Rediscovery of the Elements

Courtois and Iodine

Bernard Courtois (1777-1838) serendipitously discovered iodine in Paris while reacting varec (ashes of seaweed) with sulfuric acid. He was using varec as a source of potash (potassium carbonate), necessary in the manufacture of saltpeter (potassium nitrate) used in gunpowder.

The history of saltpeter. The 7th century Taoist alchemists in western China stumbled onto a concoction of honey, sulfur, and saltpeter which caused smoke, fire, sparks, and thunder—called huo yao, or “fire drug”—thought to be the elixir of life.1 We know the incendiary mixture today as gunpowder.2 By the mid-1200s, the Chinese discovery had spread to the Western world.3 Roger Bacon (1214—1294) of Oxford recognized it as a mixture of “saltpeter, sulfur, and charcoal” (preferably of hazelwood), with the saltpeter in excess.2 The Battle of Crécy in 1346, at the beginning of the Hundred Years War, may be marked as the formal introduction of gunpowder on the battlefield as “gonnes” were used beside the famous longbowmen.1

In an attempt to understand the nature of gunpowder, Robert Boyle (1627—1691) studied gunpowder in a vacuum, generated by pumps originally used in his gas law experiments. Boyle was puzzled by the behavior of gunpowder, which combusted in an evacuated jar just as it did in the open atmosphere, when heated by the sun’s rays concentrated by a magnifying glass.1 Boyle did not understand that saltpeter (known today to be \( \text{KNO}_3 \)) contained its own “air.” Carl Wilhelm Scheele (1742—1786) got his first hint of the nature of this supporter of combustion when he observed sparks as saltpeter was heated with oil, and he went on to discover oxygen.2 Today we understand the basic reaction of gunpowder to be

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2 \text{KNO}_3 + 3 \text{C} + \text{S} \rightarrow \text{K}_2\text{S} + \text{N}_2 + 3 \text{CO}_2
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The evolved gases expand and serve as the propellant in rockets and guns.1 Saltpeter (meaning “salt-rock”) is the most soluble of the nitrates and wicks up rapidly in soil; it was first observed in China where a climate of alternatively cold/hot wet/dry climates allowed it to form on soil surfaces.1 In Medieval Europe, on stable and barn stone walls, saltpeter grew in the form of “blooms” resembling cotton tufts. Recent analyses show such growths of saltpeter are surprisingly pure, often greater than 95% \( \text{KNO}_3 \).4 Saltpeter diggers—Salpetergraber in Germany or Salpétier in French5—soon learned to excavate soil about stables, pigeon cotes, manure piles, or barnyards to collect the saltpeter therein.

As the demand for saltpeter increased, the saltpeter plantation/refinery was born, called the Nitrary (Nitrerie in French, Nitriarv in German)4 (Figure 1). The manager of a nitrary would lay down a stratum of lime and cover with manure, garbage, corpses, and any other organic material, including a periodic dowsing with the urine of “drinkers of wine or strong beer.”1 The major qualification of a nitrary manager was his ability to tolerate the incredible stench of his enterprise. After a few months the saltpeter diggers unearthed copious amounts of saltpeter-rich soil. An aqueous washing of the soil was reacted with ashes, filtered, concentrated by boiling, and cooled to produce white crystals of saltpeter. (Figure 2)
Today we understand the chemistry of the process. Ammonia released from organic material is oxidized to nitrates. The lime (CaO) keeps the pH 7-8 required by the nitrifying bacteria. The product, calcium nitrate ("lime saltpeter"), is useless because it is hygroscopic and the resulting gunpowder would absorb moisture. Hence, the calcium nitrate is reacted with ashes (potash; potassium carbonate) to precipitate the calcium salts and leave a potassium nitrate solution: Ca(NO$_3$)$_2$ (aq) + K$_2$CO$_3$ (aq) $\rightarrow$ CaCO$_3$ (s) + 2 KNO$_3$ (aq).5,6

During the 1600s Great Britain found a cheap source of saltpeter—the East India Company could allow import of the commodity from India at one-quarter of the domestic cost.1 (Note 1) France continued to produce its saltpeter from its home nitraries and itinerant saltpeter gatherers, but by necessity had to purchase some of its supply from Holland at an exorbitant price. Louis XVI (1752-1793) established the Gunpowder Administration in 1775, and as manager he appointed Antoin-Laurent Lavoisier, who promptly increased saltpeter production and improved the gunpowder manufacture process.6 This self-sufficiency would be vital when Great Britain blockaded France during the French Revolutionary Wars (1792-1802) and the Napoleonic War (1803-1815).

**Courtois in Dijon.** Bernard Courtois was born in Dijon, France, the son of a wine merchant Jean-Baptiste Courtois (1748-1807). The Courtois home was directly across the street from the Académie de Dijon (Figure 3), where the father won the position of chemistry demonstrator, hired by Louis-Bernard Guyton de Morveau (1737-1816), a lawyer and chemist who founded the chemistry department at Dijon Academy.7 Guyton de Morveau is well known as one of the authors of the modern system of chemical nomenclature in 1787 (the others being Lavoisier,8 Claude-Louis Berthollet (1748-1822),9 and Antoine Fourcroy (1755-1809).9 In 1780 Jean-Baptiste was also assigned manager of the new Saint Médard Nittry in Dijon (1.7 km east, on the outskirts of town) (Figure 4), one of the government-sponsored plantations created to cope with the saltpeter shortage.8 As a boy the precocious Bernard Courtois absorbed considerable chemical knowledge and laboratory experience at the Académie laboratory and at the nittry.

**Courtois in Paris.** In 1798, Bernard Courtois became a laboratory assistant at the École Polytechnique, arranged by Morveau who had moved to Paris in 1791 to become a member of the Legislative assembly and also a professor at l’École Polytechnique.7 At this time the École Polytechnique was located in the Palais Bourbon, directly across the Seine from the Place de la Révolution, today la Place de la Concorde, where Lavoisier was guillotined in 1794.10

At l’École Polytechnique Bernard Courtois worked with Fourcroy and Thenard (the discoverer of hydrogen peroxide, 1777-1857), then with Armand Séguin (1767-1835).7 Séguin is perhaps best known for serving as the human guinea-pig for Lavoisier’s respiration studies; classic drawings by Mme. Lavoisier show him with the “gas-masks” in Lavoisier’s laboratory at la Petit Arsenal.7 In Séguin’s laboratory Bernard worked with opium compounds, and apparently isolated morphine, although not formally acknowledged.7 Séguin delayed publication and was scooped in 1804 by Friedrich Wilhelm Adam Sertürner (1783-1841) of Paderborn, Germany, who gave the drug its modern name.8

In 1804, Bernard terminated his studies at l’École Polytechnique, which was becoming militarized upon Napoléon’s edict and moved to its new site near the Panthéon.7 Bernard now assisted his father, who had moved to Paris in 1802 to set up a nittry and had settled on Rue Saint Marguerite (today Rue Trousseau),7 about 1 km northeast of the Seine on the outskirts of town. Six years later, Bernard assumed ownership of his own nittry another kilometer north.
The discovery of iodine. Potash was scarce because of the Napoleonic wars (1803–1815), and Bernard Courtois had resorted to varec imported from the Brittany and Normandy sea-coasts. One day in late 1811 while cleaning the discoveries. 

He remained 1810–1821. It was here that he performed some research on the crystals, and discovered a fulminating powder when he reacted the dark crystals with ammonia—he had produced nitrogen triiodide. Not being a regular member of the scientific establishment, he could not immediately publish his findings. In May 1812 he shared his experiences with his colleagues Nicolas Clément (1779–1842) and Charles Bernard Désormes (1771–1862), two scientists from Dijon who had studied at l’École Polytechnique when Bernard was there. Unfortunately, they were slow in their analysis of the new substance and did not report it to the Académie for 18 months.

Sir Humphry Davy (1778–1829) of London embarked on a two-year scientific tour of Europe in 1813, and because of his fame Napoléon granted him special permission to travel through France. While in Paris (October 15–December 23, 1813) Davy was given a sample of Courtois’ “substance nouvelle” which he analyzed using the traveling laboratory chest he carried on all his journeys. Joseph Louis Gay-Lussac (1778–1850), a member of the elite Société d’Arcueil who had formulated the law of combining volumes, was motivated by this British intruder to initiate his own research. In a few weeks the two scientists had conducted detailed independent studies on this substance which had previously been shelved for 1½ years.

In December of 1813 there appeared in Annales de Chimie (the prestigious journal of the Académie des Sciences) several back-to-back papers on “une substance nouvelle dans le Vareck” (a new substance from varec), which was called “l’iode” from the Greek word for “violet.” These articles were authored by Courtois, Gay-Lussac, and Davy (Courtois’ paper was read by Clément, since Courtois was not a member of the Académie; he was identified only as a lowly “salpetrière from Paris”). Reflecting tensions between France and Great Britain, the (London) Royal Society delighted in noting (20 January 1814) that: “[Iodine] was discovered about two years ago; but such is the deplorable state of scientific men in France, that no account of it was published till the arrival of our English philosopher there.”

One who reads these December papers of Annales de Chimie can begin to appreciate the fascination and confusion of the chemists as they tried to comprehend this “substance nouvelle”—this was half a century before the appearance and the heavy weight of a metal, it dissolved in ether! It behaved like oxygen or chlorine because it reacted with metals. In fact, because of its odor, iodine was originally thought to contain chlorine. However, it was found to react with hydrogen gas to form not hydrochloric acid but a new acid (hydroiodic acid), proven by its reaction with lead and silver nitrate to form unknown yellow salts (instead of white). Ultimately, it was decided that it must be a “corp nouveau” (a “new body”) because it could not be broken down chemically, even in the presence of red-hot carbon.

In these December papers and others that followed, vigorous competition between Gay-Lussac and Davy developed into bitter contention about who should be credited with what. Duplicate claims were published, even with regard to who named the new element. Gay-Lussac said he should be credited for naming it “l’iode.” However, Davy maintained the correct name was iodine, in analogy to chlorine, which he had named in 1810 when he recognized its elemental nature. (Previously, chlorine had been called “oxymuriatic acid,” believed to be a compound of HCl and oxygen.) Fortunately, in spite of heated arguments about proper recognitions, it was agreed that Courtois, who could have easily been lost in the shuffle, should be given credit for the initial discovery.

In 1815 the Napoleonic Wars were over, and cheap Indian saltpeter became available. Courtois turned his total attention to iodine, and with the help of Clément and Désormes developed methods for extracting iodine from seaweed in quantity. By 1820, the medicinal properties of iodine had become known and Courtois became a commercial manufacturer. In 1831, he was awarded the Montyon Prize by l’Académie for his contributions to medical science, but he never benefited financially from his discovery and he died penniless.

On November 9, 1913, the scientific community commemorated in Dijon the 100th anniversary of the discovery of iodine by holding a banquet at the Dijon Académie and mounting a plaque on the old Rue Monge house (Figures 6 and 7). No bust or portrait was prepared for the celebration, because no likeness of Courtois was known just as for Smithson Tennant (the discoverer of osmium and iridium) (Note 2).

References.
5. U. Bretcher, personal communication. Consulting Engineer, Kuster and Hager, Uznach, Switzerland.
Staying Active in Chemical Outreach After Children

By Matt & Jen Schnippert, Gamma Beta 2003

Since my wife, Jen, and I had our first child less than a year after our initiation, it has been difficult to remain active in the Fraternity. Last summer, however, we were able to include our eldest child, Liisi, in our outreach activities. She was actually the source of one of the activities. This past spring, one of her friends' fathers, a graduate student in herpetology, brought in some reptiles to show her class at the on-campus preschool. All evening after coming home, all Liisi could talk about was how she wanted her mommy and daddy to come to school so she could be a helper like her friend had been. Because I had participated in a number of Gamma Beta’s magic shows, I talked to our chapter’s Master Magician, and to Liisi’s school’s director, and set up a date to perform a magic show.

I’ve performed a number of magic shows, most of which were for school-aged children, but Liisi’s class was made up of children ranging in age from 3-5. In our usual magic shows, we try to provide the kids with reasonably thorough explanations of things, even if the material goes over their head. We hope that by exposing them to the real explanations (and not oversimplifications), it will pique their curiosity and inspire them to learn more. With this age group, however, I knew immediately I would have to stick to demonstrations that were flashy and exciting, leaving detailed explanation for a few years down the road.

Several of the experiments we did were old standbys we’d done often at our museum-shows—bubbling dry ice in colored water, dry ice balloons, shattering a ball and flower with liquid nitrogen, and—the perennial favorite—elephant toothpaste. Because I had more freedom to work outside the box, I decided to try something we’d never done before and made liquid nitrogen ice cream. I found a recipe online and made sure to try it at home before doing it in front of the kids (we wouldn’t want to embarrass ourselves in front of our daughter’s class, after all), but it makes some of the best tasting ice cream I’ve ever had. It doesn’t last long or freeze well, so you need to carefully gauge how much you need. The kids were all amazed when I poured cream, sugar, and vanilla into a big pot, added liquid nitrogen, stirred with a big wooden spoon, and had delicious ice cream that they could all try in less than two minutes. It was so much fun, we used our leftovers to make it at the chapter meeting that evening. My girls liked it so much that when I brought home the Dewar after a magic show later this summer, Kaia, my two-year-old, pointed at the container and asked “ice cream?” I then, as any well-trained parent would do, went out to the store to get the necessary ingredients to make it.

GE Note: The GE never misses an opportunity to point out that he has the distinct pleasure of hearing the first journal reference on the liquid nitrogen ice cream demo (J. Chem. Educ. 1994, 71, 1080). Following from some patents on creating machines to prepare instant ice cream, which the GE has NO part of, there is at least one café that has opened in Chicago that features liquid nitrogen ice cream (http://www.icreamcafe.com/).

ACS & Alpha Chi Sigma Luncheon