Plasma Aftertreatment for Simultaneous Control of $\text{No}_x$ and Particulates

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This article was submitted to 1999 Diesel Engine Emissions Reduction Workshop, Castine, ME, July 5-8, 1999

October 28, 1999
PLASMA AFTERTREATMENT FOR SIMULTANEOUS CONTROL OF NO\textsubscript{x} AND PARTICULATES

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Abstract

Plasma reactors can be operated as a particulate trap or as a NO\textsubscript{x} converter. The soluble organic fraction (SOF) of the trapped particulates can be utilized for the oxidation of NO to NO\textsubscript{2}. The NO\textsubscript{2} can then be used to non-thermally oxidize the carbon fraction of the particulates. This paper examines the energy density required for oxidation of the SOF hydrocarbons and the fate of NO\textsubscript{2} during the oxidation of the particulate carbon. The energy density required for complete oxidation of the SOF hydrocarbons is shown to be unacceptably large. The reaction of NO\textsubscript{2} with carbon is shown to lead mainly to backconversion of NO\textsubscript{2} to NO. These results suggest that the use of a catalyst in combination with the plasma will be required to efficiently reduce the NO\textsubscript{x} and oxidize the SOF hydrocarbons.

Introduction

Plasma reactors can be operated as a particulate trap [ref. 1-3] or as a NO\textsubscript{x} converter [ref. 4-7].

Particulate trapping in a plasma reactor can be accomplished by electrostatic precipitation. Corona-type reactors operating as electrostatic precipitators for diesel soot are reported in refs. [1-2]. A ferroelectric pellet bed reactor, also operating as an electrostatic precipitator of diesel soot, is reported in ref. [3]. Plasma-based traps need periodic regeneration just like any other particulate trap. Regeneration by the plasma can be achieved in a thermal mode or a non-thermal mode.

The plasma reactor can be operated occasionally in the arc discharge mode to thermally oxidize the particulates. In ref. [2], this mode is referred to as the self-cleaning phase of the electrostatic muffler. Thermal oxidation of the particulates precipitated on the electrode surface is achieved in very localized regions of high temperature near the tip of the arc microdischarges.

The plasma reactor can also be operated in a non-thermal corona discharge mode to provide continuous oxidation of the particulates. The non-thermal oxidation is presumably accomplished by the O and OH radicals resulting from electron-impact dissociation of oxygen and water vapor molecules, or by negative hydronium cluster ions, [(H\textsubscript{2}O\textsubscript{m}O\textsuperscript{–})\textsubscript{n}], resulting from electron attachment. Ref. [1] observed a self-cleaning effect in the corona mode that correlates with the production of CO and CO\textsubscript{2}. Ref. [8] reports on the oxidation of soot in a dielectric-barrier discharge. It was conjectured that OH radicals oxidized the soot; the soot reduction did not occur in the presence of large amounts of CO, presumably because the OH radicals were consumed by CO.
Diesel particulates are composed mainly of the carbon fraction and the soluble organic fraction (SOF). The SOF could possibly be utilized for the oxidation of NO to NO$_2$ in a plasma. As a NO$_x$ converter, the plasma requires hydrocarbons to achieve high oxidation efficiency with low electrical energy consumption. The NO$_2$ can then be used to oxidize the carbon fraction, similar to that in a Continuously Regenerated Trap (CRT) [ref. 12]. The use of NO$_2$ for the oxidation of trapped particulates in CRT devices is fairly well established. In CRT, a precious metal catalyst is used to oxidize NO to NO$_2$ upstream of a particulate filter. The CRT method requires low sulfur fuel because the catalytic oxidation of NO to NO$_2$ also leads to the oxidation of SO$_2$ to SO$_3$. The use of a plasma for the oxidation of NO to NO$_2$ can make the process more tolerant to the sulfur content of the fuel. The fate of NO$_2$ during the oxidation of carbon is important if one contemplates the use of plasma aftertreatment for the simultaneous removal of NO$_x$ and particulates.

This paper examines the energy density required for oxidation of the SOF hydrocarbons and the fate of NO$_2$ during the oxidation of the particulate carbon.

Results

Both chemical kinetics calculations and experimental measurements are presented in this section. We used propene as the hydrocarbon surrogate in the calculations because the chemical reaction database for propene is more established and facilitates comparison of the modeling to experiments. The plasma reactor used in the experiments is a pulsed corona discharge reactor consisting of a metal wire inside a metal cylinder. The plasma chemistry is not peculiar to this type of plasma processor; all electrical discharge plasma reactors accomplish essentially the same gas-phase plasma chemistry under the same gas conditions [ref. 9-10]. The important control parameter in the plasma reactor is the electrical energy density delivered to the plasma [ref. 9-11].

Figure 1 shows the calculated concentration of aldehydes formed during plasma processing of 100 ppm NO in a simulated diesel exhaust at 200°C with a C$_1$/NO$_x$ of 6. Formaldehyde is the major product of the partial oxidation of propene in the plasma. Note that the electrical energy density required to convert NO to NO$_2$ is much less than that required to fully oxidize the aldehydes. Whereas it is possible to convert 100 ppm of NO to NO$_2$ with less than 10 J/L, conversion of the resulting 50 ppm of formaldehyde to H$_2$O and CO$_2$ requires more than 150 J/L. This result shows that the electrical power required by the plasma to completely oxidize the hydrocarbons is unacceptably large.

Some of the minor products shown in Figure 1 include CH$_3$CHCO and C$_2$H$_5$CHO. Species CH$_3$CHCO is methyl ketene and species C$_2$H$_5$CHO is acrolein. Acrolein is experimentally observed as an intermediate product in the low temperature oxidation of propene [ref. 13]. The addition of O atom to propene to form a biradical and its decomposition forms methyl ketene. The rate of this process is based on the current understanding of the O + propene reaction as reviewed in Ref. [14]. Reactions to consume acrolein and methyl ketene by radical attack are included in the detailed chemical kinetic mechanism.

Figure 2 shows the calculated concentration of other species formed during plasma processing of 100 ppm NO in a simulated diesel exhaust at 200°C with a C$_1$/NO$_x$ of 6. Large amounts of CO are formed during the plasma oxidation of hydrocarbons.
Figure 1. Chemical kinetics calculation of the concentration of aldehydes formed during plasma processing of 100 ppm NO in a simulated diesel exhaust with 10% O₂, 10% CO₂, 5% H₂O, balance N₂. Propene additive, C₁/NOₓ = 6. Gas temperature = 200°C.

Figure 2. Chemical kinetics calculation of the concentration of various species during plasma processing of 100 ppm NO in a simulated diesel exhaust with 10% O₂, 10% CO₂, 5% H₂O, balance N₂. Propene additive, C₁/NOₓ = 6. Gas temperature = 200°C.
Figure 3 shows the comparison between the model predictions and experimental measurements of the concentration of formaldehyde and CO formed during plasma processing of 500 ppm NO in a gas mixture with 10% O₂, balance N₂, at 100°C with a C₁/NOₓ of 6. This level of NO is typical of that in heavy-duty diesel engine exhaust. There is fairly good agreement between the modeling and the experiment. With a C₁/NOₓ of 6, about 25 J/L is required to get maximum oxidation of NO to NO₂. At this energy density, about 150 ppm of formaldehyde has already been formed. The amount of aldehydes increases further as the energy density is increased. Oxidation of these aldehydes will require very large electrical energy density input to the plasma.

The oxidation of the carbon by NO₂ can lead to reduction of NOₓ or backconversion of NO₂ to NO. The following experiments measure the NOₓ reduction efficiency and the amount of backconversion of NO₂ to NO.

Figure 4 shows the fate of NO₂ during its reaction with carbon pellets at 250°C in a gas stream containing only N₂. There is about 20% NOₓ reduction. The NOₓ remaining after the reaction is composed mostly of NO.

Figure 5 shows the same type of experiment conducted in a gas stream containing 10% O₂, balance N₂. The oxygen seems to have promoted the NOₓ reduction efficiency to about 30%. Again the NOₓ remaining after the reaction is composed mostly of NO.
Figure 4. Oxidation of carbon by NO₂ in N₂. Gas temperature = 250°C.

Figure 5. Oxidation of carbon by NO₂ in 10% O₂, balance N₂. Gas temperature = 250°C.

Figure 6 shows the result of another experiment to simulate the effect of the volatile organic fraction of particulates on the NO₂ reduction. The gas stream in this experiment contains 10% O₂ and 1500 ppm C₁ kerosene. The NOₓ reduction efficiency is about 35%, with the remaining NOₓ being composed mostly of NO.
The results in Figures 4-6 show that even though the NO\textsubscript{2} could be utilized for the oxidation of the carbon fraction of the particulates, the reaction with the carbon fraction cannot provide a high level of NO\textsubscript{x} reduction. Another means will have to be provided to achieve a high NO\textsubscript{x} reduction efficiency. If one is going to use plasma-assisted SCR for the NO\textsubscript{x} reduction, the regeneration of the particulate trap will have to be done without the NO\textsubscript{2}. A way to do this would be to operate the plasma reactor occasionally in the arc mode to thermally oxidize the carbon fraction of the particulates. Ref. [2] describes an example of the effectiveness of this technique and how it can be implemented.

Conclusions

The electrical energy density required for complete oxidation of hydrocarbons in a plasma is much greater than that required to achieve maximum NO\textsubscript{x} conversion. Aldehydes and CO will be formed. The energy density required to oxidize the aldehydes will be unacceptably large even for light-duty applications. A catalyst in combination with the plasma will be required to take care of the aldehydes and CO.

The NO\textsubscript{2} from the plasma can be used to non-thermally oxidize the carbon fraction of trapped particulates. However, the reaction of NO\textsubscript{2} with carbon cannot provide a high level of NO\textsubscript{x} reduction and leads mostly to the backconversion of NO\textsubscript{2} to NO. If one is going to use plasma-assisted SCR for the NO\textsubscript{x} reduction, the regeneration of the particulate trap will have to be done without the NO\textsubscript{2}. A way to do this would be to operate the plasma reactor occasionally in the arc mode to thermally oxidize the carbon fraction of the particulates.

Acknowledgments

This work was performed at Lawrence Livermore National Laboratory under the auspices of the U.S. Department of Energy under Contract Number W-7405-ENG-48, with support from the Chemical Sciences Division of the DOE Office of Basic Energy Sciences, the DOE Office of Fossil Energy, the
DOE Office of Transportation Technologies, the Strategic Environmental Research and Development Program, and a Cooperative Research and Development Agreement with Cummins Engine Company.

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