

Solubility of Anthracene in Binary Alcohol + 2-Propoxyethanol Solvent Mixtures

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Experimental solubilities are reported for anthracene dissolved in seven binary mixtures containing 2-propoxyethanol with 1-propanol, 2-propanol, 1-butanol, 2-butanol, 1-pentanol, 1-octanol, and 3-methyl-1-butanol at 25 °C. Results of these measurements are used to test two mathematical representations based upon the combined nearly ideal binary solvent (NIBS)/Redlich–Kister equation and modified Wilson model. For the seven systems studied, both equations were found to provide an accurate mathematical representation of the experimental data, with an overall average absolute deviation between measured and calculated values being on the order of 0.5%.

Introduction

Solid–liquid equilibrium data of organic nonelectrolyte systems are becoming increasingly important in the petroleum industry, particularly in light of present trends toward heavier feedstocks and known carcinogenicity/mutagenicity of many of the larger polycyclic aromatic compounds. Solubility data for a number of polycyclic aromatic hydrocarbons (*i.e.*, anthracene and pyrene) and heteroatom polynuclear aromatics (*i.e.*, carbazole, dibenzothiophene, and xanthene) have been published in the recent chemical literature (for a listing of references, see Acree (1994, 1995a,b)). Despite efforts by experimentalists and scientific organizations, in terms of both new experimental measurements and critically-evaluated data compilations, there still exist numerous systems for which solubility data are not readily available.

To address this problem, researchers have turned to group contribution methods and semiempirical expressions to predict desired quantities. Group contribution methods have proved fairly successful in estimating solid solubility in pure and binary solvent mixtures from structural information. Practical application, though, is limited to systems for which all group interaction parameters are known. Interaction parameters can be evaluated from liquid–vapor, liquid–liquid, and solid–liquid equilibrium data. It is important that the data base contain as many different functional groups as possible, preferably with adequate representation from both mono- and multifunctional solute/solvent molecules to permit evaluation of potential synergistic effects. The data base should contain sufficient experimental values near infinite dilution in the event that one wishes to determine separate interaction parameters for finite concentration and infinite dilution activity coefficient predictions.

Continued development of solution models for describing the thermodynamic properties of a solute in binary solvent systems requires that a large data base be available for assessing the applications and limitations of derived expressions. Currently, only a limited data base exists for crystalline nonelectrolyte solubility in binary solvent mixtures. For this reason, anthracene solubilities were determined in seven binary alcohol + 2-propoxyethanol solvent mixtures. Results of these measurements are used to further test the descriptive abilities of several previously derived expressions.

Experimental Methods

Anthracene (Aldrich, 99.9+%) was used as received. 2-Propoxyethanol (Aldrich, 99+%), 1-propanol (Aldrich, 99+%, anhydrous), 2-propanol (Aldrich, 99+%, anhydrous), 1-butanol (Aldrich, HPLC, 99.8+%), 2-butanol (Aldrich, 99+%, anhydrous), 1-pentanol (Aldrich, 99%), 3-methyl-1-butanol (Aldrich, 99+%, anhydrous) and 1-octanol (Aldrich, 99+%, anhydrous) were stored over both anhydrous sodium sulfate and molecular sieves before being fractionally distilled. Gas chromatographic analysis showed solvent purities to be 99.7 mol % or better. Karl Fischer titration gave water contents (mass/mass %) of <0.01% for all seven alcohols used. Binary solvent mixtures were prepared by mass so that compositions could be calculated to 0.0001 mole fraction.

Excess solute and solvent were placed in amber glass bottles and allowed to equilibrate in a constant temperature water bath at (25.0 ± 0.1) °C for at least three days (often longer). Attainment of equilibrium was verified by repetitive measurements after a minimum of three additional days and by approaching equilibrium from supersaturation by preequilibrating the solutions at a higher temperature. Aliquots of saturated anthracene solutions were transferred through a coarse filter into a tared volumetric flask to determine the amount of sample and diluted quantitatively with methanol for spectrophotometric analysis at 356 nm on a Bausch and Lomb Spectronic 2000. Concentrations of the dilute solutions were determined from a Beer–Lambert law absorbance versus concentration working curve derived from measured absorbances of standard solutions of known molar concentration. Molar absorptivities of the nine standard solutions varied systematically with molar concentration, and ranged from ca. $\epsilon/(\text{L mol}^{-1} \text{cm}^{-1}) = 7450$ to $\epsilon/(\text{L mol}^{-1} \text{cm}^{-1}) = 7150$ for anthracene concentrations ranging from $C/M = 6.75 \times 10^{-5}$ to $C/M = 2.25 \times 10^{-4}$. Experimental anthracene solubilities in the seven binary alcohol + 2-propoxyethanol mixtures studied are listed in Table 1. Numerical values represent the average of between four and eight independent determinations, with the measured values being reproducible to within $\pm 1.3\%$.

Results and Discussion

Acree and co-workers (Acree and Zvaigzne, 1991; Acree *et al.*, 1991; Acree, 1992) suggested possible mathematical representations for isothermal solubility data based upon either a combined NIBS/Redlich–Kister model

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Table 1. Experimental Mole Fraction Solubilities of Anthracene (x_A^{sat}) in Binary Alcohol (B) + 2-Propoxyethanol (C) Solvent Mixtures at 25.0 °C

x_C^0	x_A^{sat}	x_C^0	x_A^{sat}
1-Propanol (B) + 2-Propoxyethanol (C)			
0.0000	0.000 591	0.4950	0.001 770
0.0709	0.000 735	0.7276	0.002 482
0.1486	0.000 900	0.8745	0.002 971
0.3066	0.001 276	1.0000	0.003 343
0.3965	0.001 502		
2-Propanol (B) + 2-Propoxyethanol (C)			
0.0000	0.000 411	0.5071	0.001 696
0.0696	0.000 541	0.7321	0.002 430
0.1504	0.000 712	0.8729	0.002 918
0.3099	0.001 115	1.0000	0.003 343
0.3980	0.001 365		
1-Butanol (B) + 2-Propoxyethanol (C)			
0.0000	0.000 801	0.5458	0.002 078
0.0829	0.000 969	0.7610	0.002 678
0.1666	0.001 153	0.8717	0.003 012
0.3408	0.001 549	1.0000	0.003 343
0.4386	0.001 805		
2-Butanol (B) + 2-Propoxyethanol (C)			
0.0000	0.000 585	0.5489	0.001 990
0.0903	0.000 775	0.7613	0.002 617
0.1743	0.000 963	0.8726	0.002 956
0.3419	0.001 398	1.0000	0.003 343
0.4364	0.001 659		
1-Pentanol (B) + 2-Propoxyethanol (C)			
0.0000	0.001 097	0.5881	0.002 352
0.1038	0.001 320	0.7874	0.002 849
0.1988	0.001 511	0.8883	0.003 095
0.3924	0.001 930	1.0000	0.003 343
0.4867	0.002 133		
3-Methyl-1-butanol (B) + 2-Propoxyethanol (C)			
0.0000	0.000 727	0.5879	0.002 092
0.1006	0.000 927	0.7922	0.002 735
0.2008	0.001 130	0.9069	0.003 080
0.3876	0.001 560	1.0000	0.003 343
0.4876	0.001 828		
1-Octanol (B) + 2-Propoxyethanol (C)			
0.0000	0.002160	0.6722	0.003168
0.1317	0.002510	0.8438	0.003406
0.2656	0.002651	0.9220	0.003487
0.4758	0.002885	1.0000	0.003343
0.5803	0.003039		

$$\ln x_A^{\text{sat}} = x_B^0 \ln(x_A^{\text{sat}})_B + x_C^0 \ln(x_A^{\text{sat}})_C + x_B^0 x_C^0 \sum_{i=0}^N S_i (x_B^0 - x_C^0)^i \quad (1)$$

or modified Wilson equation (Comer and Kopecni, 1990)

$$\ln[a_A(s)/x_A^{\text{sat}}] = 1 - x_B^0 \{1 - \ln[a_A(s)/(x_A^{\text{sat}})_B]\} / (x_B^0 + x_C^0 \Lambda_{BC}^{\text{adj}}) - x_C^0 \{1 - \ln[a_A(s)/(x_A^{\text{sat}})_C]\} / (x_B^0 \Lambda_{CB}^{\text{adj}} + x_C^0) \quad (2)$$

where the various S_i and $\Lambda_{ij}^{\text{adj}}$ "curve-fit" parameters can be evaluated via least squares analysis. In eqs 1 and 2 x_B^0 and x_C^0 refer to the initial mole fraction composition of the binary solvent calculated as if the solute were not present, $a_A(s)$ is the activity of the solid solute, and $(x_A^{\text{sat}})_i$ is the saturated mole fraction solubility of the solute in pure solvent i . The numerical value of $a_A(s) = 0.010 49$ (Acree and Rytting, 1983) used in the modified Wilson computations was calculated from

$$\ln a_A(s) = -\Delta_{\text{fus}} H_A (T_{\text{mp}} - T) / (RTT_{\text{mp}}) \quad (3)$$

the molar enthalpy of fusion $\Delta_{\text{fus}} H_A / (\text{J mol}^{-1}) = 28 860$

Table 2. Mathematical Representation of Anthracene Solubilities in Several Binary Alcohol (B) + 2-Propoxyethanol (C) Solvent Mixtures

binary solvent system (B + C)	eq 1		eq 2	
	S_i^a	% dev ^b	$\Lambda_{ij}^{\text{adj} c}$	% dev ^b
1-propanol + 2-propoxyethanol	0.952	0.1	1.803	0.6
	0.294		0.565	
	0.309			
2-propanol + 2-propoxyethanol	1.419	0.1	1.981	0.1
	0.477		0.694	
	0.201			
1-butanol + 2-propoxyethanol	0.715	0.3	1.694	0.4
	0.154		0.568	
	0.166			
2-butanol + 2-propoxyethanol	1.098	0.2	1.914	0.2
	0.364		0.554	
	0.125			
1-pentanol + 2-propoxyethanol	0.478	0.2	1.585	0.6
	0.157		0.655	
	0.205			
3-methyl-1-butanol + 2-propoxyethanol	0.689	0.3	1.585	0.7
	0.152		0.694	
	0.246			
1-octanol + 2-propoxyethanol	0.296	0.6	1.176	1.7
	-0.020		11.470	
	0.952			

^a Combined NIBS/Redlich-Kister curve-fit parameters are ordered as S_0 , S_1 , and S_2 . ^b Dev (%) = $(100/N) \sum |(x_A^{\text{sat}})^{\text{calc}} - (x_A^{\text{sat}})^{\text{exp}}| / (x_A^{\text{sat}})^{\text{exp}}$. ^c Adjustable parameters for the modified Wilson equation are ordered as $\Lambda_{BC}^{\text{adj}}$ and $\Lambda_{CB}^{\text{adj}}$.

(Weast, 1983) at the normal melting point temperature of the solute, $T_{\text{mp}}/\text{K} = 490.0$.

The ability of eqs 1 and 2 to mathematically represent the experimental solubility of anthracene in the seven binary alcohol + 2-propoxyethanol solvent systems is summarized in Table 2 in the form of curve-fit parameters and percent deviations in back-calculated solubilities. Careful examination of Table 2 reveals that both equations provide an accurate mathematical representation for how the solubility of anthracene varies with solvent composition. For the seven anthracene systems studied, the overall average absolute deviation between the experimental and calculated values is 0.3% and 0.6% for eqs 1 and 2, respectively, which is less than the experimental uncertainty.

From a computational standpoint, eq 1 will likely be preferred because most research groups involved in reporting thermodynamic properties have computer programs for evaluating the Redlich-Kister coefficients. With this idea in mind, we recommend not only that the future presentation of experimental isothermal solubility data for slightly soluble solid solutes dissolved in binary solvent mixtures include a tabulation of the actual observed values but if possible that the solubility data be mathematically represented by eq 1. Realizing that a single equation will not be applicable to all systems encountered, we further suggest eq 2 as an alternative mathematical representation for systems having extremely large solubility ranges and/or highly asymmetrical $\ln x_A^{\text{sat}}$ versus x_B^0 curves, such as the carbazole + alkane + tetrahydropyran systems reported previously (Acree *et al.*, 1991).

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