\)-propylbenzene. Using 5 mol % of CpRu(PPh\(_3\))\(_2\)(Ph), an approximate 5:1 molar ratio of \(n\)-propylbenzene to cumene is produced (eq 11).

The reaction of (PCP)Ru(CO)(OTf) (PCP = 2,6-(CH\(_2\))\(_2\)Bu\(_2\))C\(_6\)H\(_3\)) with PhLi allows the isolation of (PCP)Ru(CO)(Ph) (5) (eq 12). We anticipated complex 5 as a potential catalyst for the hydroarylation reactions since it has similar features to complex 1. For example, complex 5 possesses a tridentate (six-electron donating) monoanionic ligand (cf. the ligand Tp) as well as Ru-CO and Ru-Ph bonds. In addition, an open coordination site for olefin binding is available since complex 5 is coordinatively and electronically unsaturated. Complex 5 differs from 1 in that the tridentate ligand is coordinated in a meridional fashion (versus facial for the Tp ligand) and the metal center is a more proficient \(\pi\)-base (\(\nu_{CO} = 1900\) cm\(^{-1}\) for 5 compared with 1935 cm\(^{-1}\) for 1). A 0.5 mol % solution of 5 in benzene under 25 psi of ethylene pressure at temperatures from 90 to 150 °C failed to produce detectable quantities of ethylbenzene; however, styrene was observed in 10–15% yield based on complex 5. Although the source of the lack of activity for olefin hydroarylation using 5 has not been definitively determined, the steric bulk due to the presence of four tBu groups as well as the potential for a trans orientation of coordinated olefin and the phenyl ligand could contribute to the poor activity. Consistent with the former, the coordination of Lewis bases (e.g., ammonia) to the (PCP)Ru(CO)(R) framework is particularly weak.\(^{60}\)

**Computational Studies.** In concert with the experimental studies, the catalytic hydroarylation of ethylene has been studied at the B3LYP/SBK(d) level of theory. As a model of the full tris(pyrazolyl)borate (Tp) ligand, the tris(azo)borate (Tab) ligand, [HB(\(N=NH\))\(_3\)], was used. In previous research, Tab was shown to reproduce the structure and energetics of the full Tp models for C–H activation potential energy surfaces.\(^{61}\) In addition, the parent nitrile HCN was substituted for acetonitrile. The C–H activation event was broken down into three steps: benzene coordination, activation of benzene with elimination of methane, and coordination of HCN to the resulting phenyl complex (Scheme 8). Oxgaard and...
Goddard have recently communicated a computational study of the catalytic hydrophenylation of ethylene by TpRu(CO)(NCMe)(Ph); in cases where Oxgaard and Goddard studied similar reactions, computed enthalpies were comparable to those obtained with the smaller model employing Tab and HCN instead of Tp and MeCN, respectively.44

Benzene C–H Activation. The displacement of HCN by benzene is endergonic by 16.1 kcal/mol. Given a calculated free energy of binding of HCN to ((Tab)Ru(Me)(CO)) of 14.9 kcal/mol, this corresponds to binding free energy to ((Tab)Ru(Me)(CO)) of +1.2 kcal/mol for benzene and is consistent with the anticipated weak binding of a π-heteroaromatic system to a closed-shell, d⁶ ML₅ fragment. Subsequent C–H activation of benzene to produce methane and ((Tab)Ru(CO)(Ph)) is exergonic by 6.6 kcal/mol with a calculated free energy barrier of 21.2 kcal/mol. The coordination of nitrile to ((Tab)Ru(CO)(Ph)) is favorable by −13.7 kcal/mol, and the overall transformation of (Tab)Ru(CO)(Me)(NCH) and benzene to (Tab)Ru(CO)(Ph)(NCH) and methane is exergonic by 4.2 kcal/mol. If the overall entropy change for benzene C–H activation by (Tab)Ru(CO)(Me)(NCH) is considered negligible, using a bond dissociation energy of methane of 105 and 113 kcal/mol for benzene indicates that the Ru–phenyl bond is approximately 12 kcal/mol stronger than the Ru–methyl bond.62

Ethylene Insertion. In the proposed catalytic cycle, dissociation of acetonitrile produces (TpRu(CO)(Ph))(η²-ethylene). Subsequent ethylene insertion provides a route for C–C bond formation, and these reaction steps were also studied using DFT with ((Tab)Ru(CO)(Ph)) Scheme 9). The free energy of ethylene binding to (Tab)Ru(CO)(Ph) is −10.4 kcal/mol. The minimized structure of (Tab)Ru(CO)(Ph)(η²-ethylene) is shown in Figure 4. Consistent with predictions based on dπ back-bonding (see above), the ethylene C=C bond is oriented approximately parallel to the Ru–Ph bond. Combining the calculated olefin binding free energies with the HCN dissociation energies yields a free energy for HCN/ethylene exchange of −3.3 kcal/mol. The more favorable coordination of ethylene is consistent with the suggestion that TpRu(CO)(CH₂CH₂Ph)(η²-ethylene) is the catalyst resting state for the hydrophenylation of ethylene. The calculated ethylene insertion barrier is 18.6 kcal/mol, and the insertion process is exergonic by 5.5 kcal/mol.

Agostic Interactions and Resting State. The β-phenethyl complex, (Tab)Ru(CO)(CH₂CH₂Ph), which is the product of ethylene insertion into the Ru–Ph bond of ((Tab)Ru(CO)(Ph)), has a weak π-interaction involving the phenyl substituent, as evidenced by the calculated short distances between the ruthenium and ipso carbon (Ru–C = 2.75 Å) as well as one of the ortho carbons (Ru–C = 2.61 Å). The minimized structure of (Tab)Ru(CO)(CH₂CH₂Ph) is shown in Figure 5. This conformation leads to a Ru–Cα–Cβ–Cipso dihedral angle of −32°. An alternative conformation with this dihedral angle

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**Scheme 9.** Calculated Energetics for Ethylene Coordination to (Tab)Ru(CO)(Ph), Ethylene Insertion into the Ru–Ph Bond, and β-Hydride Elimination with Energy Differences Given in kcal/mol

![Scheme 9](image)

**Figure 4.** B3LYP/SBK(d) calculated structure of (Tab)Ru(CO)(Ph)(η²-ethylene). Atoms of the Tab ligand are shown in wire frame for clarity.

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Figure 5. B3LYP/SBK(d) calculated structure of (Tab)Ru(CO)(CH2CH2Ph). Atoms of the Tab ligand are shown in wireframe for clarity.

equal to 180° was constructed, and geometry optimized at the same level of theory, resulting in a β-agostic (RuH = 1.93 Å; RuCα = 2.44 Å; RuHCβ = 100°) interaction, which was calculated to be only 1.3 kcal/mol higher than the π-bonded conformation. The binding of ethylene to [(Tab)Ru(CO)(CH2CH2Ph)] to form the proposed resting state (Tab)Ru(CO)(CH2CH2Ph) equals ethylene is calculated to be close to thermoneutral with ΔG = +0.8 kcal/mol.

β-Hydride Elimination. The formation of (Tab)Ru(CO)(H)(η2-styrene) by β-H elimination from the most stable conformation of (Tab)Ru(CO)(CH2CH2Ph) has a calculated free energy barrier of 4.2 kcal/mol and is exergonic by 2.9 kcal/mol (Scheme 9). Given the likely assumption that a β-agostic minimum is the immediate precursor to the β-hydride elimination pathway, the free energy barrier is reduced and the driving force is enhanced by the calculated 1.3 kcal/mol energy difference between the π-coordinated and β-agostic forms of [(Tab)Ru(CO)(CH2CH2Ph)]. The calculations are consistent with the experimental observation that β-hydride elimination is kinetically facile and reversible for TpRu(CO)(NCMe)(CH2CH2Ph) (2).

Isolation of the Transition State for C−H Activation. Detailed studies of stoichiometric C−H activations by transition metal centers have revealed two common reaction pathways. Oxidative addition reactions dominate the chemistry of late transition metal complexes in low oxidation states, whereas σ-bond metathesis transformations occur when d8 transition metal complexes initiate C−H activation.36,36, Recent reports of Ir(III) systems that activate C−H bonds have prompted speculation on the mechanism of the C−H bond-breaking step of late transition metal systems for which oxidative addition would yield an unusually high formal oxidation state.41,67-73 To assist efforts to discern the C−H activation pathway for catalysis using complex 1, DFT calculations have been employed for benzene C−H activation by the [(Tab)RuII(CO)(Me)] fragment. Oxidative addition to form the Ru(IV) complex [(Tab)RuIV(CO)(Me)(H)Ph)] (followed by reductive elimination of methane) and a non-oxidative addition reaction with simultaneous C−H bond-breaking and bond-forming have been considered (Scheme 10).

Multiple pathways were investigated for the C−H activation step including oxidative addition of a C−H bond to [(Tab)Ru(CO)(Me)] (A), oxidative addition of a methane C−H bond to [(Tab)Ru(CO)(Ph)] (B), and non-oxidative addition metathesis of a C−H bond of benzene with the Ru−CH3 bond of [(Tab)Ru(CO)Me] (C) (Scheme 11). Several different starting geometries were investigated for each of the proposed transition states. In all cases, the transition states collapsed to C (the non-oxidative addition transition state). The identity of the transition state was confirmed by calculation of the intrinsic reaction coordinate along the imaginary frequency, the primary motion of which corresponded to transfer of the transannular hydrogen from the methyl to the aryl carbon. The transition state geometry for C−H activation of benzene is depicted in Figure 6. The transannular hydrogen is only slightly closer to the ipso carbon (1.49 Å) than the methyl carbon (1.52 Å). The ruthenium carbon distances in the transition state (Ru−

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Figure 6. B3LYP/6-31G(d) calculated transition state for C–H activation of benzene by (Tab)Ru(Me)(CO) (atoms of the Tab ligand are shown in wire frame for clarity). View of calculated transition state illustrates the “out-of-plane” position of the hydrogen atom undergoing transfer.

Scheme 11. Pathways for C–H Activation Investigated by DFT Studies

C_{IPSO} = 2.23 Å; Ru–C_{Me} = 2.31 Å are only slightly longer (ca. 5–6%) than ground state bond lengths.

Although the calculations indicate that the pathway for benzene C–H activation does not involve an oxidative addition pathway (i.e., no Ru(IV) intermediate was found), the calculated distance between the Ru metal center and hydrogen (1.72 Å) in the C–H activation transition state is relatively short. Calculations by Oxgaard and Goddard reveal a Ru–H distance of 1.61 Å in the transition state for benzene C–H activation by TpRu(CO)(CH2CH2Ph).44 Closely related results have been reported from calculations to probe the mechanism of C–H activation for TpM(PPh3)2(Me) (M = Fe or Ru) complexes in which the authors suggested that the reaction pathway is intermediate between oxidative addition and η-bond metathesis processes.74 Computational studies of methane activation by Pt(H2O)Cl2 indicate a similar four-center transition state with a Pt–H distance of 1.99 Å.75 Calculations on the transition state of benzene C–H activation by (acac)2Ir(CH2CH2Ph) reveal an Ir–H distance of 1.58 Å, and Goddard et al. have differentiated this transition state (in which there is a metal–hydrogen interaction) from a η-bond metathesis transition state using the label oxidative hydrogen migration.73

Summary and Conclusions

TpRu(CO)(NCMe)(Ph) (1) catalyzes the hydroarylation of olefins through a pathway that likely involves olefin binding and insertion into the Ru–Ph bond followed by metal-mediated C–H activation. Reactions that incorporate α-olefins are mildly selective for production of linear over branched alkyl benzene products. Although the yields are poor, the selectivity for the hydrophenylation of propene using CpRu(PPh3)2(Ph) (5:1 linear-to-branched ratio) indicates that ligand architecture can be used to control the regioselectivity of olefin insertion. For reactions that incorporate ethylene, η-hydride eliminations to ultimately yield free styrene are not competitive with hydrophenylation, and this observation is attributed to reversible η-hydride elimination under catalytic conditions. The reversibility is likely a result of a relatively strong Ru–styrene bond due to the metal π-basicty with suppression of associative exchange due to the coordinative saturation of TpRu(CO)(H)2styrene. In contrast, following insertion of an α-olefin into the Ru–Ph bond, η-hydride elimination and olefin dissociation are competitive with catalytic hydrophenylation of the olefin. The similarities between catalytic olefin hydrophenylation using TpRu(CO)(Ph)(NCMe) or Ir(III) phenyl complexes reported by Periana et al. are noteworthy, and computational/experimental studies indicate the likelihood of closely related reaction pathways including the C–H activation step.40,41,73 However, evidence for η-hydride elimination and olefin dissociation has not been obtained for the Ir(III) systems, whereas such reactions have been observed with the Ru(II) systems reported herein. Similar to our experimental results with Ru(II), calculations by Oxgaard and Goddard suggest that η-hydride elimination for Ir(III) is facile and reversible.73

DFT studies suggest that the C–H activation steps of the catalytic cycles do not proceed through a Ru(IV) oxidative addition intermediate, and the Tp ligand may deter processes that lead to seven-coordinate Ru(IV) systems.51 As previously discussed,44,76 for C–H activations that proceed without an oxidative addition intermediate a distinction can be made between η-bond metathesis and electrophilic aromatic substitution. The calculated transition state for benzene C–H activation by (Tab)Ru(CO)(H)(Me)2 reveals that the hydrogen atom undergoing transfer to the methyl ligand is out of the aromatic plane (Figure 6). The hydrogen atom is calculated to be 0.87 Å removed from the “best” plane of the aromatic ring with a H–C–Ru angle of 50.3°, and the “out-of-phase” transannular hydrogen could be

(76) Bulls, A. R.; Schafer, W. P.; Serfas, M.; Bercaw, J. E.; Organometallics 1987, 6, 1219–1226.
indicative of a sp$^2$ to sp$^3$ rehybridization of the ipso carbon that would be consistent with an electrophilic substitution pathway. Alternatively, the out-of-plane hydrogen could reflect progress along the reaction coordinate and incipient Ru–C bond formation with the arene. The distinction between oxidative addition, σ-bond metathesis, oxidative hydrogen migration, and electrophilic aromatic substitution could have important implications for reactivity, and in order to probe this issue in more detail, we are currently studying the electronic influence of substituents on the arene.

**Experimental Section**

**General Methods.** All procedures were performed under inert atmosphere of dinitrogen in a Vacuum Atmospheres glovebox or using standard Schlenk techniques. The glovebox was purged with dinitrogen in a Vacuum Atmospheres glovebox or using standard Schlenk techniques. The glovebox was purged with an inert atmosphere of dinitrogen in a Vacuum Atmospheres glovebox or using standard Schlenk techniques. The glovebox was purged with an inert atmosphere of dinitrogen in a Vacuum Atmospheres glovebox or using standard Schlenk techniques. The glovebox was purged with an inert atmosphere of dinitrogen in a Vacuum Atmospheres glovebox or using standard Schlenk techniques. The glovebox was purged with an inert atmosphere of dinitrogen in a Vacuum Atmospheres glovebox or using standard Schlenk techniques.

Gas chromatography was performed on a Hewlett-Packard 5890 GC using either a J&W SE-30 or an HP-5 capillary column (30 m x 0.25 mm) with 0.25 μm film thickness and an FID detector. Chromatograms were produced using either a Hewlett-Packard 3396A integrator or a Perkin-Elmer TotalChrom 6.2 software. GC–MS was performed using a HP GC/MS system with a 30 m x 0.25 mm HP-5 column with 0.25 μm film thickness. Lecture bottles of ethylene (99.5%) and propylene (99.0%), 1-hexene, and decane were received. Ethylene (99.5%) was also received in a gas cylinder from MWSC High-Purity Gases and used as received. Ethylene (99.5%) was also received in a gas cylinder from MWSC High-Purity Gases and used as received. Ethylene (99.5%) was also received in a gas cylinder from MWSC High-Purity Gases and used as received. Ethylene (99.5%) was also received in a gas cylinder from MWSC High-Purity Gases and used as received. Ethylene (99.5%) was also received in a gas cylinder from MWSC High-Purity Gases and used as received.

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Ru(II)-Catalyzed Addition of C–H Bonds to Olefins

(1H, t, 8 Hz, phenyl meta or para), 3.74 (2H, dt, 7 Hz, phenyl ortho), 3.67 (2H, d, 7 Hz, phenyl ortho). Thus, the reaction solution was removed under a purge of dinitrogen, and the tube was placed in an oil bath and heated to 70 °C for 48 h. The reaction vessel was removed from the oil bath, cooled to room temperature, and purged into an inert atmosphere glovebox. The volatiles were removed under reduced pressure, leaving a dark amber residue. Washing of the residue with hexanes resulted in the formation of solid that was collected via vacuum filtration. Analysis of the solid by 1H NMR spectroscopy revealed clean formation of CpRu(PPh3)2(Me) (Ph) (2).

Conversion of CpRu(PPh3)2(Me) to CpRu(PPh3)2(Ph). A thick-walled screw-cap pressure tube was charged with 0.204 g of CpRu(PPh3)2(Me) (0.280 mmol) and 15 mL of benzene. The tube was placed in an oil bath and heated to 70 °C for 48 h. The reaction vessel was removed from the oil bath, cooled to room temperature, and purged into an inert atmosphere glovebox. The volatiles were removed under reduced pressure, leaving a dark amber residue. Washing of the residue with hexanes resulted in the formation of solid that was collected via vacuum filtration. Analysis of the solid by 1H NMR spectroscopy revealed clean formation of CpRu(PPh3)2(Ph).

X-ray Diffraction Study of TpRu(CO)(NCMe)(CH2CH2Ph) (2). Complex 2 was recrystallized by slow evaporation of a diethyl ether solution. A pale yellow crystal of TpRu(CO)(NCMe)(CH2CH2Ph) was collected in the perfluoropolyether PFO-XR75 (Lancaster) and sealed under nitrogen in a glass capillary. The crystal was optically aligned on the four-circle routine available in SAINT. The data were corrected for Lorentz and polarization effects as well as any crystal decay. Data collection consisted of the measurement of a total of 1650 frames in five different runs covering a hemisphere of data. The program SMART (version 5.6) was used for diffractometer control, frame scans, indexing, orientation matrix calculations, least-squares refinement, and data collection. The resulting data were reduced to produce a total of 1568 reflections and their intensities and estimated standard deviations. An absorption correction was applied using the SADABS routine available in SAINT. The data were corrected for Lorentz and polarization effects as well as any crystal decay. Two different data sets were collected. Data preparation was carried out by using the program XPREP, which gave 5140 unique reflections (Rint = 0.0455) with indices −15 ≤ h ≤ 15, −15 ≤ k ≤ 15, −20 ≤ l ≤ 19. The monoclinic space group was determined to be P2₁/c (No. 14).

The structure was solved by a combination of direct methods and Fourier methods with the use of the SHELXTL 6.1. Idealized positions for the hydrogen atoms were included as fixed contributions using a riding model with isotropic temperature factors set at 1.2 (aromatic protons) or 1.5 (methyl protons) times that of the adjacent carbon. The position of the B-H proton was refined with a fixed isotropic temperature factor. All hydrogen atoms were subsequently included in the final difference Fourier map.

A representative catalytic reaction is described. TpRu(CO)(NCMe)(Ph) (1) (0.021 g, 0.046 mmol) was dissolved in 4.1 mL (0.0456 mol) of benzene. To the homogeneous solution was added decane (0.269 mL, 1.38 mmol) as an internal standard. The solution was placed in a thick-walled glass reaction vessel and charged with 25 psi of ethylene pressure. The tube was then placed in an oil bath heated to 90 °C. Periodically the tube was removed from the oil bath and plunged into an ice bath. A 0.1 mL aliquot of the reaction solution was removed under a purge of dinitrogen, and the tube was quickly returned to the oil bath and ethylene pressure was restored. Samples (~1 μL) were removed from the aliquot and analyzed by GC-FID. With application of the appropriate correction factor (determined from regression plots of at least three sets of standard samples), the peak areas of the aliquot and analyzed by GC-FID. With application of the appropriate correction factor (determined from regression plots of at least three sets of standard samples), the peak areas of the aliquot and the tube were used with the internal standard to calculate product yields.

Determination of Kinetic Isotope Effect. TpRu(CO)(NCMe)(Ph) (1) (0.0511 g, 0.111 mmol) was dissolved in an equimolar solution of C6H6 (0.49 mL, 5.5 mmol) and CpRu(Ph)2 (0.53 mL, 5.5 mmol). The solution was placed in a glass thick-walled pressure tube and placed under 25 psi ethylene pressure. The tube was then placed in an oil bath heated to 90 °C. Periodically the tube was removed from the oil bath and plunged into an ice bath. A 0.1 mL aliquot of the reaction solution was removed under a purge of dinitrogen, and the tube was quickly returned to the oil bath and ethylene pressure was restored. Samples (~1 μL) were removed from the aliquot and analyzed by GC-FID. With application of the appropriate correction factor (determined from regression plots of at least three sets of standard samples), the peak areas of the aliquot and analyzed by GC-FID. With application of the appropriate correction factor (determined from regression plots of at least three sets of standard samples), the peak areas of the aliquot and the tube were used with the internal standard to calculate product yields.

Kinetic Study of Catalytic Reactions. A sample determination of the rate of the formation of ethylbenzene catalyzed by TpRu(CO)(NCMe)(Ph) (1) at 70 °C is given. A stock solution consisting of 0.069 g (0.149 mmol) of TpRu(CO)(NCMe)(Ph) (1) in 13.33 mL (0.149 mol) of benzene was made. Decane (0.087 mL, 0.45 mmol) was added as an internal standard. Stock solution (2.0 mL) was added to a thick-walled glass reaction vessel and charged with 25 psi of ethylene pressure and heated to 70 °C in an oil bath. Periodically, the tube was removed from the oil bath and plunged into an ice bath. A 0.1 mL aliquot of the reaction solution was removed under a purge of dinitrogen, and the tube was quickly returned to the oil bath and ethylene pressure was restored. The aliquot of the solution was removed under a purge of dinitrogen, and the tube was quickly returned to the oil bath and ethylene pressure was restored. Samples (~1 μL) were removed from the aliquot and analyzed by GC-FID. With application of the appropriate correction factor (determined from regression plots of at least three sets of standard samples), the peak areas of the aliquot and analyzed by GC-FID. With application of the appropriate correction factor (determined from regression plots of at least three sets of standard samples), the peak areas of the aliquot and the tube were used with the internal standard to calculate product yields.

Reaction of TpRu(CO)(NCMe)(Ph) with Ethylene in Acetonitrile. An acetonitrile solution (~35 mL) of TpRu(CO)(NCMe)(Ph) (1) (0.118 g, 0.256 mmol) was placed under 250 psi of ethylene. The solution was heated to 90 °C for a period of 6 h. The volatiles were removed under reduced pressure, leaving a yellow residue. A portion of the residue was dissolved in C6D6 and a 1H NMR spectrum was acquired. The spectrum revealed two diamagnetic TpRu complexes including approximately 50% of TpRu(CO)(NCMe)(CH2CH2Ph) (2).

The second product is observed only at high ethylene pressures and remains uncharacterized.
lated from values from the International Tables for X-ray Crystallography.\textsuperscript{79}

**Computational Methods.** Quantum calculations were carried out using the Gaussian 98 package.\textsuperscript{80} The B3LYP hybrid functional was employed for all calculations.\textsuperscript{81} Heavy atoms were described with the Stevens relativistic effective core potentials (ECPs) and valence basis sets (VBSs).\textsuperscript{82,83} The valence basis sets of main group elements were augmented with a d polarization function. This ECP/VBS combination, termed SBK(d), has been validated for the calculation of a wide variety of transition metal properties in previous studies.\textsuperscript{84,85}

As a model of the full (tris-pyrazolyl)borate (Tp) ligand, the tris(azo)borate (Tab) ligand, $[\text{HB}(\text{N}_3\text{NH})_3]^{-}$, was used. In previous research, Tab was shown to faithfully reproduce the structure and energetics of the full Tp models for C–H activation potential energy surfaces.\textsuperscript{61}

All stationary points were fully optimized without symmetry constraint. Several conformations of the different ligands were investigated by torsion about the appropriate metal–ligand bonds; the lowest energy conformers found are used in the analyses. The energy Hessian was calculated at all stationary points to characterize them as minima (no imaginary frequencies) or transition states (one and only one imaginary frequency). The quoted energies include zero-point, enthalpy, and entropic corrections determined from unscaled vibrational frequencies calculated at the B3LYP/SBK(d) level of theory. All energetic determinations were done at 298.15 K and 1 atm. Geometries of all isolated stationary points are given as Supporting Information.

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**Supporting Information Available:** Cartesian coordinates for all optimized minima are given. Complete tables of crystal data, collection and refinement data, atomic coordinates, bond distances and angles, and anisotropic displacement coefficients for TpRu(CO)(NCMe)(CH$_2$CH$_2$Ph) ($\text{2}$). $^1$H NMR spectra of complexes $\text{4}$ and $\text{5}$ and a $^{13}$C NMR spectrum of complex $\text{5}$. This material is available free of charge via the Internet at http://pubs.acs.org.

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