

PLATINUM RESISTANCE THERMOMETRY AT HIGH TEMPERATURES.

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I. INTRODUCTION.

The present investigation was undertaken to determine the advantages and limitations of platinum resistance thermometers at high temperatures, especially with reference to reproducibility of scale, method of calibration, formulæ best expressing the relation between resistance and temperature, effect of impurities in the platinum wire, changes due to high temperatures, and the most satisfactory method of construction and use. The reproducibility of high temperatures by the melting or freezing points of metals obtained from different sources and furnished as chemically pure was also studied.

We realize that a very considerable amount of most valuable work has been done by several investigators in the domain of platinum thermometry of precision at high temperatures, yet it seemed to us that with the facilities at our disposal there was opportunity to study some of the questions above referred to with still greater exactness, especially in view of the publication of several recent researches on the gas thermometer at high temperatures, which afford, through the temperatures of the freezing points of pure metals, an indirect but very certain comparison of the gas and platinum scales.

The platinum resistance thermometer was first seriously proposed and used for the measurement of high temperatures by Siemens,* who in 1871 submitted several such thermometers to the British association for test. The results of these tests⁸ showed large and progressive changes in resistance after each exposure to high temperatures. The further investigation of the platinum thermometer by Callendar,¹² fifteen years later, showed that if the thermometer was so constructed that the platinum wire was protected from contamination by the supporting frame, the changes in resistance after some hours' exposure to high temperatures (red heat), were very small (a few parts in 10,000).

Callendar¹² has shown that if we define the *platinum temperature*, pt , by the relation

* References will be found in the Bibliography, Sec. XI, p. 223.

$$pt = 100 \frac{R - R_0}{R_{100} - R_0} \quad (1)$$

where R = observed resistance at the temperature, t° ,

$$\begin{aligned} R_0 &= \text{“ “ “ } 0^\circ \text{ C.,} \\ R_{100} &= \text{“ “ “ } 100^\circ \text{ C.,} \end{aligned}$$

the relation between the platinum temperature, pt , and the temperature, t , on the scale of the gas thermometer is represented by the relation

$$t - pt = \delta \left\{ \frac{t}{100} - 1 \right\} \frac{t}{100} \quad (2)$$

where δ is a constant for any given sample of platinum, having a value of about 1.50 for pure platinum and a higher value for impure platinum.

The *fundamental interval*, FI , of a thermometer, is defined by the relation

$$FI = R_{100} - R_0 \quad (3)$$

and the *fundamental coefficient*, c , by the relation

$$c = FI / 100 R_0 \quad (4)$$

Three temperatures only are necessary for the standardization of a platinum thermometer. Two of these temperatures generally chosen are the melting point of ice and the temperature of saturated steam at known pressure. The third point recommended by Callendar and Griffiths,¹⁵ and generally used on account of the numerous careful determinations that have been made of it, is the boiling point of sulphur, observed under carefully specified conditions.

The work of numerous investigators has established the validity of the above relation (Eq. 2) and has shown that the platinum thermometer, standardized by observations in ice, steam, and sulphur vapor, will serve to reproduce temperatures on the gas scale within the range -80° to $+1100^\circ$ C almost, if not quite, within the limits of accuracy of the gas thermometer. The scale defined by the constant δ , determined by such a calibration, begins to depart appreciably from the scale of the Hydrogen or

Helium thermometer at temperatures below -100° , and for the measurement of temperatures in the interval 0° to -200° , it is best to choose as calibration temperatures melting ice, solid CO_2 and ether, and the boiling point of liquid oxygen. If still lower temperatures are to be measured, it is necessary to use the temperatures of boiling or solidification of hydrogen, and if the highest attainable accuracy is desired it will probably be necessary to use more than three calibration temperatures, and an equation of a higher degree connecting the resistance and the temperature. For the calibration of calorimetric resistance thermometers intended for the measurement of small temperature changes, the transition point of sodium sulphate ($32^{\circ}384$) may be used with advantage as the third calibration temperature.¹¹⁷

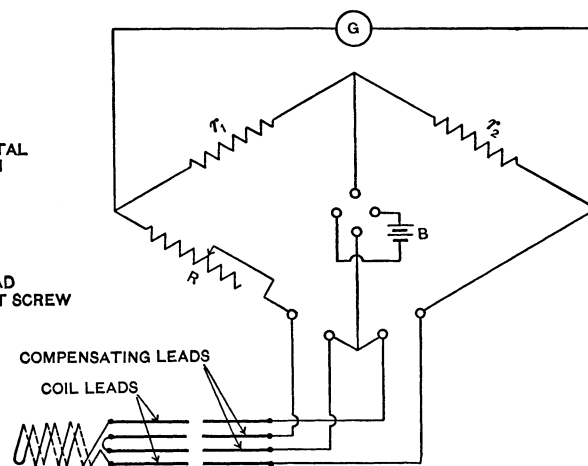
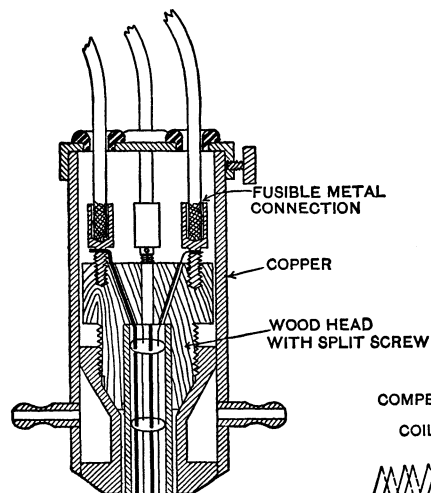
II. THERMOMETERS AND METHODS OF USE

In the present investigation 11 resistance thermometers were used, made of wire of varying degrees of purity and of various sizes from 0.1 to 0.6 mm diameter, and with resistances, R_0 , ranging from 0.11 ohm to 21.5 ohms. Two of these thermometers were made of pure palladium wire 0.3 mm diameter, while the other 9 were of platinum.

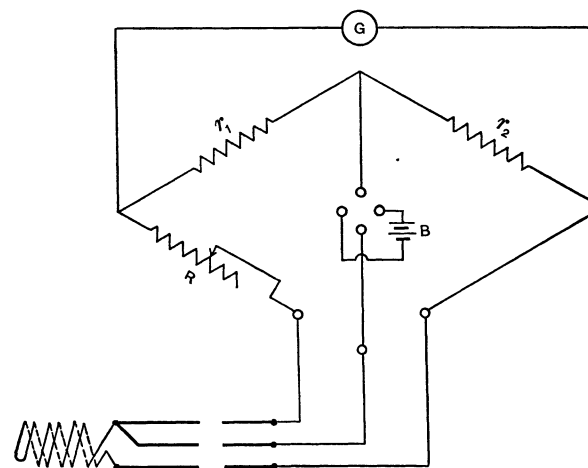
Three types of thermometer were used, which may be briefly designated as

- a. Compensated lead type.
- b. Potential terminal type.
- c. Combined compensated lead and potential terminal type.

These three types are illustrated diagrammatically in Plates I and II. In the compensated lead type, due to Callendar,¹⁸ two fairly large platinum leads (preferably over 0.6 mm diameter), as nearly identical as possible with the leads to the coil, run down the stem of the thermometer, and are connected at their lower ends. When these leads are put into the arm of the Wheatstone bridge adjacent to the arm in which the thermometer coil is inserted the thermometer is compensated for variations in temperature of the leads to the coil, due to variable depth of immersion. To compensate for the effects of heat conduction from the coil along the heavier platinum leads, the ends of the compensating leads are



4 LEAD COMPENSATED THERMOMETER

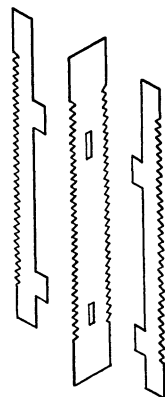


3 LEAD COMPENSATED THERMOMETER

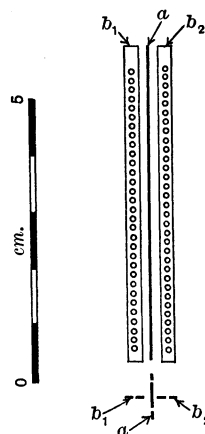


PORCELAIN
TUBE

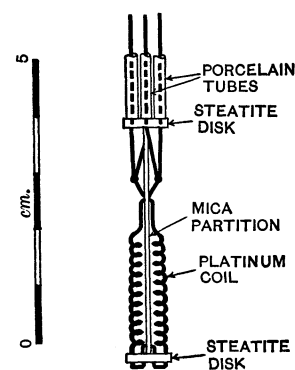
MICA DISKS



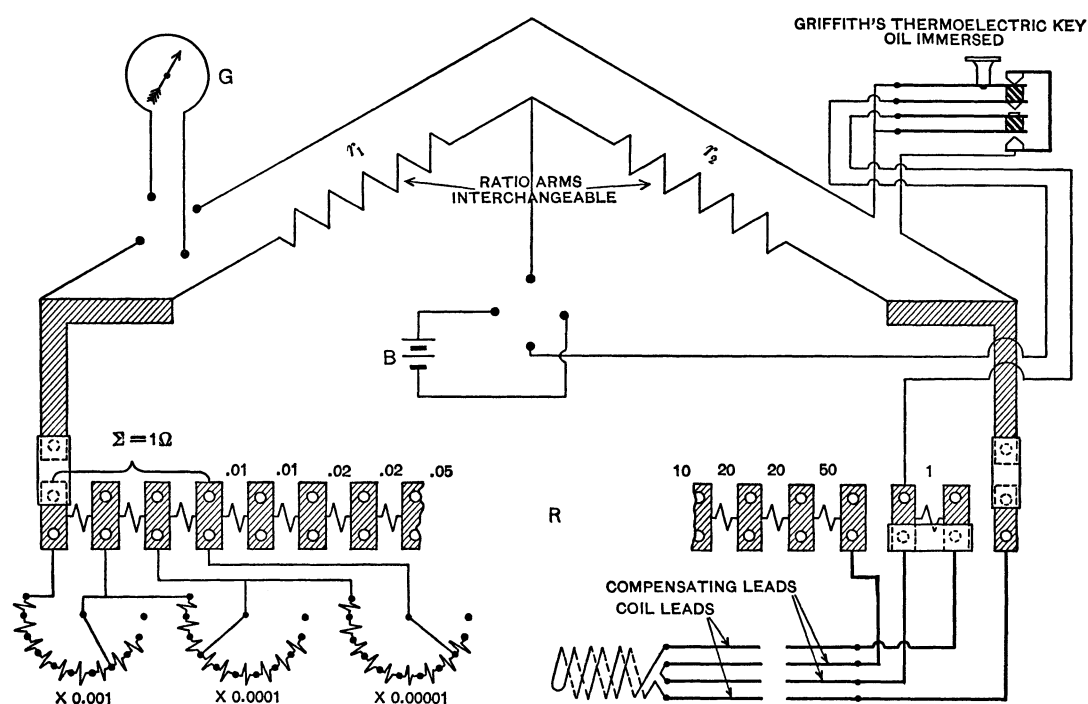
SERRATED MICA FRAME



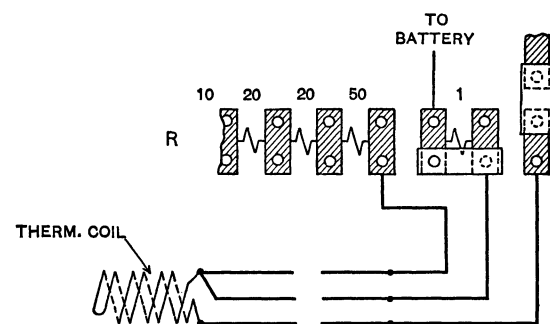
MICA FRAME FOR
THERM. NO.1787F & 1787G



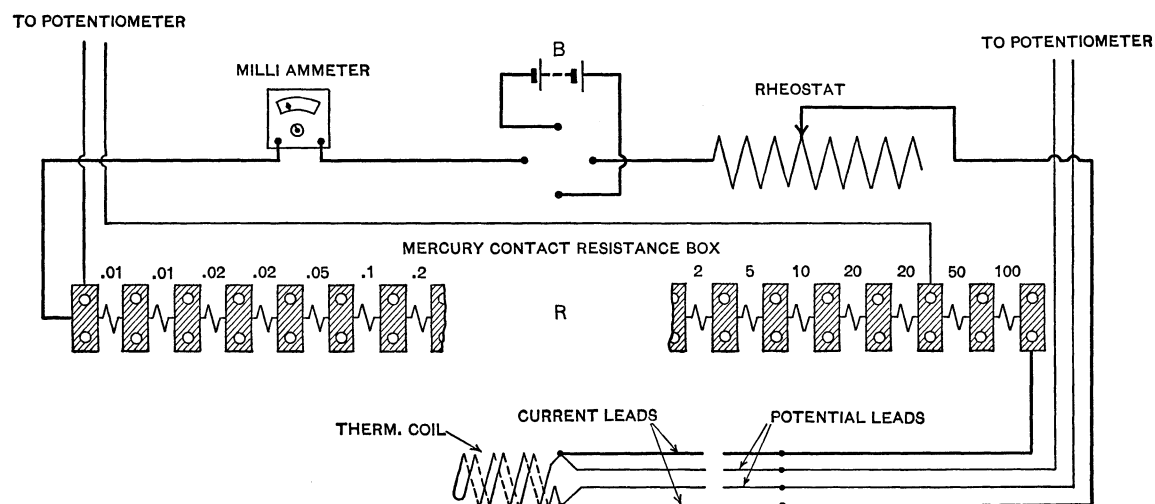
THERMS.NO.
9837 & 9838



DIAGRAMMATIC SKETCH OF BRIDGE AND 4 LEAD COMPENSATED THERMOMETER.



CONNECTIONS FOR 3 LEAD COMPENSATED THERMOMETER.



CONNECTIONS FOR POTENTIAL TERMINAL THERMOMETER.

joined by a short length of fine platinum wire of the same size as the wire of which the coil is made. In thermometers of the potential terminal type two leads are fused to each end of the coil. In this type of thermometer both the current and potential leads may obviously be made of much smaller wire than for the compensating lead type. The method of using this thermometer, which consists in comparing the potential difference at the terminals of the coil with the potential difference at the terminals of a known resistance carrying the same current, will be understood from the diagram. In thermometers of the combined compensated lead and potential terminal type the current leads should be made of fairly large wire to secure better compensation. If one of the potential leads is omitted we have a three-lead compensated thermometer of the Siemens type.⁵

Methods of Measurement.—The compensated lead thermometers (type *a*, and type *c* when so used) were used (see Plate II) with a specially designed Wheatstone bridge,²² having mercury contact plugs and manganin resistance coils ranging from 0.01 to 50 ohms. The coils were immersed in a motor-stirred and thermostatically controlled oil bath. The finer steps were obtained on three shunt dials giving steps of 0.001, 0.0001, and 0.00001 ohm, respectively. The 1000-ohm ratio coils were used throughout this work. The ratio coils could be quickly reversed by interchanging two mercury contact links, thus eliminating the effect of inequality of the ratio coils. The bridge was used with a Griffiths' oil immersed key,²⁴ which kept the galvanometer circuit always closed except at the instant of making the battery circuit, thus eliminating the effects of small thermoelectric currents in the circuits. As a further precaution measurements were made with the battery and galvanometer circuits alternately reversed, as well as with the ratio arms interchanged, in the following sequence:

r_1/r_2		r_2/r_1
1. Battery +, Galv. +		3. Battery -, Galv. +
2. " - " +		4. " + " +
7. " - " -		5. " + " -
8. " + " -		6. " - " -

The connections are shown diagrammatically in Plate II.

This bridge has been calibrated several times a year since its construction in 1903 and at intervals of about eight weeks during the progress of this investigation. The small changes, amounting to a few parts in 100,000, are apparently seasonal and are undoubtedly due to the effects of moisture.

The method of using the potential terminal thermometers is shown in Plate II. The potential difference at the terminals of the thermometer coil was compared with the potential difference at the terminals of the resistance R by means of a special Leeds and Northrup type K potentiometer. In some of the later work a Diesselhorst potentiometer, constructed by Otto Wolff, of Berlin, was also used. The resistance R , which was taken from a mercury contact resistance box of manganin coils in a stirred and thermostatically controlled oil bath, could be adjusted, to the nearest 0.01 ohm, to equality with the resistance of the thermometer coil, so that nearly equal potential differences were compared.

The low resistance thermometers ($R_0 = 0.11$ ohm) were used as potential terminal instruments. In some of the earlier tests these thermometers were calibrated in connection with a special slide-wire double Kelvin bridge submitted with the thermometers by the makers, Messrs. Leeds and Northrup, who have applied this method of resistance thermometry to industrial work.

The order of accuracy of the resistance measurements throughout the work was 1 or 2 parts in 100,000.

Construction of Thermometers.—The thermometers, with the exception of Nos. 479 and 480, from the Cambridge Scientific Instrument Company, and the two low-resistance thermometers ($R_0 = 0.11$ ohm) 9837 and 9838, from Leeds and Northrup, were very kindly constructed for this investigation by Mr. H. C. Dickinson and Mr. J. J. Crowe of this Bureau. The essential details of construction are shown in Plate I. The coil is wound on a notched mica frame. The several leads are insulated from one another by mica washers. The platinum leads terminate in the head of the thermometer in small copper cups containing a fusible alloy by means of which connection is made to stranded leads (carefully adjusted to equality for the compensated lead thermometers)

which serve to connect the thermometer to the measuring apparatus. The head of the thermometer is made of wood thoroughly dried and boiled in paraffine. This head terminates at the bottom in a split screw which serves to grip the porcelain stem when the surrounding double walled copper head is screwed on to the wooden head. The thick metal head insures a quite uniform temperature distribution throughout the inclosure and thus reduces to a minimum the thermal electromotive forces within the head of the thermometer. Undue heating of the head of the thermometer at high temperatures, with consequent increase in these emf's., can be avoided by passing a current of air through the space between the walls.

The containing tubes were of Berlin porcelain glazed on the outside only, except for thermometers 9837 and 9838, which were inclosed in quartz.

A modified form of mica frame was used in the construction of thermometers 1787 F and 1787 G, with a view to diminishing the possible effects of strains in the platinum wire of the coil due to thickening of the mica after repeated exposures to high temperatures. The construction of the mica frame for these thermometers will be understood from the illustration. The wire was threaded through holes in the mica, and the two independent mica strips, b_1 , b_2 , not being in contact with the mica strip a , gave a flexible mounting for the wire.

The wire of which the thermometer coils were constructed was of varying degrees of purity, with values of δ ranging from 1.50 to 1.80. In Table I will be found a summary of the constants of the thermometers.

Heating by Measuring Current.—The measuring current through the coils of the thermometers was from 0.004 to 0.010 ampere for the different thermometers, except for the low-resistance thermometers. For any given thermometer the same measuring current was always used. Measurements made with other currents showed that the resulting values of pt were practically identical with the pt corresponding to infinitely small current. Two different currents were used with the low resistance thermometers, viz., 0.1 and 0.5 ampere. While with the larger cur-

rent the excess of temperature of the coil above its surroundings was more than 1° in the observed value of R_0 , the value of FI found was practically identical for both these currents.

TABLE 1.
Range of Constants of Platinum Thermometers.

No. of Thermometer	R_0	FI	c	δ	Method of Measurement	Meas. current, amp.	Diam. of wire, mm
1787 A	21.3476 21.0617	4.4067 4.4203	.00206426 210065	1.571 1.569	Poten- tiometer.	.004	0.1
1787 C	3.48779 3.48174	1.34298 1.34114	.00385052 385192	1.504 1.504	Poten- tiometer.	.010	.15
1787 F	2.85355 2.85567	1.11228 1.11312	.00389787 389793	1.503 1.503	Bridge, 3 leads.	.006	.2
1787 F	2.84782 2.84987	1.11069 1.11119	.00390014 389909	1.503 1.503	Poten- tiometer.	.010	.2
479	4.27342 4.28823	1.64585 1.64513	.00385136 383638	1.516 1.508	Bridge, 4 leads.	.006	.2
478	5.15665 5.16566	1.99745 2.00000	.00387354 387172	1.507 1.507	Poten- tiometer.	.007	.2
480	2.62081	1.00503	.00383480	1.551	Bridge, 4 leads.	.006	.2
1787 E	2.92482 2.92226	0.50511 .50540	.00172630 172948	1.803 1.803	Poten- tiometer.	.005	.3
9837	0.11275 .11307	.04324 .04325	.0038350 38251	1.54 1.54	Poten- tiometer.	.50	.6
9838	.10951 .10954	.04192 .04184	.0038280 39196	1.56 1.56	Poten- tiometer.	.50	.6

Range of Constants of Palladium Thermometers.

1787 G	1.87254 1.86284	0.62761 .62696	0.00335163 33590	2.890	Bridge, 3 leads.	.006	.3
1787 G	1.87345 1.86802	.62809 .62890	.00335258 336667	2.890	Poten- tiometer.	.015	.3
Pd II	1.30897	.43591	.00333017	2.947	Poten- tiometer.	.01	.35

The effect of using different measuring currents (with thermometer 1787 C.) ranging from 0.0025 to 0.100 ampere, is shown in Table 2*. The current normally used with this thermometer

* See Appendix, p. 198.

throughout the investigation was 0.01 ampere. It will be seen that a measuring current almost, if not quite, five times as great might have been used without appreciably affecting the results.

For a given small excess in temperature of the platinum coil above the temperature of its surroundings, the energy radiated in steam is 3.7 and in sulphur 52 times that radiated at 0° C, assuming that the radiation from platinum is proportional to the fifth power of the absolute temperature. For constant measuring current the energy supplied to the coil at the S. B. P. is only 2.6 times the energy supplied at 0° C. Hence, since it is observed that the excess of temperature of the coil at the S. B. P. is only about 25 per cent less than the excess of temperature at 0° C, it follows that only a relatively small portion of the energy supplied to the coil by the measuring current is lost by radiation, by far the greater part of the loss being due to convection and conduction.

Calibration of Thermometers.—The thermometers were calibrated by the Callendar-Griffiths method by observations of the resistance in melting ice, steam, and the vapor of boiling sulphur.

The resistance in ice, R_0 , was determined in a finely divided mixture of pure ice saturated with distilled water. The resistance in steam was determined in the usual laboratory form of Regnault hypsometer. The temperature of the steam in this type of hypsometer was compared with the temperature of the steam in the International Bureau type of hypsometer, designed by Chappuis, by taking measurements with the same thermometer in the two hypsometers alternately. The observed temperature of the steam in the Regnault agreed with that in the Chappuis hypsometer to within 0°.01, being about 0°.003 higher for slow rate of boiling and about 0°.006 higher for fairly rapid rate of boiling. The observations in steam were accompanied by simultaneous observations of the atmospheric pressure with a Fuess standard barometer, accurate to 0.02 mm. The temperature of the steam was then obtained from the observed barometric height reduced to 0°, latitude 45°, sea level, and from Broch's steam tables.

The observations in sulphur, together with simultaneous barometer observations, were carried out with the apparatus and

in the manner fully described under the head of "The boiling point of sulphur," p. 184.

The S. B. P. was taken as $444^{\circ}.70$ at standard atmospheric pressure.

Each thermometer was recalibrated at frequent intervals throughout the progress of the work. Several consecutive freezing (or melting) point determinations were usually made, requiring several hours, during which time the temperature of the crucible of metal varied only a few degrees above and below the melting point. Before and after each group of such determinations the thermometer was recalibrated. For the observations at the higher temperatures, where the changes in the constants of the thermometer become appreciable, the resulting freezing points are given in Tables* 8-10, as deduced from both the calibration before and the calibration after the observations in the metal.

Constants of the Thermometers.—In Table 1 (p. 156) are indicated, for each thermometer, the values of its characteristic constants, R_0 , FI , c , and δ , as observed at the beginning of this investigation and at its close, and also the method of measurement used with each instrument. The changes in the constants due to use of the thermometers at high temperatures are discussed subsequently (p. 178).

III. MEASUREMENT AND REPRODUCIBILITY OF FREEZING AND MELTING POINTS.

The determinations with thermometers of pure platinum of the freezing and melting points of the several metals used in this investigation are given below. The results of measurements, by means of thermocouples calibrated at the temperatures of freezing of zinc, antimony, and copper, made by one of the authors,† over a period of several years, on the reproducibility of these fixed points, as given by metals from different makers, are also included. The thermoelectric determinations at the lower temperatures are less precise than those with resistance thermometers.

* All tables except Nos. 1, 11, 12, 15, and 25, are in the appendix.

† Burgess.

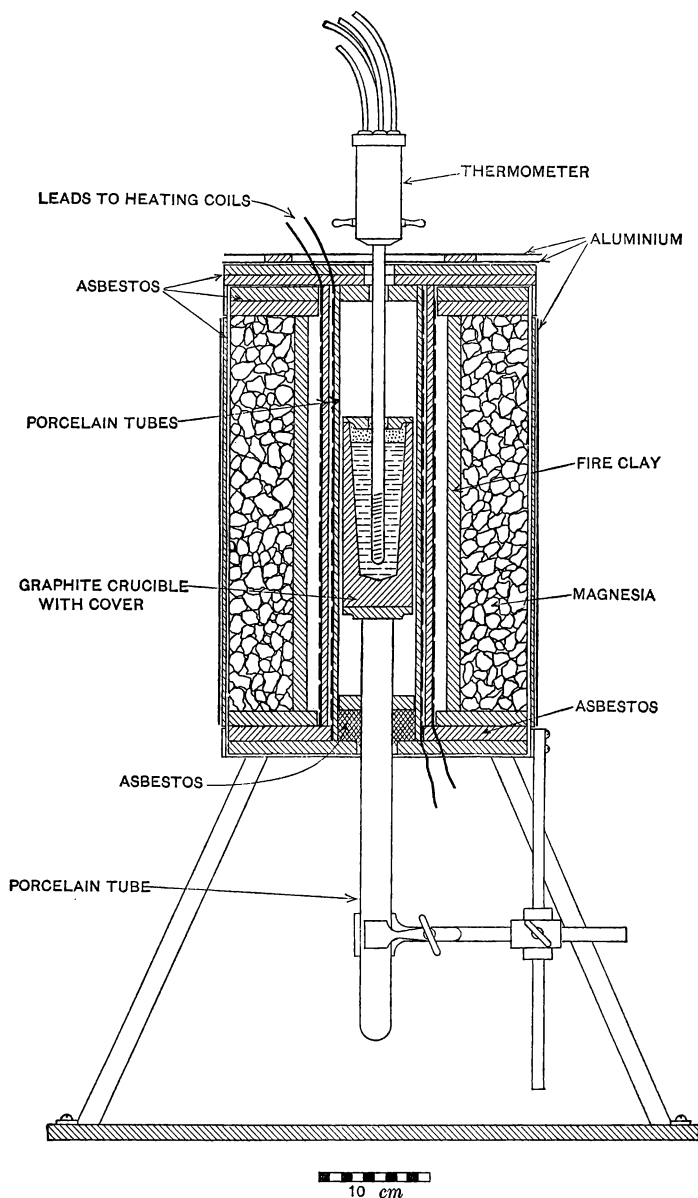
With the exception of the silver from the United States Mint in Philadelphia, the copper from the Baltimore Copper Works, and the aluminium from the Pittsburgh Reduction Works, all the materials were purchased without specifying other than the usual purest product of the several chemical manufacturers. All the metals were protected from oxidation and the measurements were carried out as explained below. A very sharp freezing point, i. e., a constant temperature extending over some minutes, can be obtained with all the metals used except aluminium, if they are pure and of sufficient bulk. This fact eliminates all ambiguity in the definition of the temperature of a freezing point.

Electric Furnaces.—The freezing and melting points were taken in well-insulated resistance furnaces, mounted vertically, with two separate heating coils of platinum ribbon, 0.01 mm thick and 2 cm wide, wound on porcelain tubes, with the turns so spaced as to allow the establishment of a uniform temperature over a considerable length of furnace by using a separate rheostat control in each circuit. The arrangement of a furnace containing a crucible of metal with a thermometer in place is shown in Plate III.

Furnace Manipulation.—With the above set-up, it is possible to hold a metal at its freezing point for an indefinite time. In practice it was the custom, when taking a freezing point, to so adjust the currents in the heating circuits as to obtain a freeze extending over ten to fifteen minutes, during which interval, for a pure metal, the temperature would usually remain constant to a few hundredths of a degree. At the higher temperatures it was deemed desirable to prevent overheating of the thermometer head in order to reduce to a minimum the thermal emf's. This was effected by passing an air current through the double-walled head of the thermometer. When this was done the head could be grasped by the hand for stirring without discomfort. In order to avoid breaking the porcelain containing tube, it was preheated in a gas furnace before introducing it into the metal.

Crucibles.—Two kinds of crucible were used, of Dixon graphite and of Acheson graphite, the latter being turned from rods. The crucibles were about 13.5 cm inside depth, 5 cm inside diameter at the top, and 3.5 cm at the bottom, holding therefore about

PLATE III



CRUCIBLE RESISTANCE FURNACE

180 cc of metal, or about 1.5 kilograms. A chemical analysis of the Acheson graphite, cut from several crucibles, was made by Dr. W. F. Hillebrand of this bureau. The total ash was found to be 0.3 per cent, one-half of which was Fe_2O_3 (equivalent to 0.11 per cent Fe). No differences in the observed values of any of the freezing points could be traced to the effects of impurities introduced from the crucibles.

Surface Oxidation and sublimation of the metals were practically eliminated, even for antimony, by putting a layer of powdered graphite and a graphite crucible cover over the surface.

Depth of Immersion.—Measurements were usually taken with the thermometer immersed to within 1 cm of the bottom of the crucible. The effect of varying depths of immersion was carefully studied by taking several consecutive freezes in each metal, everything else remaining the same. No difference greater than a few hundredths of a degree in the freezing points could be detected when the thermometer was raised 2 cm from its normal position in zinc, antimony, silver, and copper. Raising a thermometer 4 cm in zinc gave a freezing point too low by $1^\circ.3$.

Undercooling.—To avoid any considerable undercooling, as well as to insure a more certain uniformity of temperature over the length of the thermometer coil, especially when working with such metals as antimony, which, when left quiet, naturally undercools considerably and which has also a low heat conductivity, it was the practice to stir the liquid metal as long as possible by means of the thermometer itself.

Most, if not all, of these metals develop some undercooling, which is in general the more marked the higher the temperature to which the metal is brought above its freezing point, or the more rapid the rate of cooling, unless the liquid metal is stirred. When undercooling occurs, the temperature to which the metal rises is not appreciably different from the freezing point as determined when undercooling is practically eliminated by stirring.

The Observations.—In tables 3 to 10 in the appendix are given the necessary data relating to all the freezing and melting point determinations with thermometers of *pure platinum* [$\delta = 1.50$ to 1.51], of *impure platinum* [Therm. No. 1787A ($\delta = 1.57 \pm$) and No. 1787E ($\delta = 1.80$)] and of *palladium* [Therm. No. 1787G ($\delta =$

2.88±)]. No observations are omitted except a single day's work on copper, when a leak was found in the bridge circuits.

It will be seen that the thermometers of impure platinum and the palladium thermometer invariably give a lower temperature, t , than those of pure platinum for temperatures below the S. B. P. and a higher temperature above the S. B. P.

Tin.—Eleven determinations of the freezing point were made on three samples of tin: Eimer & Amend C. P. rods of 1905 in a Dixon graphite crucible, Baker & Adamson C. P. rods of 1905 in a Dixon crucible, and "Kahlbaum" tin of 1909 in an Acheson graphite crucible. The essential details of these determinations are given in Tables 3, and 12 (p. 174), from which it would appear that on the scale of the resistance thermometer of pure platinum the freezing point of "Kahlbaum" tin is $231^{\circ}.92 \pm 0.02$, and that of each of the other two samples is about $0^{\circ}.05$ lower. Undercooling, which is considerable when the metal cools without stirring from a relatively high temperature, could not be completely eliminated by stirring for any of these samples.

Measurements made with *Pt-Rh* and *Pt-Ir* thermocouples in 1903-1905, assuming "Kahlbaum" tin to freeze at $231^{\circ}.92$, gave the following results:

F. P. of Tin.—Thermoelectric Determinations.

Source	Number of Couples Used	Number of Dets.	Freezing Point
"Kahlbaum"	4	6	$231^{\circ}.92$
Eimer & Amend (Metallic, Henderson Bros.).....	4	8	$231^{\circ}.99$
Eimer & Amend C. P. sticks	4	4	$231^{\circ}.85$
Baker and Adamson C. P. sticks	2	2	$231^{\circ}.97$

In these earlier experiments no undercooling was observed for the Baker & Adamson tin.

The freezing point of tin is therefore readily reproducible to within $0^{\circ}.1$ with material purchased from various chemical firms. Even "Metallic" tin, i. e., tin of commercial purity, seems to have about the same freezing point as the pure metal. This may be due to the fact that the effect of some impurities on the freezing

point of tin is first to lower this temperature very slightly, and then to raise it with increasing percentage of impurity. This would explain why the commercial tin appears to have the highest freezing point of all the samples. Its freezing point curve is less flat than for the purer tin, as should be the case when impurities are present.

Cadmium.—The samples of cadmium used were “Kahlbaum,” Baker & Adamson, and J. T. Baker,* all purchased in 1909. As shown in Tables 4 and 12, the freezing point of the “Kahlbaum” sample is $321^{\circ}.01 \pm 0.04$; of the Baker & Adamson metal, $320^{\circ}.39$, and of J. T. Baker’s, $320^{\circ}.44$.

The undercooling of cadmium before freezing can be nearly eliminated by stirring.

Lead.—Three samples of this metal were used, “Kahlbaum,” J. T. Baker’s “Lead Granulated, Test Lead,” stated to contain no silver; and Baker & Adamson’s “Test Lead, free from Ag,” all of 1909. As shown in Tables 5 and 12, the mean value of the freezing point of lead is $327^{\circ}.43$, the determinations for a single sample agreeing to $0^{\circ}.03$. The “Kahlbaum” lead is about $0^{\circ}.3$ lower than Baker’s and $0^{\circ}.2$ lower than Baker & Adamson’s. Lead undercools slightly even with stirring.

Determinations with four thermocouples, in 1903 and 1905, of the freezing point of Eimer & Amend C. P. lead gave from $327^{\circ}.3$ to $327^{\circ}.5$. Pig lead had a less sharp freezing point some $1^{\circ}.5$ low, and lead pipe was about 2° low. The data for these commercial samples of lead are given for the reason that lead hardening and annealing baths are used in the industries, and the question has often arisen whether this grade of lead can be used as a control point for industrial pyrometers.

Zinc.—More determinations of the freezing point were made with zinc than with any other metal, and the results obtained with the platinum thermometers, as shown in Tables 6 and 12, indicate a remarkable degree of concordance, both for determinations by means of different thermometers of the freezing point of the same material and for determinations with material from different sources. Thus, samples of “Kahlbaum” zinc bought in 1903, 1907,

* Marked “Cadmium Metal” with an indicated analysis showing iron 0.004 per cent and arsenic, trace.

and 1908 give a freezing point of $419^{\circ}.37 \pm 0.02$, one of these being in Acheson and the others in Dixon graphite crucibles. One of these samples had been in use in the same crucible since 1903 for calibrating thermocouples.

Determinations of the freezing points were also made on the following samples:

Eimer & Amend "C. P. in sticks," two crucibles of 1903 and 1905 mixed, and Baker & Adamson "C. P. in sticks." The former,* which had been repeatedly used, gave a lower value, $419^{\circ}.25$, and the latter a slightly higher value, $419^{\circ}.42$, than the Kahlbaum zinc.

Measurements made with thermocouples in 1905 gave the following results for the freezing point taken in a gas furnace:

F. P. of Zinc—Thermoelectric Determinations.

Source	Number of Couples Used	Number of Dets.	Freezing Point
"Kahlbaum".....	8	10	$419^{\circ}.3$
Eimer & Amend "C. P. in sticks".....	8	12	$419^{\circ}.0$
Eimer & Amend "Metallic".....	4	6	$416^{\circ}.8$
Baker & Adamson "C. P. in sticks".....	4	8	$419^{\circ}.1$

Other thermoelectric determinations made in 1907 gave equally concordant results.

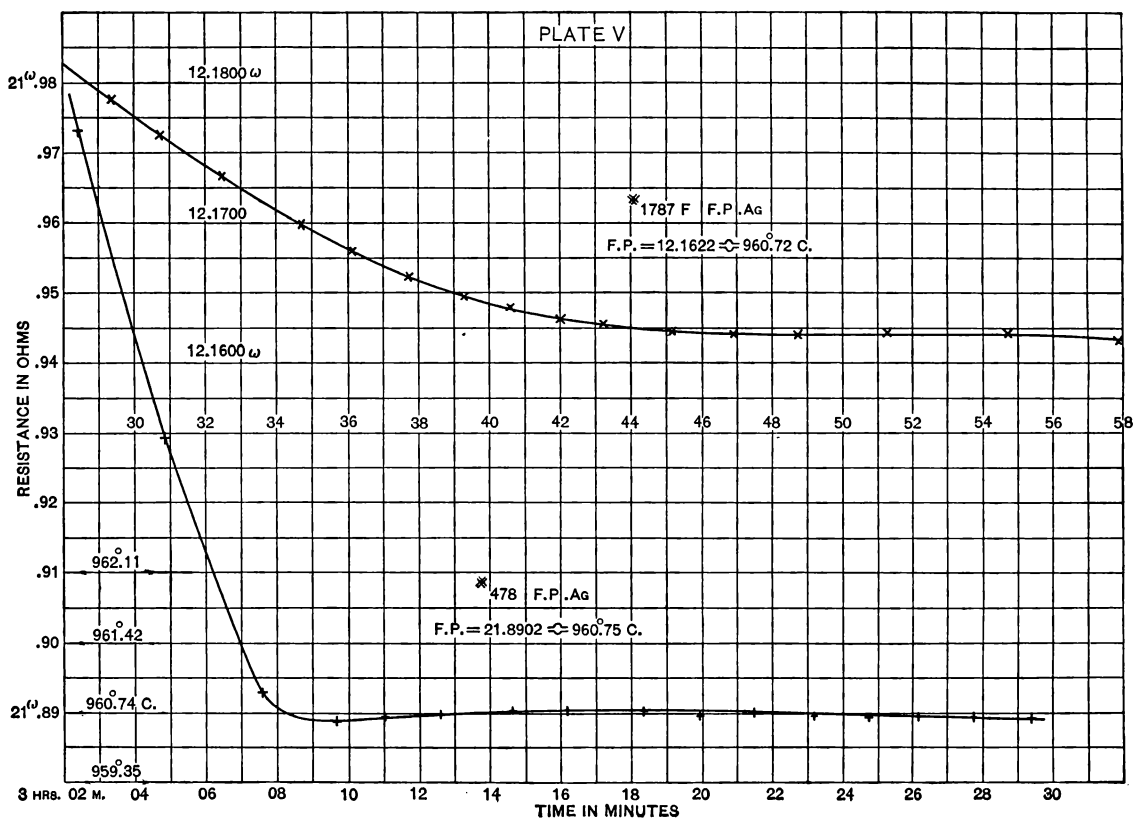
The undercooling of zinc is almost inappreciable, being only a few hundredths of a degree unless the metal has been considerably overheated. It is therefore unnecessary to stir when taking a freeze.

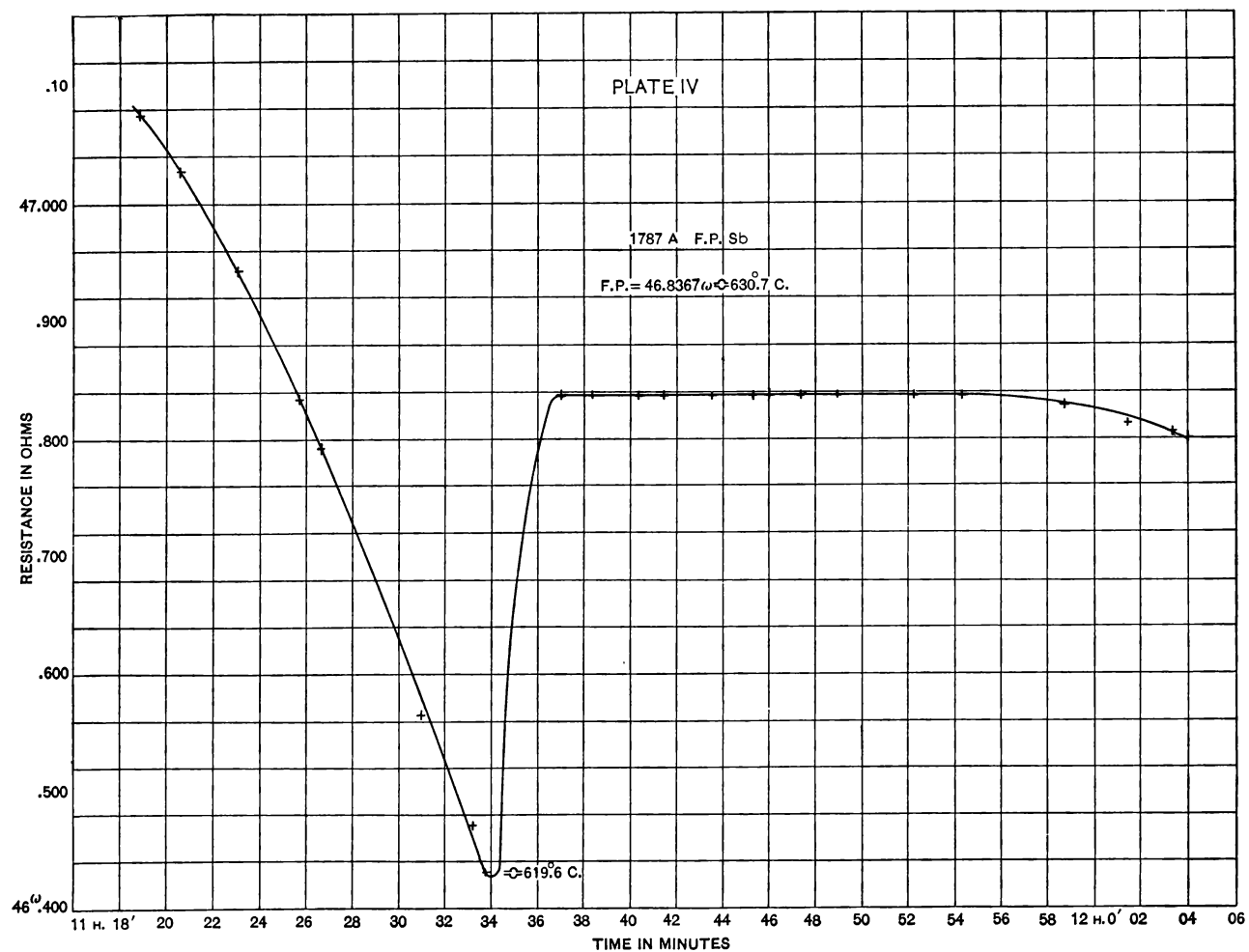
A considerable number of determinations of the melting point of zinc, made with the resistance thermometers, gave results agreeing to $0^{\circ}.1$ with the freezing point, even for fairly rapid heating.

Antimony.—"Kahlbaum" antimony† has been used as a fixed point for a number of years with satisfaction, and this material,

* Dr. E. T. Allen finds, for two samples from different lots of Eimer & Amend zinc "C. P. in sticks," total impurities 0.063 and 0.049 per cent, with lead as the principle component, it being 0.051 and 0.041; see *Phys. Rev.* **19**, p. 177; 1904; *Am. Jl. Sci.* **26**, p. 454; 1908.

† An analysis of this material has been published; Fritz Henz, *Inaugural Dissertation*, Zurich, 1903 (Pub. in Leipzig), $\text{Fe}=0.012$, $\text{Cu}=0.004$, $\text{Pb}=0.003$ per cent.





purchased at different times over a period of six years, appears to be a very uniform product as determined by the constancy of the freezing point. Antimony from other sources, however, so far as our experience goes* is not to be relied upon as being of sufficient purity to warrant its use for defining a fixed temperature by means of its freezing point.

Antimony oxidizes very rapidly in the air and somewhat even in the presence of graphite, but this oxidation does not appear to affect the freezing point appreciably. This metal normally undercools very considerably, even to the extent of 30° or more, but whether this is nearly eliminated or not, by stirring, practically the same temperature of freezing is obtained. In Plate IV is given a freezing point curve for antimony when the metal is stirred with the thermometer. The melting point agrees with the freezing point to within $0^{\circ}.1$, and they approach each other as the slowness of freeze and melt are increased.

Seven determinations of the freezing point of three samples of "Kahlbaum" antimony with two resistance thermometers gave $630^{\circ}.71 \pm 0.13$. (See tables 7 and 12.)

Thermoelectric determinations in gas and electric furnaces of the freezing point of antimony from various sources gave the following results, assuming Sb "Kahlbaum" = $630^{\circ}.7$:

F. P. of Antimony—Thermoelectric Determinations.

Source	Date of Purchase	Number of Couples	Number of Dets.	Temperature
"Kahlbaum".....	1903-1909	6	8	$630^{\circ}.7$
Eimer & Amend C. P.	1903	4	6	$621^{\circ}.0$
Eimer & Amend "Metallic".....	1903	2	6	$619^{\circ}.1$
J. T. Baker† "Antimony metal".....	1909	2	4	$626^{\circ}.2$
Baker & Adamson.....	1909	2	4	$626^{\circ}.0$
Merck "Highest Purity".....	1907	1	2	$624^{\circ}.7$

* See also Day and Allen, Phys. Rev. 19, p. 177; 1904.

† Analysis on label: Fe=0.01, Cu=0.005, As=0.002.

Pb=0.002, Zn=0.000, Sn=0.000 %.

It is evident that for this metal the determination of the freezing point is a very delicate test of the purity of the sample.

Silver.—Only one sample of silver was used, purchased in the form of drops from the Philadelphia mint, as the purest silver obtainable. Its analysis was given as 99.995 fine. An analysis of another sample of this grade of silver has also been made by Dr. E. F. Allen,* who found only 0.0032 per cent impurities, one-third of which was iron.

Eight determinations with three resistance thermometers gave a freezing point of $960^{\circ}.88 \pm 0.16$. See Tables 8 and 12, which also show how constant the indications of these thermometers remain even at such a high temperature, as is seen from a comparison of the calibrations made before and after the freezing point determinations.

The silver point is very sharply defined, as might be expected from the high conductivity and purity of this metal, and although there is some slight undercooling if the freezing point is approached from a high temperature this disappears entirely with stirring or when the rate of cooling before freezing is small. Sample freezing point determinations with and without undercooling are shown in Plate V.

Copper.—As an upper limit for the exact determination of temperatures by the usual type of platinum resistance thermometer, the freezing point of copper was chosen, and an examination of the data in Tables 9 and 12 shows that the precision attainable at this temperature is already considerably less than at the silver F. P., though still sufficiently good for work within about 1° . On account of the changes in the constants of the platinum thermometer, which occur at this high temperature, it is better to depend on the thermocouple for temperature measurements in this region, since the advantages of very great sensitiveness of the resistance instrument are masked by these changes.

Copper from several sources was used. The only product furnished with an analysis was that obtained from the Baltimore Copper Smelting and Rolling Company, which company very

* Am. Jl. Sci. (4), **26**, p. 456; 1908.

kindly furnished us, in 1908, with a specially prepared electrolytic sample in the form of cathode plates.*

This copper had a freezing point of $1083^{\circ}.0 \pm 0.5$ from seven determinations with three resistance thermometers.

Kahlbaum electrolytic copper in a Dixon crucible gave $1081^{\circ}.4$; a Dixon crucible filled with copper of uncertain origin, but which had been in use for several years, $1082^{\circ}.7$; and Eimer & Amend's "Copper,† C. P. drops, cooled in hydrogen" of 1909, $1083^{\circ}.9$ in Acheson graphite for two determinations with the same thermometer.

It should be noticed that very slight changes in either the fundamental or the difference coefficient will cause considerable differences in the computed temperatures at the copper point. Thus a change in δ from 1.50 to 1.51 is equivalent to about 1° temperature difference at 1083° .

When the metal is protected from oxidation, the copper freezing point is very sharp, the undercooling being at most only $0^{\circ}.2$ to $0^{\circ}.3$ (see Plate VI), and often inappreciable.

The copper freezing point has been used with entire satisfaction as a fixed point in calibrating pyrometers at the Bureau of Standards since 1903. This metal is so readily obtained sufficiently pure, at small cost, that its freezing point is particularly adapted for use as a standardizing temperature.

Even commercially pure copper in the form of bars or wire usually has a freezing point within a degree of that of the specially prepared material.

A great many determinations of this temperature were made by means of thermocouples from 1903 to 1909, on material from various sources, both in gas and electric furnaces. Among the kinds of copper used were copper wire, copper in bars, Baker &

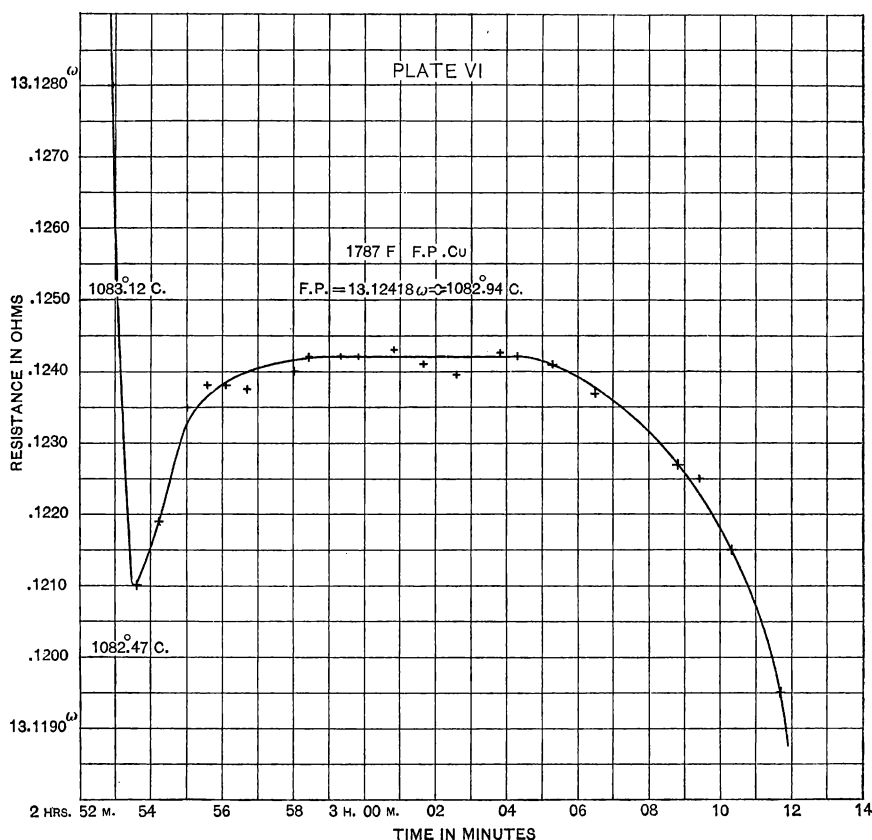
*Mr. Wm. M. Pierce reports the analysis of this copper as follows.

Ag, Au, As, Sb, Pb, none; Fe, S, trace; Cu 99.995 per cent by direct electrolytic test.

"In making the test for Ag, As, Sb, Fe, S, and Pb, large quantities of the samples were used: six assay tons for Ag, 300 grams for As and Sb, and 200 grams for Fe, S, and Pb."

† An analysis of this grade of copper is given by Day and Allen, *Phys. Rev.* **10**, p. 180, 1904, showing total impurities = 0.040 per cent, of which Fe = 0.011, Zn = 0.010, Ag = 0.018, and Pb = 0.001.

Adamson's "copper, C. P. reduced by hydrogen," and Kahlbaum's electrolytic granulated. No certain difference could be detected in the freezing points of any of these materials, but when using the powdered copper reduced by hydrogen great care has to be taken not to oxidize the metal. When determinations are made in a gas furnace the metal tends to oxidize partially, even in graphite,



unless a reducing flame is used. The presence of the soluble oxide, Cu_2O , lowers the copper freezing point until, at saturation, the copper-cuprous oxide eutectic point is reached at about $1,063^{\circ}$. In a gas furnace, however, unless special precautions are taken, neither the freezing point of pure copper nor that of the eutectic of $\text{Cu-Cu}_2\text{O}$, is obtained alone, but an intermediate temperature, which is usually nearer the Cu F.P. The eutectic point can always be detected except for very small percentages of the oxide.

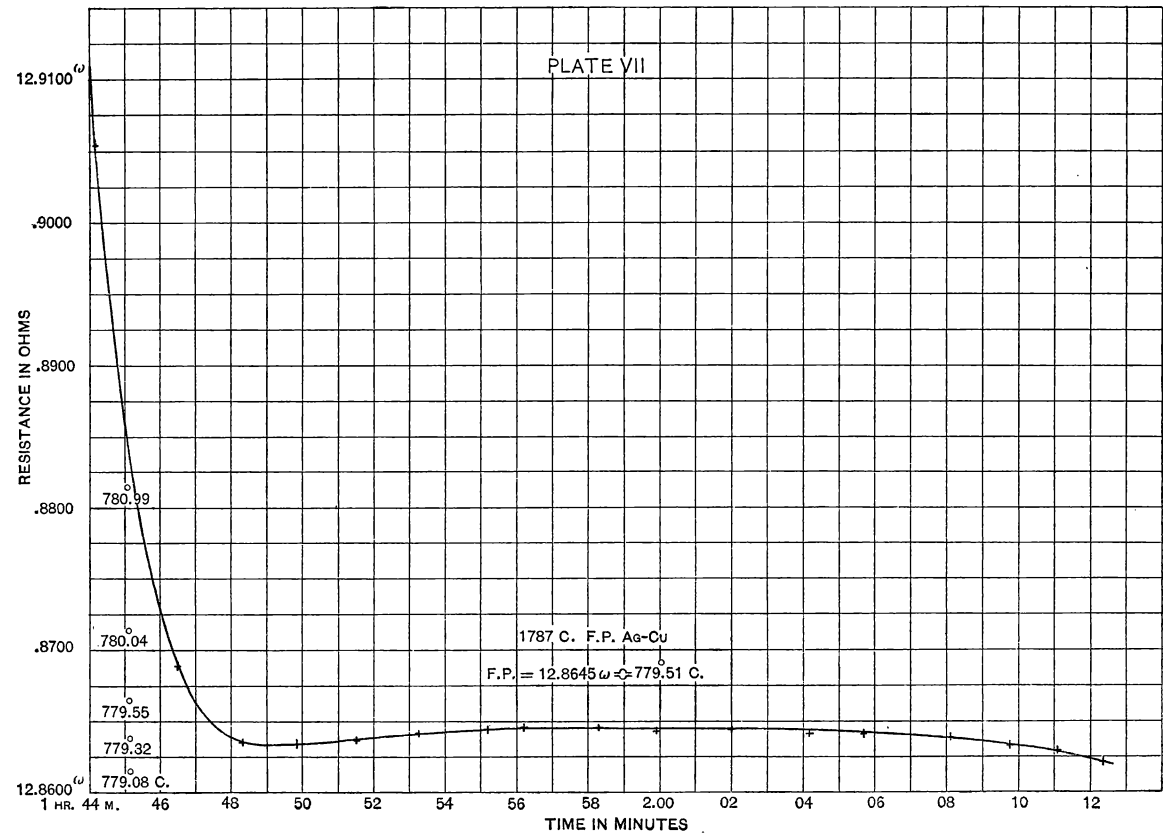
Silver-Copper Eutectic.—With the exception of the “Kahlbaum” antimony point at $630^{\circ}.7$ there is no satisfactory fixed point between zinc and silver or copper, and it is especially desirable to have such a temperature located in the neighborhood of 800° . We have found that the Ag-Cu eutectic alloy ($\text{Ag}_3 \text{Cu}_2$) meets this requirement in a very satisfactory way.

A preliminary study of the conditions necessary for the constancy of this eutectic temperature was made with a thermocouple and gave the following results, using as components silver from the Philadelphia mint, and copper, “C. P. drops cooled in hydrogen,” from Eimer & Amend:

Silver-Copper Eutectic—Thermoelectric Determinations.

Per cent Ag-Cu	Composition in Grams Ag-Cu	Range of Transformation Temperature	Alloy, cooled from	Eutectic Temperature	Remarks
50-50	150-150	$1^{\circ}.5$	828°	$777^{\circ}.50$	
50-50	150-150	.3	812	$777^{\circ}.89$	
50-50	150-150	$1^{\circ}.1$	821	$778^{\circ}.72$	
50-50	150-150	.2	906	$778^{\circ}.23$	{ Slight undercool.
50-50	150-150	.5	977	$777^{\circ}.89$	
50-50	150-150	.4	857	$776^{\circ}.33$	Stirred.
50-50	150-150	.4	929	$776^{\circ}.43$	Stirred.
60-40	255-150	.8	945	$773^{\circ}.07$	Stirred.
60-40	255-150	.7	818	$774^{\circ}.99$	
71-29	355-145	.0	825	$779^{\circ}.49$	{ Slight but long undercool.

These results show that the presence of an excess of copper is disadvantageous in two ways: first, the eutectic temperature is not sharply marked, as shown by the column headed “range of transformation temperature”; and second, the value of the eutectic temperature observed depends upon the degree of mixture of the excess component with the eutectic, as is shown by comparing the freezes taken with and without stirring. Although the lowest temperature is obtained when stirring the noneutectic mixture, yet this appears to be considerably below the true eutectic temperature, as obtained when the alloy of the eutectic composition



is used. The reason for this retarding of the separation of the eutectic by the admixture of copper is uncertain, but it may be due to the greater conductivity of the copper.

For the measurements with the resistance thermometers, an Acheson crucible was filled with 1200 grams of the alloy 71Ag-29Cu; measurements were also made with the eutectic composition 72Ag-28Cu, and with 73Ag-27Cu. The observations are shown in Table 10 and a summary is given in Table 12. The best value to assign to the eutectic temperature from these observations is $779^{\circ}.20 \pm 0.15$.

This temperature is as sharply marked as the freezing point of most of the pure metals; a sample freezing curve is given in Plate VII. It will be noticed that the undercooling is slightly different from that of a pure metal. This undercooling is of the order of $0^{\circ}.1$; but instead of the temperature rising rapidly to its maximum, as in the case of a metal, the rise is very slow, taking almost half of the time of duration of a freeze. This slight undercooling could not be avoided by stirring.

Copper-cuprous Oxide Eutectic.—Some measurements were made with thermocouples in order to determine the temperature and reproducibility of the Cu-Cu₂O eutectic point in both gas and electric furnaces. In agreement with the work of Heyn,* we find that the cooling curve for copper in the presence of its oxide when the mixture has been heated in a gas furnace does not, in general, correspond to the composition of the components originally put into the crucible, on account of the varying reducing and oxidizing action of the atmosphere in such a furnace. Unless the atmosphere is strongly reducing, however, the eutectic point is sharp to within 1° at 1063° . When the alloy of the eutectic composition (Cu₂O = 3.5 per cent) is introduced into an electric furnace of the type used in these experiments (Plate III), the eutectic temperature is as sharply defined as the freezing point of pure copper and is normally preceded by an undercooling of 2 or 3° . From six observations we find this eutectic temperature to be $1063^{\circ}.2 \pm 0.3$ on the temperature scale of this investigation (on which F.P. of Cu = $1083^{\circ}.0$).

*C. Heyn, Kupfer und Sauerstoff, Mitt. Königl. Tech. Versuchsanstalten 18, p. 315; 1900.

Aluminium.—No determinations of the freezing point of aluminium were made with the platinum-resistance thermometer. Two series of measurements were carried out, however, with thermocouples, one in 1905 and the other in 1909, on the same material but with separate samples. The Pittsburgh Reduction Company in 1903 furnished us with a specially prepared sample of exceptional purity. The analysis of this sample was given as: Al, 99.78; Si, 0.12; Fe, 0.10; Cu, trace. The other samples were in the form of bars from Eimer & Amend and from Kahlbaum.

On the temperature scale used in this investigation the following values were found for the freezing points of these aluminium samples:

F. P. of Aluminium—Thermoelectric determinations.

Source	Al F. P.
Pittsburgh Reduction Co.	658.0
Kahlbaum, bars.	654.5
Eimer & Amend, bars.	656.0

The observations of 1905 and 1909 agree to within 1°, which is about the uncertainty in successive observations with the same thermocouple. The freezing point of aluminium is not sharp, even for the purest of these samples, so that this metal is the least adapted of all we have tried for defining a fixed temperature. Small quantities of an impurity, such as iron or silicon, lower the aluminium freezing point very considerably.

Metals from Different Sources.—The data on the freezing points of the metals from various sources may be grouped together so as to show the reproducibility of the fixed temperature that each defines. This has been done in Table 11, in which are indicated, for each sample, the source of the material, a brief description, the date of purchase, and the freezing point. In the last column is given the “reproducibility factor” for each metal, which we have taken as the average deviation of a single observation in degrees C from the mean F. P., including only those samples of C. P. grade.

TABLE 11.

Reproducibility of Fixed Points by Means of Metals from Different Sources.

Metal	Source	Description	Date of Purchase	Freezing Point °C.	Reproducibility Factor
Tin	Kahlbaum	"Kahlbaum"	1903, 1905, 1909	231°.92	0°.06
	Eimer & Amend	"C. P. in sticks"	1905	231°.85	
	Henderson Bros.	"Metallic" in bars	1905	*(231°.99)	
	Baker & Adamson	"C. P. in sticks"	1905	231°.97	
	Baker & Adamson	"C. P. in sticks"	1909	231°.85	
Cadmium	J. T. Baker Co.	"Mossy metal"	1909	231°.84	0.26
	J. T. Baker Co.	"Cadmium metal"	1909	320°.44	
	Kahlbaum	"Kahlbaum"	1909	321°.01	
	Baker & Adamson	"Cadmium"	1909	320°.39	
	Baker & Adamson	"Test lead free from Ag"	1909	327°.47	
Lead	J. T. Baker Co.	"Lead granulated, test lead"	1909	327°.58	0.10
	Kahlbaum	"Kahlbaum"	1909	327°.26	
	Eimer & Amend	"C. P. lead"	1905	327°.40	
		Pig lead	1905	(326°.0)	
		Lead pipe	1905	(325°.5)	
Zinc	Eimer & Amend	"C. P. in sticks"	1903-1905	419°.25	0.06
	Eimer & Amend	"Metallic"	1903	(416°.8)	
	Kahlbaum	"Kahlbaum"	1903, 1907, 1908	419°.37	
	Baker & Adamson	"C. P. in sticks"	1903	419°.42	
	Kahlbaum	"Kahlbaum"	1903-1909	630°.71	
Antimony	Eimer & Amend	"C. P."	1903	621°.0	2.3
	Eimer & Amend	"Metallic"	1903	(619°.1)	
	J. T. Baker Co.	"Antimony metal"	1909	626°.2	
	Baker & Adamson	"Antimony"	1909	626°.0	
	Merck	"Highest purity"	1909	624°.7	
Aluminium	Pittsburgh Reduction Co.	99.78 aluminium	1903	658°.0	1.2
	Kahlbaum	In bars	1905	654°.5	
	Eimer & Amend	In bars	1905	656°.0	
	Eimer & Amend	"C. P. drops, cooled in hydrogen."	1909		
	Kahlbaum	Electrolytic, granulated	1907		
Copper	Baker & Adamson	"C. P. reduced by hydrogen"	1905	1083°.0	†1.0
	Baltimore Copper Works.	99.995 copper	1908		
		In bars			
		Wire			

*Numbers in parentheses are not included in computing the "Reproducibility Factor."

†This uncertainty is due more to oxidation than to impurities.

TABLE 12.
Scale of Platinum Thermometer.
CALENDAR FORMULA.
[Calibration data: Ice, steam, sulphur (S. B. P. = 444° 70'.)]

Metal	$\delta=1.505$ $c=0.00385$ to 0.00390					$\delta=1.570 \dagger$ $c=0.00206$ to 0.00210				$\delta=1.803 \dagger$ $c=0.00173$			Heycock and Neville, corrected to S. B. P. =444.70	Gas scale Holborn and Day
	Temp.	Range deg. C.	Number of—			Temp.	Range deg. C.	Number of—		Temp.	Range deg. C.	No. of dets.		
			Dets.	Samples	Thermometers			Dets.	Samples					
Sn.....	{ 231.92	0.04	4	1K	2	{ 231.82	0.08	4	3				231.94	
	231.89	.09	7	4	2									
Cd.....	{ 321.01	.02	2	1K	2	{ 320.95	.02	2	1					321.7
	320.62	.65	5	3	2									
Pb.....	{ 327.43	.35	6	3	2	{ 327.26	.03	2	1K	1				326.9
	327.26	.03	2	1K	1									
Zn.....	419.37	{ .04	6	3K	1	{ 419.32	{ .08	8	3K				419.4	419.0
		.11	9	3	1									
	419.30	{ $\delta=1.516$				{ 419.28	{ $\delta=1.551$							
		.06	5	1K	1			1	1K					
Sb.....	630.71	.57	7	3K	2	631.25	.17	4	3K	632.65	0.09	2	*630.0	630.6
Sb M. P.....	630.80	.15	4	2K	1	631.06	.08	3	3K					
Ag-Cu 71-29.....	779.55	.08	3	1	2	781.21	.06	2	1					
Ag-Cu 72-28.....	779.18	.45	6	1	3								779.0	
Ag-Cu 73-27.....	779.23	.30	3	1	2					784.61	.14	2		
Ag.....	960.88	.39	8	1	3	966.21	.02	2	1	975.28	.08	2	961.9	961.5
Cu.....	{ 1083.04	1.61	7	B. C. W.	3	{ 1091.92	1.15	4	3	1106.2		1	1082.0	1084.1
	1082.88	2.26	13	4	4									

"Range"=difference between highest and lowest F. P. determinations. "Number of samples" refers to materials from different chemical firms.
 "K"="Kahlbaum." B. C. W.=Baltimore Copper Works purest product.
 *This is not antimony "Kahlbaum."
 \dagger Thermometers of impure platinum.

IV. THE SCALE OF THE PLATINUM THERMOMETER.

The scale defined by the platinum thermometer, calibrated in ice, steam, and sulphur, may be best expressed in terms of the freezing points of the pure metals. In Table 12 are grouped the results of the determinations of a series of such freezing points as obtained with platinum of several grades of purity.

The temperatures determined with thermometers of pure platinum, for which the fundamental coefficient is 0.0039, and the difference coefficient 1.50, are evidently in very close agreement with temperatures on the gas scale. When, however, thermometers are made of impure platinum having $\delta = 1.57$ or 1.80, as shown in the table, and the Callendar method of reduction is used, the agreement of the two temperature scales, platinum resistance and gas, is seen to be far from satisfactory.

The precision at the various fixed points is also shown in Table 12, where the column headed "Range," when a single sample is used, gives a measure of this precision in terms of the difference between the highest and the lowest determinations, omitting none.

An idea of the reproducibility of these fixed points in terms of samples from several sources is also given by this same column, "range," for those cases in which several samples are used, as indicated in the column headed "number of samples."

The freezing points determined by Heycock and Neville²⁷ with the resistance thermometer are also included in Table 12. They are computed for S. B. P. = 444°.70 instead of for S. B. P. = 444°.53.

Correction for Thermometers of Impure Platinum.—From the measurements with thermometers made of wires with various difference coefficients it is possible to find, graphically, the corrections necessary to reduce the indications of any thermometer of impure platinum to the scale of the thermometer of pure platinum.

Such a series of corrections is given in Table 13 for values of δ varying from 1.525 to 1.900, and for temperatures from 200° to 1100°. The corrections are in degrees C and reduce the apparent temperatures, calculated by the Callendar formula from readings at the ice, steam, and sulphur points, to the scale of the pure platinum thermometer of $\delta = 1.505$.

Precision of Measurement with Platinum Thermometer.—The range of values of all the F. P. determinations taken with a single thermometer in a given crucible of metal will serve as a measure of the precision of measurement attainable with that thermometer at the temperature in question.

In Table 14 is given the range of temperatures at a series of fixed points obtained with thermometers 1787C. and 1787A., the former of pure platinum and the latter of platinum for which $\delta = 1.57$. It will be seen that the greatest range observed at any temperature up to the copper point is $0^{\circ}.4$, and the average range is only $0^{\circ}.11$, or practically within the reproducibility of many of these metals from new samples. The thermometers of impure platinum give as good precision as those of pure platinum; but it is to be remembered that the constants of the impure metal undergo the greater changes with heating, necessitating more frequent recalibration of such thermometers.

Modification of Callendar's Formula.—For thermometers of impure platinum, used at high temperatures, Callendar's formula may be modified so as to give results agreeing very closely with temperatures on the scale of the thermometer of pure platinum or on the gas scale. If a fourth calibration point is taken, such as the copper or silver freezing point, the difference coefficient may be written in the form, $\delta = a + bt$. Platinum temperatures, pt , are then computed in the usual way; but the true temperatures are now obtained by using the interpolated value of δ for that temperature, which need be known only approximately.

Using the silver freezing point as the fourth calibration temperature to determine a and b , we have calculated several of the freezing points of the metals by this method for thermometers of impure platinum and of palladium, as shown in Table 15, where are also grouped for comparison the temperatures on the scale of pure platinum and for the thermometers of impure platinum, as reduced by the ordinary Callendar method of calibration (*i. e.*, at ice, steam, and S. B. P.).

It is evident from an inspection of the table that for the impure platinum wires the calibration at four points by the modified Callendar method reproduces the temperature scale of pure platinum in the range 0 to 1100° more closely than the latter is known in terms of the gas thermometer scale.

TABLE 15.

Temperature Scale for Impure Platinum and Palladium. Modification of Callendar Formula.

$$[\delta = a + bt \text{ (from S. B. P. and Ag F. P.)}]$$

Metal	F. P. on scale of Therm. of Pure Pt $\delta = 1.505$	Thermometers of Impure Pt				Pd Therm.	
		No. 1787A		No. 1787E		No. 1787G	
		$\delta = 1.570$ from S. B. P. calibration	$\delta = a + bt$ $a = 1.608$ $b = -0.04853$	$\delta = 1.803$ from S. B. P. calibration	$\delta = a + bt$ $a = 1.900$ $b = -0.04219$	$\delta = 2.890$ from S. B. P. calibration	$\delta = a + bt$ $a = 3.020$ $b = -0.04327$
Sn.....	231°.90	231°.82	231°.89	231°.62	231°.80
Cd.....	321°.01	320°.95	321°.05	320°.52	320°.74
Sb.....	630°.7	631°.2	630°.5	632°.6	630°.8	633°.0	629°.2
Ag-Cu.....	779°.2	781°.2	779°.1	784°.6	779°.3	787°.6	777°.1
Ag.....	960°.9	966°.2	(961°.0)	975°.3	(961°.0)	992°.5	(961°.0)
Cu.....	1083°.0	1091°.9	1083°.3	1106°.2	1082°.3	1152	1091

For the palladium thermometer the modified Callendar formula is not sufficient to give true temperatures, although up to the silver point the greatest divergence is only about 2° .

Other Calibration Formulæ.—It was thought worth while to see if other calibration formulæ would represent the resistance-temperature relation for platinum and palladium. Dickson³⁸ suggested the formula for platinum:

$$(R + a)^2 = p(t + b)$$

where a , p , and b are constants, and applied it to the observations of several observers. This formula possesses the apparent theoretical advantage over Callendar's of not requiring a maximum for the resistance of platinum. We find that, for thermometers of platinum or of palladium, calibrated at 0° , 100° , and S. B. P., Dickson's formula does not give satisfactory values of the freezing points. For example, thermometer 1787C., of pure platinum, calibrated in this way gives $232^\circ.13$ for tin, and $1051^\circ.3$ for copper. For thermometer 1787G, of palladium, the freezing point of copper would be 1005° , and of the Ag-Cu eutectic, $754^\circ.4$.

The adaptability of other types of formula has been discussed at length by Callendar.¹²

Other Methods of Calibration at High Temperatures.—The temperature scale defined by the resistance thermometer (of pure platinum), as used in this investigation and as almost universally used, is based on calibration in ice, steam, and sulphur. Higher temperatures are then determined by an extrapolation of the resistance-temperature relation thus found. If the thermometer is intended for use at temperatures much above S. B. P., it could be standardized at other known temperatures, for example, ice, S. B. P. or F. P. of Zn, and the F. P. of Ag or Cu, in which case extrapolation would be avoided. Instead of representing the parabolic resistance-temperature relation by the Callendar equation, it would be represented by the equivalent equation:

$$R_t = R_0(1 + at - \beta t^2)$$

This method of standardization was not used in the present investigation for the reason that the usual method (ice, steam, and sulphur) leads to values for the freezing points of Sb, Ag, and Cu, which are in such excellent agreement with the best gas thermometer determinations of these freezing points that the differences are almost, if not quite, within the accuracy of measurements attained up to the present time in gas thermometry.

V. CHANGES IN CONSTANTS OF THERMOMETERS AT HIGH TEMPERATURES.

Changes with Discontinuous Heating.—If the calibration data of each thermometer are arranged chronologically, the changes produced in the thermometric constants R_0 , FI , c , and δ , due to the various temperatures to which the instrument has been subjected, become immediately apparent. Tables 16 to 23 show these changes for the thermometers used in this investigation.

A comparison of the behavior of thermometers made of different grades of platinum and subjected to similar heat treatment shows that the constants of pure platinum undergo less change than those of impure platinum, although it is difficult to eliminate the part played by the mica frame in causing variations in these constants. It may be noted that, generally, the pure platinum shows a tendency to become less pure, as is seen by the decrease in c for thermometers 1787F, 478, 479, although for 1787C the

value of c remained practically constant, while for thermometers made of less pure platinum the value of c increased, as if the platinum were being purified by the loss by evaporation of impurities such as iridium, due to successive heatings. In support of this view it is to be noticed that the effect of continued annealing at 1100° is the same for all of the thermometers, namely, to increase the value of c .

For the thermometers of pure platinum there is a slight rise in zero (R_0) amounting only to a few tenths of a degree, while for those of impure platinum and of palladium the zero falls, and, in general, by relatively much greater amounts. For example, with 1787A, for which $\delta=1.57$, R_0 decreased by about $6^{\circ}.5$, and with 1787C, for which $\delta=1.50$, R_0 increased by less than $0^{\circ}.15$, for about the same heat treatment.

An examination of the behavior of the thermometers made of platinum of different grades shows the advantage of employing the purest metal for work of precision. It will also be seen that the changes in thermometric constants for all types are relatively small at temperatures up to the melting point of silver, but that they are often considerable at the copper point.

Thermometers 1787F and 1787G were used by both the Bridge and Potentiometer methods, and in comparing results by these two methods, Tables 17 and 22, it should be noted that the values of R_0 and FI differ, due to the different resistance units of the measuring apparatus, and also because the thermometer is not exactly the same instrument when used by the two methods.

The changes produced in δ by heating, after the thermometer is well annealed, are extremely small, and no certain conclusions can be drawn regarding the slight variations observed, as they appear for the most part to be within the limit of experimental error.

Changes during a F. P. Determination.—In determining freezing points at the higher temperatures it is necessary to determine, by calibrating the thermometer both before and after the freezing-point measurement, the changes occurring while the freezing point was being taken. These changes become important, and allowance must be made for them in work to $0^{\circ}.1$ above 700° , and in some cases at lower temperatures. This

effect increases materially with the temperature, and is the main factor in reducing the precision attainable at high temperatures. An examination of the experiments on Ag, Cu, and Ag-Cu (Tables 8, 9, and 10) shows that changes amounting to several tenths of a degree at these temperatures may arise from this cause. These changes are indicated in condensed form in Table 24, for values of both pt and t . In general, the temperature assigned to a freezing point was the mean value obtained from the calibrations before and after the freeze.

Some measurements made on two Leeds and Northrup resistance thermometers, Nos. 9837 and 9838, the coils of which were made of much larger platinum wire (diam. = 0.6 mm), show a much greater constancy at high temperatures for thermometers made of thick wire. These coils were not mounted on mica frames but were freely suspended and were thus not subject to strains resulting from thickening of the mica frame (see Plate I). The resistance in ice was quite low, about 0.11 ohm, but a much larger measuring current could, of course, be used than with the fine wire thermometers, so that the sensitiveness was not reduced in proportion to the resistance. These two thermometers were submitted to the same heat treatment. After a very thorough annealing it was found that an exposure for five hours at 1200°, three and one-half hours at 1250°, and one hour at 1300°, in an electric furnace, caused a total increase of R_0 and a change in the FI corresponding to only a few tenths of a degree (see Table 23). The error at 1200° at the close of the experiments, due to using the calibration found at the beginning of the experiments was of the order of 3°, the two thermometers showing approximately the same changes.

These thermometers were mounted in fairly clear quartz tubes, and it may be worthy of mention that while this material withstood with only slight crystallization the exposure to temperatures of 1200° or 1300° for some ten or fifteen hours, it deteriorated very rapidly after that, and developed numerous cracks and spots where transformations in the structure of the material were evident.

TABLE 25
Comparison of Scales of Platinum Thermometer and Thermocouples.

Metals	Temp. from Res. Therms.*	Temperature from Thermocouples $E=a+bt+ct^2$ †									
		P ₂ Heraeus 1905 Pt, 90 Pt-10 Rh		P ₃ Heraeus 1908 Pt, 90 Pt-10 Rh		S ₂ Johnson Matthey 1904 Pt, 90 Pt-10 Rh		S ₃ Heraeus 1908 Pt, 90 Pt-10 Rh		W ₇ Pellin Pt, 90 Pt-10 Ir	
		I	II	I	II	I	II	I	II	I	II
Zn	419°.37	†419°.37	419°.37	419°.37	419°.37	419°.37	419°.37	419°.37	419°.37	419°.37	419°.37
Sb	630 .7	630 .7	630 .9	630 .7	631 .0	630 .7	630 .4	630 .7	631 .3	630 .7	630 .7
Ag-Cu .	779 .2	779 .1	779 .2	778 .9	779 .2	779 .6	779 .2	778 .6	779 .2	779 .2	779 .2
Ag	961 .0	960 .7	960 .8	960 .8	960 .9	961 .4	961 .2	960 .6	961 .0	960 .9	960 .9
Cu	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0	1083 .0

* Scale defined by resistance thermometer of pure platinum calibrated in terms of the Callendar equation from observations in ice, steam, and sulphur (444°.70).

† Numbers in heavy type taken as calibration points for thermocouples, i. e., for computing constants a , b , and c of the equation $E=a+bt+ct^2$. The remaining two temperatures are then the freezing points on the thermoelectric scale defined by this equation.

‡ For convenience in computation, the formula $T=a+bE+cE^2$ is sometimes used as the calibration equation for a thermocouple. It should be noted that this formula does not exactly agree with the above formula, even within the range of calibration temperatures, and differs very much for higher extrapolated temperatures.

VI. COMPARISON OF RESISTANCE AND THERMOELECTRIC SCALES.

To compare these two temperature scales a number of thermocouples were calibrated at the freezing points of the metals previously used with the resistance thermometers. The results are given in Table 25. The calibration equations of the couples, $E = a + bt + ct^2$, were computed from the observed electromotive forces, E , at three temperatures, indicated by heavy type in the table. The observations were reduced in two ways. In the first column under each thermocouple are given the computed values of the freezing point of the silver-copper eutectic alloy, and of silver, when the freezing points of zinc, antimony, and copper are taken as calibration temperatures. These are the fixed points often employed for the standardization of thermocouples. In the second column under each thermocouple are given the computed values of the freezing point of antimony and of silver, when the freezing points of zinc, Ag-Cu eutectic, and copper are taken as calibration temperatures.

The observations show that the thermocouples calibrated at three temperatures are in excellent agreement at the intermediate temperatures, the average deviation from the mean of their indications being about $0^{\circ}.2$, and the maximum deviation $0^{\circ}.5$. Furthermore, the temperature scale defined by these couples calibrated at three temperatures (as given by the scale of the platinum thermometer) is in agreement at the intermediate temperatures with the scale of the platinum thermometer, as defined by the Callendar equation, to within $0^{\circ}.3$.

The method of comparing these temperature scales here adopted, viz, the determination of the freezing points of a number of metals, is capable of great precision and has some decided advantages over the usual method of intercomparison even in a most carefully compensated electric furnace. Harker,⁸⁸ using the latter method, finds the gas, thermoelectric, and resistance scales to agree to within 2° on the average, in the range 450° to 1000° .

VII. THE PALLADIUM THERMOMETER.

It was thought of sufficient interest to include measurements made with a thermometer of pure palladium, as this would furnish evidence as to the generality of the Callendar formula for pure metals.

The constants of such a thermometer, 1787G, calibrated in the same way as the platinum thermometer, are given in Table 1; the individual observations on the freezing points of the metals in the several tables 3 to 10; the behavior of the constants R_0 , FI and c in Table 22; and a comparison of the results obtained with the platinum and palladium thermometers at the several freezing points in Tables 26 and 27.

It is evident that the palladium thermometer by no means satisfies the Callendar equation at the higher temperatures. It follows that this formula is not general in its applicability to metallic resistance thermometers, even for metals of the platinum group, in spite of the fact that for thermometers of pure platinum it may be used for extrapolation to at least 1100° and give results in excellent agreement with the gas scale.

Between 0° and nearly 600° , however, the Callendar equation gives, for the palladium thermometer, temperatures agreeing with those found by the platinum thermometer to within 1° , and for the greater part of this range to within 0.5° .

The Callendar formula, modified by making $\delta = a + bt$, gives better values for temperatures obtained with the palladium thermometer than the simple Callendar formula (see Table 15). Dickson's equation (p. 177) does not satisfy the observations with palladium at all.

An independent equation for the temperature-resistance relation of palladium to 1100° appears to require five constants, so that it was considered impracticable to indicate such an equation.

The behavior of palladium as regards changes in its constants, due to heating (to 1100°) is about the same as that of platinum wire having $c = 0.0021$ and $\delta = 1.57$, see Tables 20 and 22.

Another thermometer of palladium, Pd II, was made with metal from another source; but its calibration constants, Table 1, as well as the thermoelectric behavior of the wire of which it was made, indicated a less pure material, and no further measurements were made with it.

VIII. THE BOILING POINT OF SULPHUR.

Regnault⁴ made a series of observations on the relation between pressure and temperature of the vapor of sulphur for pressures ranging from 250 to 3000 mm. As only two of these observations were in the neighborhood of atmospheric pressure, various values for the S. B. P. have been deduced from these observations by different experimenters who have endeavored to utilize Regnault's results.* As these determinations were carried out in a thick-walled iron-tube boiling-point apparatus heated in a furnace, the observed temperatures were too high, due to superheating of the sulphur vapor, a fact which was recognized by Regnault.

After an extensive series of experiments with an iron-tube S. B. P. apparatus, Callendar and Griffiths¹⁵ were finally led to adopt a glass-tube apparatus. The sulphur was boiled in the outer tube of the well known Victor Meyer vapor-density apparatus, about 4 cm diameter, 40 cm long, and having a spherical enlargement at the lower end. The tube was protected against heat loss by a suitable asbestos covering extending from the center of the bulb to within about 4 cm of the upper end. To prevent superheating of the sulphur vapor, the heating by the Bunsen flame was limited to the lower half of the spherical bulb containing the sulphur, the depth of which was such that it somewhat more than filled the spherical enlargement. The supply of heat was regulated so that the line of condensation of the sulphur vapor remained fixed a short distance (about 2 cm) above the asbestos covering. A series of instructive experiments carried out by these observers showed that in order to find the temperature of the vapor, or at least to attain a definite and reproducible temper-

* By the boiling point of sulphur, often abbreviated to S. B. P., will be understood the temperature on the scale of the gas thermometer of the saturated vapor of sulphur at standard atmospheric pressure (i. e., 760 mm of mercury at 0° and at latitude 45°, sea level), as measured in a definite form of boiling-point apparatus and under carefully specified conditions.

Using the two observations nearest atmospheric pressure, Callendar and Griffiths¹⁵ deduce from Regnault's data 447°.48 for S. B. P. Weinhold⁷ deduces from Regnault's empirical formula 448°.4. From a recomputation of these observations Ramsay¹⁰ gives 448°.34. Taking the four observations of Regnault nearest to atmospheric pressure (467.45 mm to 1308.54 mm), Chappuis and Harker⁶¹ deduce from these data 447°.51 as the S. B. P.

ature, it is not sufficient merely to insert a thermometer in the vapor. They found that when the platinum thermometer was immersed in the sulphur vapor, unprotected by any radiation screens it read about $1^{\circ}.2$ lower than when thoroughly shielded, owing to cooler liquid sulphur condensed in the upper parts of the tube running down over the bulb of the thermometer, and to loss of heat by radiation to the cooler walls of the tube. The effect of adding the umbrella (see Plate VIII) above the coil of the thermometer was to raise the observed temperature about $0^{\circ}.3$, and the effect of adding a glass-tube radiation screen and two perforated asbestos disks below was to produce a further rise of about $0^{\circ}.5$. Adding a second coaxial glass-tube radiation screen produced an additional rise of about $0^{\circ}.2$. Replacing the inner glass tube with a brass tube gave a further rise in temperature. In the final form of radiation screen used by these observers, shown in Plate VIII, the inner glass tube is covered with platinum foil. With this apparatus in its final form, these observers found for the *S. B. P.* $444^{\circ}.53$ at standard atmospheric pressure, on the scale of the constant pressure air thermometer, the platinum thermometers used in this determination having been previously compared with the constant pressure air thermometer.

A brief note²⁵ on the results of a determination of the *S. B. P.* made at the Physikalisch-Technische Reichsanstalt, by means of several mercurial thermometers that had been compared with a constant volume porcelain bulb air thermometer, gave for the *S. B. P.* $444^{\circ}.5$. The temperature scale of the air thermometer on which this work is based was shown to be in satisfactory agreement with the temperature scale of the air thermometer of Wiebe and Böttcher.

Chappuis and Harker⁶¹ determined the boiling point of sulphur by means of two platinum thermometers in an apparatus similar in all essential details to that used by Callendar and Griffiths, using as a radiation screen a perforated asbestos cone, similar to that used by Heycock and Neville,²⁷ as shown in Plate VIII. These thermometers had been previously compared with a constant volume nitrogen thermometer, having a bulb of Berlin porcelain. This work gave for the *S. B. P.* $445^{\circ}.2$. In this work the expan-

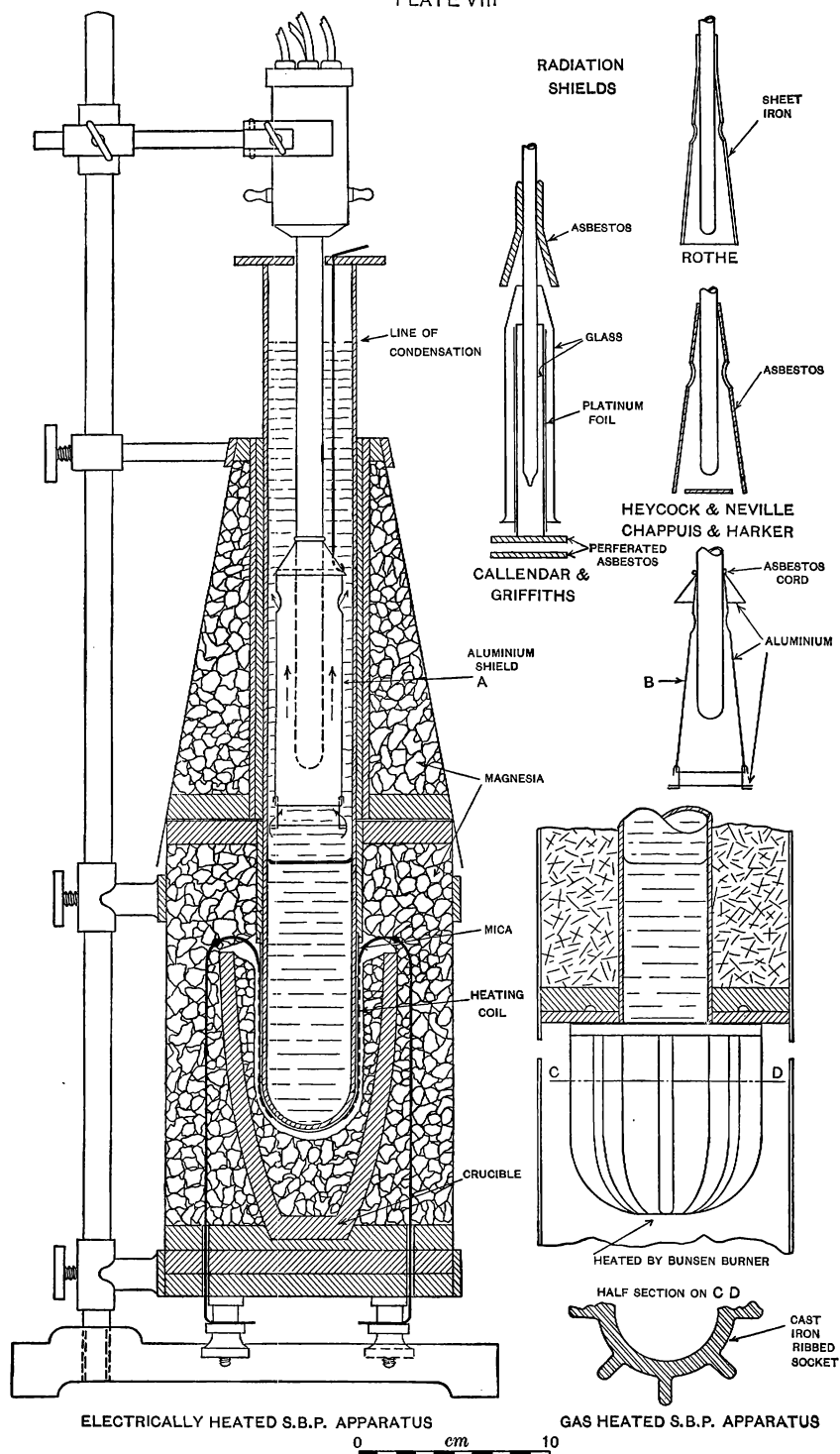
sion of the porcelain was determined only for the interval 0° to 100° , and the value at higher temperatures was determined by extrapolation. Subsequently, after Holborn and Day⁵³ had determined the expansion of Berlin porcelain at high temperatures, Chappuis⁷⁰ recomputed the results previously found in collaboration with Harker, and was led to the value *S. B. P.* = $444^{\circ}.7$ on the scale of the constant volume nitrogen thermometer.

Holborn⁵² compared several thermometers of pure platinum wire with a constant volume nitrogen thermometer in the interval 300° to 500° and found that the results of observation were in agreement with the Callendar equation to within $0^{\circ}.1$, the value of δ being 1.489 . While Holborn did not use these thermometers in an actual determination of the *S. B. P.*, it is evident that they would give $444^{\circ}.55$ for the *S. B. P.*, since platinum of the same degree of purity as that used in these thermometers has a $\delta = 1.50$, when this constant is determined by calibration in ice, steam, and sulphur, and when the boiling point of sulphur is taken as $444^{\circ}.70$.

Rothe⁸⁰ determined the boiling point of sulphur in an electrically heated *S. B. P.* apparatus similar to that described and used by the authors in the present work, using for this purpose four gas-filled mercurial thermometers, the relation of the scales defined by these thermometers to the standard gas scale of the Physikalisch-Technische Reichsanstalt being known. The correction for emergent stem of the thermometer, which amounted to $1^{\circ}.5$ to 3° , was determined by means of a Mahlke fadenthermometer. A conical radiation screen of thin sheet iron, as shown in Plate VIII, was used. Rothe found for the *S. B. P.* $444^{\circ}.7$ on the constant volume scale of the air thermometer. With two thermocouples that were calibrated on the scale of the Holborn and Day constant volume (*Pt-Ir* bulb) nitrogen thermometer, the standard high temperature scale of the Reichsanstalt, but which was not primarily intended for the highest attainable accuracy at the lower range of temperatures, Rothe found for the *S. B. P.* $445^{\circ}.0$.

Variation of *S. B. P.* with Pressure.—The following relations for the variation in temperature of the *S. B. P.* with pressure have been used:

PLATE VIII



$$t = t_{760} + 0.082 (p - 760) \dots\dots\dots (1)$$

$$t = t_{760} + 0.088 (p - 760) \dots\dots\dots (2)$$

$$t = t_{760} + 0.090 (p - 760) + 0.0002 (p - 760)^2 \dots\dots\dots (3)$$

$$t = t_{760} + 0.0912(p - 760) - 0.000042 (p - 760)^2 \dots\dots\dots (4)$$

$$t = t_{760} + 0.0904(p - 760) - 0.0000519 (p - 760)^2 \dots\dots\dots (5)$$

Equations (1) and (2) were deduced from Regnault's observations, the former by Griffiths from observations including a wide range of pressures, and the latter by Chappuis and Harker from the observations nearest atmospheric pressure. Equation (3) was computed by Chree⁵⁷ from a large number of observations of the steam points of platinum thermometers carried out at different times at Kew Observatory, the total range of barometric pressure being, however, only about 30 mm. Equations (4) and (5), the former given by Holborn and Henning¹³⁰ and the latter by the Harker and Sexton¹²⁷, are the results of experiments especially planned with a view to the determination of the temperature-pressure relation, and are, therefore, to be preferred. These equations, which are in very close agreement, were determined from experiments over a range of 100 mm above and below standard atmospheric pressure.

S. B. P. Tubes.—The construction of the electrically heated S. B. P. apparatus in which most of our observations were made is shown in sectional drawing in Plate VIII. Two forms of gas-heated tubes were also tried, one a Victor Meyer tube similar to that used by Callendar and Griffiths, in which the heating was done by a Bunsen flame applied directly to the lower part of the spherical enlargement at the bottom of the tube, the other in which the lower end of the glass tube was inclosed by the iron casting shown in Plate VIII, the heat being supplied to this iron jacket by a Bunsen flame.

The order of accuracy with which the S. B. P., observed in this way, serves to fix a definite and reproducible temperature, and the precautions that are necessary, will be seen from the following experiments which are briefly described:

Gas and Electric Heating.—A series of measurements made alternately with these three S. B. P. tubes,* using the same ther-

* We have had tubes which would last for a long time; at other times the cracking of the tubes was very annoying, and required frequent rebuilding of the apparatus due to the sulphur coming into contact with the heating coil. Heating coils of nickel, constantan, etc., were used at different times, but a heating coil of thin platinum ribbon was finally adopted. For these reasons the gas-heated type would probably prove more satisfactory for ordinary laboratory use.

monometer and the same radiation shields, has shown that the tubes serve to define practically the same temperature, the S. B. P. observed in the two gas-heated tubes being $0^{\circ}.03$ and $0^{\circ}.04$ higher, respectively, than in the electrically heated tube.

Superheating.—To avoid superheating of the sulphur vapor, the depth of liquid sulphur should be such that its upper surface is some distance (several cms at least) above the upper end of the electric heating coil or that portion of the glass tube in contact with the heating surface (flame or hot gases). For example, with an electric heating coil 7 cm high and with a depth of 8 cm of sulphur, there were still evidences of superheating for a distance of some 7 cm above the surface of the liquid, the superheating being about $0^{\circ}.1$, $0^{\circ}.05$, and $0^{\circ}.01$, respectively, at distances of 4.5, 5.5, and 6.5 cm above the surface. With a depth of only 3.5 to 4 cm of liquid sulphur in the tube, the observed superheating at 6.5, 10, 17, and 30 cm distance from the liquid surface (in a tube 55 cm long) was 20° , $8^{\circ}.3$, $2^{\circ}.5$, and $1^{\circ}.1$, respectively. With a depth of 12 cm of sulphur the maximum variation in temperature, when taken as hereafter described, for the interval 4 cm above the surface to 20 cm above the surface was only about $0^{\circ}.02$ or $0^{\circ}.03$. When naphthalene is boiled in the electrically heated tube, the observed boiling point (with thermometer No. 479) for 8 and 16 cm depth of liquid in the tube was identical (viz, $218^{\circ}.01$), while when the depth of liquid was only 4 cm the observed boiling point was $0^{\circ}.31$ higher, showing the effect of superheating due to the contact of the vapor with the portion of the wall heated by the coil above the surface of the liquid.

By lowering the thermometer, with its attached radiation shields, into the liquid sulphur, it was found that the temperature of the liquid was some $5^{\circ}.4$ higher than the temperature observed in the vapor.

Radiation Shields.—We have repeated most of the experiments of Callendar and Griffiths on the effects of shielding the thermometer in the sulphur vapor and have made many additional experiments in order to determine the conditions to be specified in order to reproduce the S. B. P. with the highest accuracy. The two types of radiation shields used in this work are shown in Plate

VIII. The temperatures observed with these two shields were very nearly identical, *A* giving a temperature very slightly lower than *B* (about $0^{\circ}.02$). With the lower disk of *A* removed, which gives the type of shield used by Rothe, the observed temperature was very slightly lower ($0^{\circ}.015$). A shield exactly similar to that used by Callendar and Griffiths was also tried. It gave a temperature about $0^{\circ}.03$ or $0^{\circ}.04$ higher than the shield *B*. The effect of adding another aluminium cylinder around *B* was to give a temperature about $0^{\circ}.03$ higher than the single cylindrical shield *B*. If all radiation shields are dispensed with the thermometer reads about $1^{\circ}.3$ low, and if the small cone alone is used above the coil, the thermometer reads about 1° low.

Temperature Distribution within S. B. P. Tube.—To determine whether such circumstances as distance of thermometer coil from liquid surface, from line of condensation, etc., has any effect on the observed S. B. P., the thermometer, together with its shield, was moved up and down along the axis of the S. B. P. tube and observations were taken with the coil of the thermometer at various distances from the liquid surface (from about 7 cm from center of coil to surface of sulphur to a position where the top of the radiation shield is just below the condensation line on the S. B. P. tube). With the thermometer in shield *B*, mounted so that the lower end of the porcelain tube was in the plane of the base of the surrounding aluminium cylinder, the temperature remained constant (within $0^{\circ}.01$) until the top of the shield was about 1 cm below the line of condensation of the sulphur, in which position it read $0^{\circ}.02$ low, the observed temperature falling off very rapidly upon further raising the thermometer. A similar experiment was also tried with the thermometer coil mounted about 4.5 cm farther up in the shield with the same result. Notwithstanding the excellent agreement ($0^{\circ}.01$ or $0^{\circ}.02$) in the observed temperatures when the coil of the thermometer was placed in these two different positions within the radiation shield, thermocouples (of *Pt*, *Pt-Ir*, and *Pt*, *Pt-Rh*), used to explore the distribution of temperature within that portion of a porcelain tube (similar to the thermometer tube) under the radiation shield, showed very appreciable point-to-point variations within this region, amounting, as will be seen later, to as much as $0^{\circ}.5$.

While the conical shield *A* gives a temperature only about $0^{\circ}.02$ lower than the cylindrical shield *B*, it must be placed farther down in the sulphur vapor, well below the line of condensation. With the conical shield *A* there is a marked drop in temperature when the thermometer, with its shield, is raised, which is noticeable when the top of the shield is still some 4 or 5 cm below the condensation line. By putting an additional small umbrella above the shield *A* this effect disappears, and the thermometer, with its shield, can be raised so that the top of the shield is within 1 cm of the condensation line without producing a change in the observed temperature exceeding $0^{\circ}.01$.

Temperature Distribution within Radiation Shield.—As has been stated, the observed resistances of a platinum thermometer when mounted in the two different positions within the shield, as already described, were almost identical. Notwithstanding this fact, when the temperature distribution inside a porcelain tube, similar to the thermometer tube, was explored with a thermocouple, it gave evidence of very appreciable point-to-point variations of temperature of the order of $0^{\circ}.4$ or $0^{\circ}.5$. If these are true variations in temperature, one would expect to find a different resistance for the coil of the platinum thermometer when mounted in different positions within the shield, especially as the thermometer set-up was sensitive to well within $0^{\circ}.01$. It is possible that a part of the variation observed with the couples is apparent rather than real, due, perhaps, to the heterogeneity of the wires of the couple, which would give slightly different electromotive forces with variable depth of immersion in the heated region. Heræus couples of *Pt*, *Pt-Rh*, and of *Pt*, *Pt-Ir*, however, both gave about the same temperature-distribution curve. Further experiments are contemplated with a special resistance thermometer, having a very short coil of a few millimeters of fine (Wollaston) platinum wire.

On the Uncertainty in the Accepted Value of the S. B. P.—On account of the satisfactory agreement of the determinations of the sulphur boiling point by different observers using different gas thermometers it has been very generally assumed that this temperature was known to $0^{\circ}.1$ or $0^{\circ}.2$. A consideration of some recent work with the gas thermometer, however, throws some

doubt on this conclusion and would seem to show that the uncertainty in the present accepted value of the S. B. P. ($444^{\circ}.5$ to $444^{\circ}.7$) may be as much as 1° .

In the present investigation the value assumed for the S. B. P.—viz, $444^{\circ}.70$ —gives $419^{\circ}.37$ for the F. P. of zinc, as determined by the platinum resistance thermometer. Day and Clement¹²⁵ have recently published the preliminary results of a series of elaborate and painstaking measurements with a constant volume nitrogen thermometer and find for the F. P. of zinc $418^{\circ}.5$. If this temperature were taken as one of our known calibration temperatures for the platinum thermometer, the other two temperatures being that of ice and of steam, it would give for the S. B. P. $443^{\circ}.7$ —that is, 1° lower than the generally accepted value. The desirability of further careful work with the gas thermometer is therefore evident. In this connection it is of interest to note the value of the S. B. P. recently found by Eumorfopoulos,¹²⁸ using the Callendar form of compensated constant pressure air thermometer, viz, $443^{\circ}.6$. The uncertainty in the coefficients of expansion used for the glass bulb of his air thermometer is such, however, that this value of the S. B. P. may be in error by more than 1° , as is recognized by both Messrs. Eumorfopoulos and Callendar.¹²⁴

If the true temperature of the vapor of sulphur, boiling at standard atmospheric pressure and measured under the conditions specified by the several investigators who have determined this point, should ultimately be found to be $443^{\circ}.7$ instead of $444^{\circ}.7$, as we have taken it, it would result in changing the value of δ from 1.505 to 1.447 and in lowering the resulting values of the freezing points of Sn, Cd, Pb, Zn, Sb, Ag-Cu, Ag, Cu, by $0^{\circ}.19$, $0^{\circ}.44$, $0^{\circ}.47$, $0^{\circ}.87$, $2^{\circ}.3$, $3^{\circ}.9$, $6^{\circ}.5$, $8^{\circ}.8$, respectively.

The Callendar method of calibration in ice, steam, and sulphur would not then be applicable for extrapolation to high temperatures, as it would give for the F. P. of copper $1,074^{\circ}$ instead of $1,083^{\circ}$ and the experimental evidence available makes it reasonably certain that the latter value can hardly be in error by more than 2° or 3° .

While the agreement of the scale of the platinum thermometer with the Holborn-Day gas scale is remarkable when we consider

that the former scale is dependent on an extrapolation from the boiling point of sulphur, yet on account of the purely empirical nature of the Callendar equation this agreement can not be used as an argument in favor either of the correctness of the present accepted values of the freezing points of the metals or to decide between these two values of the S. B. P., although, taken in connection with the other corroborative experimental evidence previously referred to, the higher value of the S. B. P. ($444^{\circ}.70$) and the resulting temperature scale as used in this investigation appear the more probable.

S. B. P. on Thermodynamic Scale.—Berthelot¹³⁷ and Buckingham¹³⁸ have recently discussed the probable differences between the temperature scales of the constant volume and the constant pressure gas thermometer and the absolute thermodynamic scale. The S. B. P. found by Callendar and Griffiths ($444^{\circ}.53$ on the constant pressure air scale) and by Harker and Chappuis ($444^{\circ}.70$ on the constant volume nitrogen scale), when reduced in accordance with the conclusions of the above authors, is $444^{\circ}.9$ on the thermodynamic Centigrade scale. In some recent work with the platinum resistance thermometer, Holborn and Henning¹³⁰ have used the value $445^{\circ}.0$.

IX. GENERAL CONCLUSIONS.

a. Temperature Scales.

1. Temperatures defined by the resistance thermometer of pure platinum calibrated by Callendar's formulæ at 0° , 100° , and 444.70° , the boiling point of sulphur, agree with temperatures on the generally accepted gas scale in the range 0 to 1100° C, to within the degree of reproducibility of the latter.

Freezing Point of:	Sn	Cd	Zn	Sb	Ag ₃ -Cu ₂	Ag	Cu
Thermometers of Pure Pt..... { $c=0.0039$ { $\delta=1.50$	231°.90	321°.01	419°.37	630°.71	779°.20	960°.88	1083°.0
Thermometer of Impure Pt 1787A.. { $c=0.0021$ { $\delta=1.57$	231°.82	320°.95	419°.32	631°.25	780°.86	966°.21	1091°.9
Thermometer of Impure Pt 1787E.. { $c=0.0017$ { $\delta=1.80$	632°.65	784°.60	975°.22	1106°.2

2. Thermometers made of impure platinum, and calibrated in the same way, do not define the same temperature scale as that of pure platinum, the divergence increasing with the impurity of the metal, as shown in the preceding table.

3. If thermometers of impure platinum are calibrated at a fourth temperature, such as the silver freezing point, and the difference coefficient of the Callendar formula is written $\delta = a + bt$, then the scale so defined by these thermometers is brought into very close agreement with that of the thermometer of pure platinum, as follows:

Freezing Point of:	Sn	Cd	Sb	Ag ₃ -Cu ₂	Ag	Cu
Thermometers of Pure Pt..... ($\delta=1.50$)	231°.90	321°.01	630°.7	779°.2	960°.9	1083°.0
Thermometer of Impure Pt 1787A.....	231°.89	321°.05	630°.5	779°.1	(961°.0)	1083°.3
Thermometer of Impure Pt 1787E.....			630°.8	779°.3	(961°.0)	1082°.3

For Thermometer 1787A, $\delta=1.608-0.00853 t$.

For Thermometer 1787E, $\delta=3.020-0.00327 t$.

4. The Dickson formula $(R + a)^2 = p(t + b)$ is not applicable to either platinum or palladium thermometers at high temperatures.

5. The palladium thermometer, calibrated by the Callendar method, gives temperatures agreeing to 1° with the scale of the platinum thermometer up to 600°. It requires a fourth degree equation to express its resistance-temperature relation from 0 to 1100°.

6. The following table of freezing points represents the scale of the resistance thermometer of pure platinum as determined from this investigation, using only the purest metals:

Metal	Sn	Cd	Pb	Zn	Sb	Al	Ag ₃ -Cu ₂	Ag	Cu	Cu-Cu ₂ O
Freezing Point.....	231°.92	321°.01	327°.43	419°.37	630°.71	658°.0	779°.20	960°.9	1083°.0	1063°.2
Average deviation for purest metal.....	.02	.02	.10	.02	.30	.5	.25	.20	.5	.3
Reproducibility factor.....	.06	.26	.10	.06	2.3	1.2			1.0	

The average deviation of a single observation is a measure of the precision attained when two or more thermometers were used in a single metal sample. The reproducibility factor is the average deviation of the metals, designated as C. P. or better, from their mean F. P.

7. Thermocouples of *Pt*, 90 *Pt*-10 *Rh* and *Pt*, 90 *Pt*-10 *Ir*, calibrated at three temperatures in terms of the formula $E = a + bt + ct^2$, define a temperature scale in the interval 200° to 1100°, which is in agreement with the platinum resistance scale at intermediate temperatures to within 0°.3.

b. Behavior of Platinum Thermometers.

8. When used at high temperatures the constants of platinum thermometers undergo gradual changes, which necessitate frequent recalibration in work of high precision. The use of thermocouples for work above the F. P. of silver is therefore much less laborious.

9. These changes in thermometric constants are greatly reduced but not completely eliminated, by annealing the thermometers for at least several hours, at a temperature higher than that at which they are subsequently to be used.

10. For thermometers of pure platinum, the changes in their constants after the thermometers have been annealed, are very much less than for those of impure platinum; thus, R_0 changes only by a few tenths of a degree for pure platinum but by several degrees for impure.

11. These changes are least for pure platinum wires of large diameter and suspended free from strains. For such thermometers, made of wire 0.6 mm in diameter, R_0 changed by only a few tenths of a degree after repeated and continued heatings at 1200° to 1300°.

12. For impure platinum wire the effect of high temperatures is to decrease R_0 and to increase the fundamental coefficient, c , that is, the effect is as if the wire became purer, possibly because of the evaporation of impurities, for example, iridium. If the platinum is pure, the slight changes indicate a contamination of the wire and the effect of strains, as is evidenced by decrease in c and the increase in R_0 . The total change observed is a resultant of the effects of strains, of annealing, and of contamination and purification. In the use to which these thermometers were subjected in this investigation, the changes in the value of δ were almost within the limits of experimental error. For impure platinum there is evidence of a slight decrease in this constant.

c. Calibration Temperatures.

13. The freezing points of the pure metals are the most satisfactory calibration temperatures, being superior to boiling points in constancy, ease of manipulation, and reproducibility. Unless

a very slow rate of heating is used, melting points are less satisfactorily fixed and apparently slightly higher than freezing points.

14. The temperature of the separation of the silver-copper eutectic is an entirely satisfactory fixed point at $779^{\circ}.2$, when the alloy is of the eutectic composition, or about 72Ag-28Cu. The eutectic point of copper-cuprous oxide is also a satisfactory calibration temperature at $1063^{\circ}.2$, when proper experimental precautions are taken to maintain the eutectic composition.

d. The Boiling Point of Sulphur.

15. The temperature defined by the vapor of sulphur boiling at known pressure and under carefully specified conditions is a most satisfactory fixed point, reproducible to a few hundredths of a degree. These conditions are briefly:

(a) Sulphur to be boiled in a glass tube 4 to 5 cm diameter and of sufficient length to insure ample depth of immersion in the vapor (tube about 40 cm satisfactory);

(b) Sufficient depth of liquid sulphur to avoid superheating;

A criterion for the above conditions is absence of change in the observed resistance when the thermometer with its attached radiation shield is moved some centimeters up and down the tube;

(c) Surrounding coil of thermometer with a suitable radiation shield to avoid the cooling effects of the condensed sulphur running down the stem of the thermometer and of radiation to the cooler parts of the apparatus.

16. Both gas and electrically heated S. B. P. tubes were used; the former gave a temperature $0^{\circ}.03$ or $0^{\circ}.04$ higher than the latter.

17. The determinations of the S. B. P. by different observers is discussed at length. The determinations of these observers are in satisfactory agreement, the values found ranging, with one exception, from $444^{\circ}.5$ to $444^{\circ}.7$. The bearing of some recent work with the gas thermometer on the value of the S. B. P. is considered, which shows that the above value may be too high. The evidence, however, is not sufficient as yet to warrant any change in the generally accepted and widely used value of the S. B. P.

We take great pleasure in acknowledging the assistance rendered by Messrs. J. J. Crowe and H. A. Ginsburg, both in the experimental work and in the computations involved in this investigation.

WASHINGTON, June 26, 1909.

X. APPENDIX.*

TABLE 2.

Heating Effect of Measuring Current.

[Thermometer 1787C.]

Amperes	R ₀	R ₁₀₀	R _{444.33}	F. I.	pt _s	δ
2.50x10 ⁻³	3.48160	4.82275	9.13220	1.34115	421°.33	1.503
5.00x10 ⁻³	3.48164	4.82277	9.13220	1.34113	421 .33	1.503
10.00x10 ⁻³	3.48174	4.82287	9.13213	1.34113	421 .31	1.505
25.0 x10 ⁻³	3.48286	4.82401	9.13308	1.34115	421 .36	1.505
50.0 x10 ⁻³	3.48705	4.82819	9.13608	1.34114	421 .21	1.511
100.0 x10 ⁻³	3.50373	4.84546	9.14832	1.34173	420 .69	1.545

Amperes	Heating of Pt Coil above Surrounding Temperature		
	ΔT ₀	ΔT ₁₀₀	ΔT _s
2.50x10 ⁻³	0°.001±	0°.001±	
5.00x10 ⁻³	.004	.002	
10.00x10 ⁻³	.011	.010	
25.0 x10 ⁻³	.095	.095	0°.067
50.0 x10 ⁻³	.41	.41	.29
100.0 x10 ⁻³	1 .65	1 .69	1 .20

R at F. P. of Silver		pt _{A_R}	δ	t _{A_R}
0.100 amp.	0.0100 amp.			
14.7079		835°.06	1.545	963°.64
	14.6995	836 .43	1.504	960 .83
	14.7009	836 .54	1.504	960 .98
14.7104		835 .24	1.545	963 .89

*The following tables are in the text: 1 on p. 156, 11 on p. 173, 12 on p. 174, 15 on p. 177, 25 on p. 181.

TABLE 3.
Freezing Point of Tin.

Date	Therm. No.	R ₀	FI	δ	R _{Sn}	pt _{Sn}	t	Remarks
Feb. 5, '09	1787C., P. T.	3.47977	1.33997	1.507	6.52466	227°.24	231°.85	F. P., E. & A., '05, D. C.
May 13, '09	"	3.48164	1.34105	1.505	6.52906	227°.24	231°.84	{ " Baker, '09, "
		3.48177	1.34105	1.505		227°.23	231°.83	
Mar. 17, '09	1787F., W. B.	2.85456	1.11265	1.500	5.38326	227°.27	231°.85	" B. & A., " "
"	"	2.85456	1.11265	1.500	5.38420	227°.35	231°.94	" (K), " A. C.
"	"	2.85456	1.11265	1.500	5.38397	227°.33	231°.92	" " " "
Mar. 18, '09	" P. T.	2.84881	1.11084	1.508	5.37370	227°.30	231°.91	" " " "
"	"	2.84881	1.11084	1.508	5.37360	227°.29	231°.90	" " " "
Jan. 30, '09	1787A., P. T.	21.3347	4.4077	1.572	31.3388	226°.97	231°.77	" B. & A., '05, D. C.
"	"	21.3347	4.4077	1.572	31.3427	227°.06	231°.87	M. P., " " "
Feb. 5, '09	"	21.3347	4.4077	1.572	31.3402	227°.00	231°.80	F. P., E. & A., " "
Mar. 18, '09	"	21.0734	4.4213	1.565	31.1129	227°.07	231°.85	" (K), '09, A. C.
Mar. 17, '09	1787G., W. B. Palladium.	1.87759	0.63158	2.890	3.28485	222°.81	231°.62	" " " "

F. P. = Freezing point. M. P. = Melting point. D. C. = Dixon graphite crucible. A. C. = Acheson graphite crucible. E. & A. = Eimer and Amend.
B. & A. = Baker and Adamson (K) = "Kahlbaum." P. T. and W. B. indicate method of measurement: Potentiometer, or Wheatstone Bridge.

Where two values are given in brackets, the upper refers to the calibration before the F. P. or M. P. determination and the resulting value of the F. P. or M. P., and the lower to the calibration after and the corresponding value of the F. P. or M. P. computed therefrom.

TABLE 4.
Freezing Point of Cadmium.

Date	Therm. No.	R _o	FI	δ	R _{Cd}	pt _{Cd}	t	Remarks
Feb. 27, '09	1787C., P. T.	3.47967	1.34094	1.505	7.64123	310°.35	321°.03	F. P., (K), '09, A. C.
"	"	3.47967	1.34094	1.505	7.64015	310°.27	320°.94	M. P., " " "
May 12, '09	"	3.48164	1.34105	1.505	7.63632	309°.81	320°.44	F. P., Baker, '09, D. C.
		3.48177	1.34105	1.505		309°.80	320°.43	
"	"	3.48164	1.34105	1.505	7.63580	309°.77	320°.40	" B. & A., '09, "
		3.48177	1.34105	1.505		309°.76	320°.39	
"	"	3.48164	1.34105	1.505	7.63563	309°.76	320°.39	" " " "
		3.48177	1.34105	1.505		309°.75	320°.38	
Mar. 16, '09	1787F., W. B.	2.85456	1.11265	1.501	6.30820	310°.40	321°.05	" (K), " "
Feb. 27, '09	1787A., P. T.	21.1732	4.4165	1.568	34.8576	309°.85	320°.97	" " " "
"	"	21.1732	4.4165	1.568	34.8559	309°.81	320°.93	M. P., " " "
Mar. 16, '09	1787G., W. B. Palladium.	1.87759	0.63158	2.890	3.77290	300°.09	320°.52	F. P., " " "

TABLE 5.
Freezing Point of Lead.

Date	Therm. No.	R ₀	FI	δ	R _{Pb}	pt _{Pb}	t	Remarks
Mar. 25, '09	1787F., W. B.	2.85455	1.11265	1.500	6.37505	316°.41	327°.59	F. P., Baker, Ag free, test.
"	"	2.85455	1.11265	1.500	6.37501	316 .40	327 .58	F. P., Baker, Ag free, test.
Apr. 22, '09	1787C., P. T.	3.48164	1.34105	1.505	7.72003	316 .05	327 .24	F. P., (K), '09.
"	"	3.48164	1.34105	1.505	7.72043	316 .08	327 .27	" " "
May 13, '09	"	{ 3.48164	1.34105	1.505	{ 7.72278	{ 316 .26	327 .47	{ M. P., B. & A. test, Ag free, D. C.
		{ 3.48177	1.34105	1.505		{ 316 .25	327 .46	
		{ 3.48164	1.34105	1.505	{ 7.72278	{ 316 .26	327 .47	{ F. P., B. & A. test, Ag free, D. C.
"	"	{ 3.48177	1.34105	1.505		{ 316 .25	327 .46	

TABLE 6.
Freezing Point of Zinc.

Date	Therm. No.	R _o	F I	δ	R _{Zn}	pt _{Zn}	t	Remarks
Nov. 19, '08	1787A., P. T.	21.3476	4.4067	1.571	38.9003	398°.32	419°.36	Fairly rapid freeze.
20, '08	"	21.3476	4.4067	1.571	38.8972	398 .25	419 .28	" " "
Dec. 15, '08	"	21.3361	4.4071	1.571	38.8887	398 .28	419 .31	Slow freeze.
"	"	21.3361	4.4071	1.571	38.8880	398 .26	419 .29	" "
"	"	21.3361	4.4071	1.571	38.8880	398 .26	419 .29	" "
"	"	21.3361	4.4071	1.571	38.8886	398 .28	419 .31	" "
Dec. 5, '08	480 W. B.	2.62081	1.00503	1.551	6.62603	398 .52	419 .28	Rapid freeze.
"	"	2.62081	1.00503	1.551	6.62482	398 .40	(419 .16)	
Dec. 8, '08	479 W. B.	4.27342	1.64585	1.516	10.84077	399 .02	419 .32	Bottom of therm. 5 cm from bottom of cru- cible.
"	"	4.27340	1.64585	1.516	10.83975	398 .97	419 .27	
"	"	4.27338	1.64585	1.516	10.84036	399 .00	419 .30	
Dec. 9, '08	"	4.27323	1.64585	1.516	10.84046	399 .02	419 .32	
"	"	4.27323	1.64585	1.516	10.82075	397 .82	(417 .99)	3 cm from bottom.
"	"	4.27323	1.64585	1.516	10.83951	398 .96	(419 .26)	
Jan. 4, '09	1787C., P. T.	3.48779	1.34298	1.506	8.84889	399 .19	419 .36	All above dets. (K) Zn in A. C.
"	"	3.48779	1.34298	1.506	8.84902	399 .20	419 .37	

Jan. 21, '09	"	3.47971	1.33989	1.506	8.82960	399 .28	419 .46	B. & A., C. P. Zn in D. C.
22, '09	1787A., P. T.	21.3396	4.4076	1.571	38.8982	398 .37	419 .41	B. & A., C. P. Zn in D. C.
"	"	21.3396	4.4076	1.571	38.8975	398 .36	419 .40	B. & A., C. P. Zn in D. C.
Jan. 23, '09	"	21.3396	4.4076	1.571	38.8964	398 .33	419 .37	Ⓚ Zn in D. C. In use since 1903 for testing couples.
"	1787C., P. T.	3.47971	1.33989	1.506	8.82896	399 .23	419 .40	
"	"	3.47971	1.33989	1.506	8.82896	399 .23	419 .40	
Jan. 25, '09	1787A., P. T.	21.3395	4.4076	1.571	38.8942	399 .28	419 .32	Ⓚ Zn in D. C. In use since 1907.
"	1787C., "	3.47964	1.33989	1.506	8.82841	399 .19	419 .36	
"	"	3.47964	1.33989	1.506	8.82843	399 .19	419 .36	
Jan. 26, '09	"	3.47970	1.33991	1.506	8.82713	399 .09	419 .25	E. & A., C. P. Zn in D. C. Mixture of 1903 and 1905 samples.
"	1787A., P. T.	21.3395	4.4076	1.571	38.8915	398 .22	419 .25	

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Platinum Thermometry at High Temperatures.

TABLE 7.
Freezing Point of Antimony.

Date	Therm. No.	R _o	FI	δ	R _{eb}	p _{eb}	t	Remarks
Nov. 21, '08	1787A., P. T	21.3384	4.4067	1.571	46.8340	578°.56	631°.24	(K) 1908, in A. C.
Jan. 26, '09	"	21.3384	4.4073	1.572	46.8334	578 .47	631 .17	" " " " "
Jan. 28, '09	"	21.3360	4.4077	1.572	46.8367	578 .55	631 .26	(K) 1905, " D. C.
Jan. 29, '09	"	21.3348	4.4077	1.572	46.8384	578 .61	631 .34	(K) 1907, " "
Nov. 21, '08	1787C., P. T	3.4831	1.34163	1.506	11.2688	580 .32	630 .73	(K) 1908, in A. C.
"	"	3.4831	1.34163	1.506	11.2664	580 .14	630 .52	" " " "
Jan. 27, '09	"	3.4797	1.33994	1.507	11.2567	580 .40	630 .87	" " " "
Jan. 28, '09	"	3.4798	1.33994	1.507	11.2573	580 .44	630 .92	(K) 1905, " "
Jan. 29, '09	"	3.4798	1.33994	1.507	11.2566	580 .39	630 .86	(K) 1907, " "
Jan. 27, '09	479 W. B	4.2736	1.64580	1.508	13.8182	579 .94	630 .35	(K) 1908, " "
Feb. 1, '09	"	4.2735	1.64573	1.508	13.8175	579 .93	630 .34	" " " "
Feb. 3, '09	"	4.2735	1.64573	1.512	13.8208	580 .13	630 .74	(K) 1907, " "
Feb. 1, '09	1787 G. W. B	1.8750	0.62861	2.890	5.2405	535 .38	632 .82	(K) 1908, " "
"	" P. T	1.8725	0.62761	2.890	5.2332	535 .48	632 .99	(K) 1908, " "
Feb. 3, '09	" W. B	1.8741	0.62817	2.890	5.2386	535 .60	633 .16	(K) 1907, " "
Apr. 16, '09	1787E., P. T	2.92415	0.50480	1.802	5.81135	571 .95	632 .68	(K) 1907, " D. C. No cover.
		2.92382	0.50511	1.807		571 .66	632 .53	
		2.92415	0.50480	1.802	5.81178	572 .03	632 .78	" " "
		2.92382	0.50511	1.807		571 .75	632 .64	

Melting Point of Antimony.

Nov. 21, '08	1787A., P. T	21.3384	4.4067	1.571	46.8261	578°.38	631°.02	Ⓚ 1908, in A. C.
Jan. 28, '09	"	21.3354	4.4077	1.572	46.8286	578.38	631.07	Ⓚ 1905, " D. C.
Jan. 29, '09	"	21.3347	4.4077	1.572	46.8293	578.41	631.10	Ⓚ 1907, " "
Nov. 21, '08	1787C., P. T	3.4831	1.34163	1.506	11.2691	580.34	630.75	Ⓚ 1908, " A. C.
"	"	3.4831	1.34163	1.506	11.2707	580.46	630.90	" " " "
Jan. 27, '09	"	3.4797	1.33994	1.507	11.2556	580.32	630.78	" " " "
Jan. 29, '09	"	3.4798	1.33997	1.507	11.2556	580.30	630.75	Ⓚ 1907, " D. C.
Feb. 3, '09	1787 G. W. B	1.8741	0.62817	2.890	5.2379	535.49	633.00	" " " "
Feb. 1, '09	" P. T	1.8750	0.62861	2.890	5.2389	535.13	632.45	Ⓚ 1908, " A. C.

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Burgess.] Platinum Thermometry at High Temperatures.

TABLE 8.
Freezing Point of Silver.

Date	No. of Therm.	R _o	FI	δ	R _{Ag}	pt _{Ag}	t	t mean	Remarks
Mar. 11, '09	478 P. T	{ 5.16514 5.16618	{ 2.00011 1.99957	{ 1.507 1.503	{ 21.8904	{ 836°.22 836 .39	{ 960°.88 960 .65	{ 960°.76	A. C.
"	" "	{ 5.16514 5.16618	{ 2.00011 1.99957	{ 1.507 1.503	{ 21.8898	{ 836 .19 836 .36	{ 960 .84 960 .60	{ 960 .72	A. C.
June 11, '09	1787C., P. T	3.48174	1.34114	1.504	14.7009	836 .54	960 .98	"
"	" "	3.48174	1.34114	1.504	14.6995	836 .43	960 .83	"
Mar. 8, '09	1787 F. W. B	{ 2.85459 2.85456	{ 1.11274 1.11265	{ 1.503 1.501	{ 12.16260	{ 836 .49 836 .56	{ 960 .79 960 .66	{ 960 .72	"
"	" P. T	{ 2.84890 2.84881	{ 1.11077 1.11084	{ 1.504 1.508	{ 12.14004	{ 836 .46 836 .41	{ 960 .86 961 .25	{ 961 .06	"
"	" P. T	{ 2.84890 2.84881	{ 1.11077 1.11084	{ 1.504 1.508	{ 12.1401	{ 836 .46 836 .42	{ 960 .86 961 .26	{ 961 .06	" M. P.
"	" P. T	{ 2.84890 2.84881	{ 1.11077 1.11084	{ 1.504 1.508	{ 12.13953	{ 836 .41 836 .37	{ 960 .79 961 .20	{ 961 .00	"
"	" W. B	{ 2.85459 2.85456	{ 1.11274 1.11265	{ 1.503 1.501	{ 12.16214	{ 836 .45 836 .52	{ 960 .73 960 .61	{ 960 .67	"
"	479 W. B	4.28823	1.64513	1.513	18.0387	835 .83	961 .02	" Broke after F. P.

Mar. 11, '09	1787A., P. T	{ 21.0862 21.0734	{ 4.4203 4.4213	{ 1.568 1.565	{ 57.9980	{ 835 .05 835 .15	{ 966 .31 966 .10	{ 966 .20	" Impure Platinum.
"	" "	{ 21.0862 21.0734	{ 4.4203 4.4213	{ 1.568 1.565	{ 57.9982	{ 835 .06 835 .16	{ 966 .32 966 .11	{ 966 .22	" " "
Apr. 17, '09	1787E., P. T	{ 2.92382 2.92276	{ 0.50511 0.50517	{ 1.804 1.803	{ 7.07231	{ 821 .30 821 .41	{ 975 .30 975 .35	{ 975 .32	" " "
"	" "	{ 2.92382 2.92276	{ 0.50511 0.50517	{ 1.804 1.803	{ 7.07200	{ 821 .24 821 .36	{ 975 .22 975 .27	{ 975 .24	" " "
Mar. 15, '09	1787 G. W. B	{ 1.87869 1.87759	{ 0.63124 0.63158	{ 2.890 2.890	{ 6.52849	{ 736 .61 736 .39	{ 992 .74 992 .27	{ 992 .50	" Palladium.
"	" "	{ 1.87869 1.87759	{ 0.63124 0.63158	{ 2.890 2.890	{ 6.52861	{ 736 .63 736 .41	{ 992 .78 992 .29	{ 992 .54	" "
"	" P. T	{ 1.87498 1.87397	{ 0.63093 0.63084	{ 2.883 2.883	{ 6.52460	{ 736 .95 737 .21	{ 992 .14 992 .68	{ 992 .41	" "

TABLE 9.—Freezing Point of Copper.

Date	Therm. No.	R ₀	F. I.	δ	R _{Cu}	pt _{Cu}	t	t mean	Remarks
Feb. 19, '09	1787C., P. T	3.47967 3.47904	1.34058 1.34076	1.508 1.508	15.8439	922°.30 922 .23	1082°.77 1082 .69	1082°.7 ₃	Old Pot., D. C.
Feb. 19, '09	" "	3.47967 3.47904	1.34058 1.34076	1.508 1.508	15.8488	922 .67 922 .59	1083 .31 1083 .21	1083 .2 ₆	M. P. Old Pot., D. C., quite rapid.
Feb. 20, '09	" "	3.47904 3.47988	1.34076 1.34094	1.506 1.506	15.8538	922 .97 922 .78	1083 .43 1083 .17	1083 .3 ₀	K. elect., D. C.
"	" "	3.47904 3.47988	1.34076 1.34094	1.506 1.506	15.8558	923 .11 922 .93	1083 .63 1083 .38	1083 .5 ₀	M. P., K. elect., D. C., fast melt.
Mar. 1, '09	" "	3.47947 3.48041	1.34094 1.34080	1.505 1.508	15.8560	922 .97 923 .00	1083 .27 1083 .77	1083 .5 ₂	B. C. W., A. C.
Mar. 5, '09	" "	3.48041 3.48128	1.34080 1.34086	1.508 Not redet.	15.8565	923 .04 922 .93	1083 .85 1083 .68	1083 .7 ₆	B. C. W., A. C.
"	" "	3.48041 3.48128	1.34080 1.34086	Not redet. Not redet.	15.8570	923 .07 922 .97	1083 .89 1083 .74	1083 .8 ₁	B. C. W., A. C.
Mar. 6, '09	" "	3.48128 3.48156	1.34086 1.34104	Not redet. 1.509	15.8593	923 .14 923 .00	1083 .99 1083 .94	1083 .9 ₆	E. A., & Drops, '09, A. C. with cover.
"	" "	3.48128 3.48156	1.34086 1.34104	1.508 1.509	15.8591	923 .12 922 .98	1083 .97 1083 .91	1083 .9 ₄	E. & A., Drops, '09, A. C. with cover.
Feb. 24, '09	478 P. T	5.16070 5.16248	1.99969 2.00013	1.501 1.505	23.6006	922 .14 921 .85	1081 .47 1081 .65	1081 .5 ₆	K. elect., D. C.
"	" "	5.16070 5.16248	1.99969 2.00013	1.501 1.505	23.6011	922 .17 921 .87	1081 .50 1081 .67	1081 .5 ₈	M. P., K. elect., D. C.
Mar. 1, '09	" "	5.16248 5.16514	2.00013 2.00011	1.505 1.507	23.6120	922 .42 922 .29	1082 .46 1082 .58	1082 .5 ₂	B. C. W., A. C.

Mar. 1, '09	478 P. T	{ 5.16248 5.16514	{ 2.00013 2.00011	{ 1.505 1.507	{ 23.6021	{ 921 .92 921 .80	{ 1081 .74 1081 .88	{ 1081 .8 ₁	{ M. P. very rapid, A. C., B. C. W.
Feb. 24, '09	479 W. B	{ 4.27916 4.28715	{ 1.64497 1.64500	{ 1.505 1.514	{ 19.4283	{ 920 .94 920 .44	{ 1080 .33 1080 .99	{ 1080 .7 ₀	{ K. elect., D. C.
"	" "	{ 4.27916 4.28715	{ 1.64497 1.64500	{ 1.505 1.514	{ 19.4343	{ 921 .30 920 .80	{ 1080 .85 1081 .53	{ 1081 .2 ₀	{ M. P., K. elect., D. C.
Mar. 2, '09	" "	{ 4.28715 4.28823	{ 1.64500 1.64513	{ 1.514 1.513	{ 19.4484	{ 921 .66 921 .52	{ 1082 .76 1082 .40	{ 1082 .5 ₃	{ B. C. W., D. C.
Mar. 2, '09	1787F., W. B	{ 2.85344 2.85459	{ 1.11239 1.11274	{ 1.494 1.509	{ 13.1241 ₄	{ 923 .30 922 .91	{ 1082 .06 1083 .81	{ 1082 .9 ₃	{ B. C. W., D. C.
"	1787F., P. T	{ 2.84782 2.84890	{ 1.11069 1.11077	{ 1.504 1.504	{ 13.0929 ₀	{ 922 .41 922 .24	{ 1082 .32 1082 .08	{ 1082 .2 ₀	{ B. C. W., D. C.
Feb. 20, '09	1787A., P. T	{ 21.2475 21.1734	{ 4.4132 4.4165	{ 1.575 1.568	{ 61.9037	{ 921 .24 922 .23	{ 1091 .77 1092 .12	{ 1091 .9 ₄	{ K. elect., D. C.
"	" "	{ 21.2475 21.1734	{ 4.4132 4.4165	{ 1.575 1.568	{ 61.8890	{ 920 .90 921 .90	{ 1091 .28 1091 .63	{ 1091 .4 ₅	{ M. P., K. elect., D. C.
Mar. 1, '09	" "	{ 21.1730 21.1418	{ 4.4165 4.4179	{ 1.568 1.570	{ 61.8642	{ 921 .35 921 .76	{ 1090 .83 1091 .75	{ 1091 .2 ₉	{ B. C. W., A. C.
Mar. 6, '09	" "	{ 21.1418 21.0862	{ 4.4179 4.4203	{ 1.570 1.568	{ 61.8733	{ 921 .96 922 .72	{ 1092 .04 1092 .85	{ 1092 .4 ₄	{ E. & A., Drops, A. C.
"	" "	{ 21.1418 21.0862	{ 4.4179 4.4203	{ 1.570 1.568	{ 61.8593	{ 921 .65 922 .41	{ 1091 .60 1092 .40	{ 1092 .0 ₀	{ E. & A., Drops, A. C.
Apr. 20, '09	1787E., P. T	{ 2.92226	{ 0.50540	{ 1.803	{ 7.49870	{ 905 .5	{ 1106	{	{ B. C. W., A. C.
Feb. 18, '09	{ Palladium 1787G., W. B	{ 1.86284	{ .62696	{ 2.878	{ 6.89936	{ 803 .32	{ 1151	{	{ Old Pot., D. C.
"	" "	{ 1.86284	{ .62696	{ 2.878	{ 6.89978	{ 803 .4	{ 1151	{	{ M. P. Old Pot., D. C.
"	" P. T	{ 1.86802	{ .62890	{ 2.884	{ 6.91594	{ 802 .7	{ 1153	{	{ Old Pot., D. C.

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TABLE 10.—Freezing Point of Silver-Copper Alloy.

Date	Therm. No.	R ₀	F. I.	δ	R _{Ag Cu}	pt _{Ag-Cu}	t	t mean	Remarks
71 Ag-29 Cu.									
Mar. 26, '09	1787C., P. T	{ 3.48156	1.34104	1.509	{ 12.8645	{ 699°.68	779°.64	{ 779°.52	F. P.
		{ 3.48153	1.34099	1.505		{ 699°.70	779°.39		
"	"	{ 3.48156	1.34104	1.509	{ 12.8646	{ 699°.68	779°.64	{ 779°.52	"
		{ 3.48153	1.34099	1.505		{ 699°.71	779°.40		
Mar. 27, '09	1787F., P. T	{ 2.84881	1.11084	1.508	{ 10.6244	{ 699°.97	779°.94	{ 779°.60	"
		{ 2.84974	1.11121	1.504		{ 699°.65	779°.26		
"	1787A., P. T	{ 21.0734	4.4213	1.565	{ 51.9268	{ 697°.84	781°.10	{ 781°.24	"
		{ 21.0620	4.4239	1.572		{ 697°.68	781°.38		
"	"	{ 21.0734	4.4213	1.565	{ 51.9251	{ 697°.80	781°.05	{ 781°.18	"
		{ 21.0620	4.4239	1.572		{ 697°.64	781°.32		
"	"	{ 21.0734	4.4213	1.565	{ 51.9286	{ 697°.88	781°.15	{ 781°.29	M. P. Fast melt.
		{ 21.0620	4.4239	1.572		{ 697°.72	781°.43		
72 Ag-28 Cu (Eutectic).									
Mar. 31, '09	1787C., P. T	{ 3.48153	1.34099	1.505	{ 12.8623	{ 699°.54	779°.18	{ 779°.16	F. P.
		{ 3.48164	1.34105	1.505		{ 699°.50	779°.13		
Mar. 29, '09	1787F., W. B	{ 2.85542	1.11295	1.504	{ 10.6425	{ 699°.68	779°.28	{ 779°.20	"
		{ 2.85567	1.11312	1.504		{ 699°.55	779°.13		
"	" P. T	{ 2.84974	1.11121	1.508	{ 10.6249	{ 699°.70	779°.60	{ 779°.46	"
		{ 2.84987	1.11119	1.504		{ 699°.70	779°.32		

Mar. 29, '09	1787F., W. B	{ 2.85542 1.11295 1.504	{ 10.6434 { 699 .76 779 .40	{ 779 .32	"
		{ 2.85567 1.11312 1.504		{ 699 .63 779 .24	
"	" P. T	{ 2.84974 1.11121 1.508	{ 10.6249 { 699 .70 779 .60	{ 779 .46	M. P.
		{ 2.84987 1.11119 1.504		{ 699 .70 779 .32	
Mar. 31, '09	478, P. T	{ 5.16618 1.99957 1.503	{ 19.15225 { 699 .45 778 .93	{ 778 .92	F. P.
		{ 5.16566 2.00000 1.505		{ 699 .33 778 .91	
"	"	{ 5.16618 1.99957 1.503	{ 19.15424 { 699 .55 779 .06	{ 779 .05	"
		{ 5.16566 2.00000 1.505		{ 699 .43 779 .04	

73 Ag-27 Cu.

Apr. 1, '09	1787C., P. T	{ 3.48153 1.34099 1.505	{ 12.8620 { 699°.52 779°.15	{ 779°.12	F. P.
		{ 3.48164 1.34105 1.505		{ 699 .48 779 .10	
"	"	{ 3.48153 1.34099 1.505	{ 12.8651 { 699 .75 779 .45	{ 779 .42	M. P. Fast melt.
		{ 3.48164 1.34105 1.505		{ 699 .71 779 .40	
"	1787A., P. T	{ 21.0620 4.4239 1.572	{ 51.9222 { 697 .58 781 .25	{ 781 .14	F. P.
		{ 21.0617 4.4243 1.570		{ 697 .52 781 .03	
Apr. 19, '09	1787E., P. T	{ 2.92276 0.50517 1.803	{ 6.39844 { 688 .02 784 .96	{ 784 .82	M. P.
		{ 2.92226 .50540 1.803		{ 687 .81 784 .67	
"	"	{ 2.92276 .50517 1.803	{ 6.39792 { 687 .92 784 .82	{ 784 .68	F. P.
		{ 2.92226 .50540 1.803		{ 687 .70 784 .53	
"	"	{ 2.92276 .50517 1.803	{ 6.39741 { 687 .82 784 .69	{ 784 .54	"
		{ 2.92226 .50540 1.803		{ 687 .60 784 .39	
Mar. 27, '09	1787G., P. T	{ 1.87397 .63084 2.883	{ 5.85843 { 631 .61 787 .85	{ 787 .59	"
		{ 1.87312 .63101 2.878		{ 631 .58 787 .33	

Note.—All determinations in Table 10 in Acheson graphite crucible.

TABLE 13.—Temperature Corrections for Platinum of Different δ .

[Thermometer calibrated by Callendar method, ice, steam, and S. B. P.]

Temp. °C	Correction in °C for values of δ given below							
	1.525	1.550	1.575	1.600	1.650	1.700	1.800	1.900
200°	+0.02	+0.05	+ 0.08	+ 0.10	+ 0.14	+ 0.16	+ 0.20	+ 0.21
300°	+ .02	+ .05	+ .08	+ .11	+ .19	+ .27	+ .45	+ .55
400°	.00	.00	+ .01	+ .03	+ .08	+ .14	+ .29	+ .37
500°	— .02	— .05	— .09	— .11	— .18	— .24	— .39	— .57
600°	— .09	— .18	— .30	— .40	— .62	— .88	— 1.42	— 1.96
700°	— .33	— .70	— 1.03	— 1.32	— 1.78	— 2.2	— 2.9	— 3.5
800°	— .90	— 1.65	— 2.24	— 2.7	— 3.6	— 4.4	— 5.8	— 7.1
900°	— 1.90	— 3.1	— 4.0	— 4.9	— 6.5	— 8.1	— 10.8	— 13.5
1000°	— 3.3	— 5.2	— 6.8	— 8.2	— 10.7	— 13.1	— 17.1	— 20.8
1100°	— 5.5	— 8.1	— 10.3	— 12.2	— 15.7	— 18.7	— 24.3	— 29.1

TABLE 14.—Precision with Thermometer 1787C.

Metal	Freezing Point	Number of Dets.	Range
Cd, Baker & Adamson.....	320° .39	2	0° .01
Pb, Baker & Adamson.....	327 .47	2	.01
Zn, "Kahlbaum".....	419 .36	4	.01
Sb, "Kahlbaum," 1908.....	630 .71	3	.35
Sb, "Kahlbaum," M. P.....	630 .81	3	.15
71 Ag-29 Cu.....	779 .52	2	.00
Ag, Mint.....	960 .90	2	.10
Cu, B. C. W....	1083 .30	3	.29
Cu, E. & A. drops.....	1083 .95	2	.02
Cu, Kahlb. Electrolytic.....	1083 .40	2	.10

Precision with Thermometer 1787A.

Sn, Baker & Adamson.....	231° .89	2	0° .10
Cd, "Kahlbaum".....	321 .05	2	.04
Zn, Baker & Adamson.....	419 .40	2	.01
Sb, "Kahlbaum".....	630 .50	2	.07
71 Ag-29 Cu.....	779 .40	2	.04
Ag, Mint.....	961 .00	2	.02
Cu, E. & A. drops.....	1083 .3	2	.40

TABLE 16.—Changes of Zero, F. I. and c of 1787C. ($\delta = 1.503$.)

Date	Obs.	R ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer 1787C
Nov. 21, 1908....	3.4831	1.34163	0.00385183	After Sb. 2 times (previously annealed)
Jan. 4, 1909....	3.48779	1.3429800385052	0	" Zn 2 "
21.....	1	3.47971	0	1.33989	0	058	+ .001 ₅	" " 1 " (mica frame repaired)
23.....	2	71	.000	89	.000	058	+ .001 ₅	" " 1 "
25.....	3	64	-.005	89	.000	066	+ .003 ₆	" " 2 "
26.....	4	70	-.001	91	+ .001 ₅	065	+ .003 ₄	" " 1 "
27.....	5	70	-.001	94	+ .003 ₇	073	+ .005 ₄	" Sb 1 "
28.....	6	80	+ .007	94	+ .003 ₇	062	+ .002 ₆	" " 1 "
29.....	7	80	+ .007	94	+ .003 ₇	062	+ .002 ₆	" " 1 "
Feb. 6.....	8	84	+ .010	" Sn 1 "
8.....	9	68	-.002	1.34058	+ .051	260	+ .054	" Annealing 2 hrs. at 1100° +
20.....	10	04	-.050	76	+ .065	383	+ .086	" Cu 1 time
24.....	11	88	+ .013	94	+ .078	339	+ .074	" " 2 "
Mar. 1.....	12	47	-.018	" Cd 2 "
2.....	13	3.48041	+ .052	80	+ .068	242	+ .049	" Cu 1 "
6.....	14	3.48128	+ .117	86	+ .072	163	+ .029	" " 2 "
12.....	15	56	+ .138	104	+ .085	184	+ .034	" " 2 "
21.....	16	53	+ .136	099	+ .082	173	+ .031	" Ag-Cu. 2 "
26.....	17	64	+ .144	105	+ .087	178	+ .033	" " 2 "
Apr. 1.....	18	64	+ .144	105	+ .087	178	+ .033	" " 1 "
June 7.....	19	74	+ .152	114	+ .093	192	+ .036	" Cd, 3 times; Pb, 2 times.

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TABLE 17.

Changes of Zero, F. I. and c of 1787F. ($\delta=1.503$.)

Date	Obs.	R ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer 1787 F
ON WHEATSTONE BRIDGE.								
1909.								
Feb. 10	1	2.85355	0	1.11228	0	0.00389787	0	After annealing 2 hrs. at 1100°
25	2	344	— .010	249	+ .019	876	+ .023	“ S 2 times
Mar. 3	3	459	+ .093	274	+ .041	807	+ .005	“ Cu 2 “
9	4	456	+ .091	265	+ .033	780	— .002	“ Ag 4 “
29	5	542	+ .168	295	+ .060	904	+ .030	“ Cd 1; Sn 4, Pb 2, Ag—Cu 1
30	6	567	+ .190	312	+ .076	793	+ .0016	“ Ag—Cu 2 times
ON POTENTIOMETER.								
Feb. 9	1'	2.84782	0	1.11069	0	0.00390014	0	After annealing 2 hrs. at 1100°
25	2'	776	— .005	“ S 2 times
Mar. 3	3'	890	+ .097	077	+ .007	.00389894	— .031	“ Cu 2 “
16	4'	881	+ .089	087	+ .016	942	— .018	“ Ag 4 “
29	5'	974	+ .173	121	+ .047	933	— .021	“ Cd 1, Sn 4, Pb 2, Ag—Cu 1
30	6'	987	+ .184	119	+ .045	909	— .027	“ Ag—Cu 2 times

TABLE 18.
Changes of Zero, F. I. and c. of 478. ($\delta = 1.507$.)

Date	Obs.	R ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer, 478
1909.								
Feb. 5.....	1	5.15665	0	1.99745	0	0.00387354	0	After annealing 2 hrs. at 1100°.
23.....	2	5.16070	+.202	969	+.112	464	+.028	" 5 mins. in Cu. Coil straightened.
25.....	3	248	.292	2.00013	.134	436	+.022	" Cu 1 time.
Mar. 6.....	4	514	.424	11	.133	232	-.032	" " 1 "
12.....	5	618	.476	1.99957	.106	050	-.078	" Ag 2 times.
31.....	6	566	.450	2.00000	.128	172	-.047	" Ag-Cu 2 "

TABLE 19.
Changes of Zero, F. I. and c. of 479. ($\delta = 1.512$.)

Date	Obs.	R ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer, 479
Dec. 8, 1908....	1	4.27342	0	1.64585	0	0.00385136	0	Old thermometer reannealed.
9.....	2	23	-.012	85	.000	154	+.005	After Zn 4 times.
Jan. 26, 1909....	3	83	+.025	85	.000	100	-.009	" " 3 "
Feb. 1.....	4	45	+.002	" Sb 2 "
1.....	5	53	+.007	72	-.008	096	-.010	" " 1 time.
2.....	6	04	-.023	55	-.018	101	-.009	" " 1 "
3.....	7	50	+.005	" " 1 "
4.....	8	56	+.008
8.....	9	4.26905	-.271	04	-.049	341	+.053	" 2 hrs. at 1100°+.
17.....	10	812	-.321	46	-.024	525	+.101	" S.
23.....	11	4.27916	+.348	497	-.053	4414	-.187	" Cu 2 times.
25.....	12	4.28715	+.832	500	-.052	3705	-.372	" Cu 1 time.
Mar. 3.....	13	823	+.898	13	-.044	638	-.389	" Ag 1 "

TABLE 20.—Changes of Zero, F. I. and c. of 1787A. ($\delta = 1.570$.)

Date	Obs.	R. ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer, 1787 A
Nov. 19, 1908....	1	21.3476	0	4.4067	0	0.00206426	0	First annealed at 1100°.
21.....	2	384	— .21	67	.000	515	+ .043	After Zn 1 time and Sb 1 time.
Dec. 15.....	3	61	— .26	71	+.009	556	.063	“ Zn 4 times.
Jan. 22, 1909....	4	96	— .18	76	.020	546	.058	“ Zn 3 “
25.....	5	95	— .18	76	.020	547	.059	“ Zn 2 “
26.....	6	84	— .21	73	.014	543	.057	“ Sb 1 time.
27.....	7	72	— .24	77	.023	573	.071	“ Sb 1 “
28.....	8	48	— .29	77	.023	597	.083	“ Sb 2 times.
29.....	9	47	— .29	78	.025	602	.085	“ Sb 2 “
Feb. 6.....	10	41	— .31					“ Sn 1, and S.
6.....	11	21.2475	—2.27	4.4132	.148	.00207704	.619	“ Annealing 2 hrs. at 1100°+.
24.....	12	.1734	—3.94	.41648	.222	208586	1.046	“ Cu 1 time.
Mar. 1.....	13	.1730	—3.95					“ Cd 2 times.
2.....	14	.1418	—4.66	.41786	.254	208963	1.229	“ Cu 1 time.
9.....	15	.0827	—5.99	.42539	.424	209906	1.686	“ Cu 2 times.
11.....	16	.0862	—5.91	.03	.309	872	1.669	“ Several hrs. in S.
12.....	17	.0734	—6.20	128	.331	804	1.636	“ Ag 2 times.
27.....	18	.0620	—6.46	39	.390	210042	1.752	“ Ag-Cu 3 times.
Apr. 1.....	19	.0617	—6.46	43	.399	065	1.763	“ “ 1 time.

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TABLE 21.
Changes of Zero, F. I. and c. of 1787E. ($\delta = 1.803$.)

Date	Obs.	R ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer, 1787 E
1909.								
Apr. 15.....	1	2.92415	0	0.50480	0	0.00172630	0	After annealing 1.5 hrs. at 1100°.
17.....	2	382	— .006	511	+ .06	757	+ .08	“ Sb and Ag 2 times each.
19.....	3	226	— .039	540	+ .19	948	+ .19	“ Ag-Cu 2 times.

TABLE 22.—Changes of Zero, F. I. and c of 1787G. ($\delta = 2.890$.)

Date	Obs.	R ₀	ΔR_0 in ° C	F. I.	Per cent change in F. I.	c	Per cent change in c	History of Thermometer, 1787 G
ON WHEATSTONE BRIDGE.								
Feb. 1, 1909....	1	1.87254	0	0.627606	0	0.00335163	0	After annealing at 1100°.
2.....	2	398	+ .23	.62817	+ .089	206	+ .013	" Sb 2 times.
3.....	3	400	+ .23	
3.....	4	414	+ .26	" Sb 1 time.
8.....	5	6268	—1.57	.62696	— .104	6590	+ .426	" 2 hrs. at 1100°–1150°.
17.....	6	284	—1.55	
Mar. 15.....	7	1.87869	0	.63124	0	0.00335998	+ .241	Thermometer repaired. Annealed 5 min. in gas furnace.
16.....	8	759	— .174	158	+ .054	6378	+ .362	After Ag 2 times.
22.....	9	756	— .179	" Cd and Sn 1 time each.
ON POTENTIOMETER.								
Jan. 27.....	1'	1.87345	0	0.62809	0	0.00335258	0	After annealing.
Feb. 1.....	2'	504	+ .25	861	+ .083	252	— .005	" Sb 2 times.
5.....	3'	6802	— .86	890	+ .129	6667	+ .420	" Annealing 2 hrs. at 1100°–1150°.
Mar. 12.....	4'	1.87498	0	0.63098	0	.00336499	+ .370	Thermometer repaired. Annealed 5 min. in gas furnace.
16.....	5'	397	— .17	084	— .022	633	+ .400	After Ag 2 times.
30.....	6'	312	— .30	101	+ .005	876	+ .482	" Ag-Cu 3 "

TABLE 23.
Changes of Zero, F. I., and c. of Nos. 9837 and 9838.

Heat Treatment.	9837. $\delta=1.54_0$			9838. $\delta=1.55_7$		
	R ₀	F I	c	R ₀	F I	c
5 hrs. at 1100°	0.11275	0.04324	0.0038350	0.10951	0.04192	0.0038280
5 " " 1200°	78	32	411	48	200	363
1.5 " " 1250°	83	26	340	50	190	265
1 " " 1300°	88	50	536	56	220	518
2 " " 1250°	307	25	251	54	184	196

TABLE 24.
Changes in Thermometers at High Temperatures.

		Thermometers in copper. ¹					
		1787 C	1787 F	1787 A	478	479	1787 G ⁴
pt ₁ -pt ₂ ¹⁰	Mean	+0°.12	+0°.28	-0°.78	+0°.21	+0°.38	See 9.
	Range.....	+.03 to +.19	+.17 to +.39	-.41 to -1.00	+.12 to +.30	+.14 to +.50	
t ₁ -t ₂ ¹¹	Mean	+0.07	-0.75	-0.65	-.15	-.33	“
	Range.....	+.26 to -.50	+.24 to -1.75 ⁵	-.35 to -.92	-.12 to -.18	+.36 to -.68	

Thermometers in silver.²

$pt_1-pt_2^{10}$	Mean.....	0°.00	-0°.10	-0°.17	See 8.	+0°.06
	Range.....	+.05 ⁶ to -.07 ⁷	-.10 to -.10	-.17 to -.17		+.22 to -.26
$t_1-t_2^{11}$	Mean.....	-0.19	+0.21	+0.24	"	+0.14
	Range.....	+.13 ⁷ to -.41 ⁶	+.21 to +.21	+.23 to +.24		+.49 to -.54

Thermometers in Ag-Cu Alloy.³

$pt_1-pt_2^{10}$	Mean.....	+0°.01	+0°.12	+0°.13	+0°.12	Not used.	Not used.
	Range.....	+.04 to -.03	+.32 to .00	+.16 to +.06	+.12 to +.12		
$t_1-t_2^{11}$	Mean.....	+0.13	+0.04	-0.16	+0.02	" "	" "
	Range.....	+.25 to +.05	+.28 to -.68	+.18 to -.28	+.02 to +.02		

¹ Thermometers had been previously annealed in electric furnace at 1050-1100° for several hours.

² Thermometers had been previously used in copper.

³ Thermometers had been previously used in copper and silver.

⁴ Palladium thermometer.

⁵ This large difference is due to a too low value of δ (1.494) observed before and a too high value of δ (1.509) observed after the two determinations in molten copper.

⁶ Thermometer 1787 F used as potential terminal thermometer.

⁷ Thermometer 1787 F used as Wheatstone Bridge three-lead compensated thermometer.

⁸ Thermometer 479 used once. No calibration after, as thermometer found broken on cooling.

⁹ Thermometer 1787 G broke; could not be recalibrated after 3 determinations in copper of Feb. 18, 1909.

¹⁰ pt_1 =platinum temperature computed from calibration before determination of freezing point.

pt_2 =platinum temperature computed from calibration after determination of freezing point.

¹¹ t_1 =true temperature computed from calibration before determination of freezing point.

t_2 =true temperature computed from calibration after determination of freezing point.

TABLE 26.

Comparison of Platinum and Palladium Thermometers at Freezing Point Temperatures.

[Callendar's Formula. Calibration Data: Ice, Steam, Sulphur ($444^{\circ}.70$).]Platinum $\begin{cases} \delta=1.505 \\ c=0.00385 \text{ to } 389 \end{cases}$ Palladium $\begin{cases} \delta=2.890 \\ c=0.00336 \end{cases}$

Metal.	Temperature by Platinum Therm.	Temperature by Palladium Therm.	$\frac{1}{\text{Pd-Pt}}$	Range, ° C.	Number of—	
					Obs.	Samples
Sn.	231°.90	231°.62	+ 0°.28	1	1
Cd.	321 .01	320 .52	— .49	1	1
Pb.	327 .58	326 .90	— .68	1	1
Sb.	630 .71	632 .99	+ 2 .28	0.34	3	2 K
Ag.-Cu.	779 .20	787 .59	+ 8 .39	1	1
Ag.	960 .88	992 .49	+31 .61	.13	3	1
Cu.	1,083 .0	1,152	+69	2	2	1

TABLE 27.

Corrections to the Palladium Thermometer.

[Palladium Temperatures computed by Callendar's Formula. Calibration Data, Ice, Steam, Sulphur ($444^{\circ}.70$).]

Temperatures centigrade (Platinum Therm.)	Temperatures centigrade (Palladium Therm.)	Correction to temperatures with Palladium Therm.
200°	199°.80	+0°.20
300	299 .56	+ .44
400	399 .70	+ .30
500	500 .40	— .40
600	601 .5	— 1.5
700	704 .7	— 4.7
800	810 .5	—10.5
900	921 .0	—21.0
1,000	1,040 .8	—40.8
1,100	(1,175)	(—75)

XI. BIBLIOGRAPHY.

The following is not intended to be a complete bibliography of electrical resistance thermometry, but it is hoped that all the more important original contributions to the subject are included, as well as a sufficient number of technical papers treating of the applications of the resistance thermometer. Many duplicate references, mainly from the engineering press and reviewing journals, are omitted. Most of the English papers and many of the others will be found abstracted in the *London Electrician*.

Some references are given which deal with those properties of platinum which are of interest in connection with a study of its resistance at high temperatures.

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