Methane Activation by Group IVB Imido Complexes

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Abstract: An ab initio study of methane activation by group IVB imido complexes, when coupled with available experimental data, reveals an interesting picture of this important reaction. Initial activation of methane and (H2)2M=NH leads to the formation of a C-H alkane complexes bond by ~9 kcal mol-1. Experiment indicates that the polarity of the metal-ligand bond upon which the C-H is activated plays an important role in facilitating subsequent scission. Calculations support this hypothesis and suggest that formation of the alkane complex acts to increase C-H polarization, setting the stage for C-H cleavage. Calculated methane elimination barriers for (H2)2M(CH3)(NH2) (M = Ti, Zr, Hf) are in good agreement with experimental elimination in agreement with experimentally determined activation parameters. The TSs have a kite-shaped geometry with an obtuse angle about the H of the C-H bond being activated (H2) and a short MH1 distance, 1-2% greater than normal. The short MH1 distance suggests a stabilizing interaction, supported by a positive bond overlap population. Calculation of the intrinsic reaction coordinate demonstrates the importance of agostic interactions between N-H and M along the reaction coordinate.

Introduction

The functionalization of alkanes by discrete, well-characterized organometallic complexes has attracted the attention of numerous chemists. Much of the impetus for this work has been provided by the search for selective organic synthetic reagents and the ability to convert petrochemical feedstocks into more valuable commodity chemicals. Methane, the least reactive and most plentiful of the alkanes, has been described as the "Holy Grail" of C-H activation. Transition-metal complexes used to effect concerted C-H activation are of two main classes. The first group consists of low-valent late transition-metal complexes, and the activation mechanism involves oxidative addition. A large body of experimental and theoretical work has focused on these systems and oxidative addition. C-H activation is also effected by high-valent (d6 and d8p1) metals, predominantly of the early d or f block. Representative examples of high-valent methane activators include the lanthanide and group IIIIB metallocenes, organoactinides, and group IVB imido complexes. Their high reactivity hinders direct experimental probing of the potential energy surface for methane activation. Thus, computation can play an important role in "isolating" important species, submitting them to study, and understanding the interactions which give rise to an effective methane activation system. To this end we have initiated studies of high-valent methane activation systems to complement their more well-studied counterparts and the growing body of experimental data.

Rothwell has reviewed the literature of high-valent C-H activation complexes up to 1987. The high-valent, often d8, nature of the metal precludes oxidative addition. Proposed mechanisms, therefore, involve four-center [(2 + 2)] transition states (TSs) in which the formal oxidation state of the metal does not change. The activation of X-H bonds by early, high-valent complexes has been studied computationally, although to a lesser extent than low-valent systems. Goddard and Steigerwald have looked at the reactions of H2 with the model complexes Cl2M=MH (M = Sc, Ti) and [Cl2TiH]+. Rappe has extensively researched the reactivity of Ti and Sc complexes. Hoffmann et al. have performed an extended Huckel analysis of the Cp7LuR (R = H, Me) system. Cundari has reported an ab initio study of the reaction (NHSi')3ZrR (Si' = 1-Bu), N-Si = C6D6, cyclo-C5H11), which the C-H is activated plays an important role in facilitating subsequent scission. Calculations support this hypothesis and suggest that formation of the alkane complex acts to increase C-H polarization, setting the stage for C-H cleavage. Calculated methane elimination barriers for (H2)2M(CH3)(NH2) (M = Ti, Zr, Hf) are in good agreement with experimental elimination in agreement with experimentally determined activation parameters. The TSs have a kite-shaped geometry with an obtuse angle about the H of the C-H bond being activated (H2) and a short MH1 distance, 1-2% greater than normal. The short MH1 distance suggests a stabilizing interaction, supported by a positive bond overlap population. Calculation of the intrinsic reaction coordinate demonstrates the importance of agostic interactions between N-H and M along the reaction coordinate.

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(1) E-mail: cundari@memstvl.memst.edu.
(3) Rothwell, I. P., in ref 2, p 151.
(4) Numerous references to the original literature, experimental and computational, for the activation of C-H bonds by low-valent transition-metal complexes can be found in Saillard, I.-Y.; Hoffmann, R. J. Am. Chem. Soc. 1984, 106, 2006.
(7) The most researched family of high-valent, multiply bonded complexes which effect C-H activation are the biochemically important metal-oxo (L,M=O) complexes. Holm, R. H. Chem. Rev. 1987, 87, 1401.

(14) Preliminary results from our laboratories on the Cl2MCH3 + CH4 reaction, models of the metallocene systems, show substantial stabilizing MH interaction in the TS for methane activation (M = Sc, Y, Lu), as indicated by positive Mulliken bond overlap populations (~0.2).
when heated in the presence of methane (3 atm), yield the corresponding methyl complex (NH$_2$Si')ZrMe plus an equivalent of alkane. Reaction of the tris(amido)methyl complex with D$_2$ gives (NH$_2$Si')(ND$_2$)ZrD and methane. The data support a bis(amido)imido intermediate (NH$_2$Si')(Zr=NSi') capable of activating H$_2$ (or RH). Two pathways can be envisioned for the conversion of the bis(amido)imido to the tris(amido)alkyl. The first route is the 1,2-addition of C-H across the Zr-N bond of the bis(amido)imido to form tris(amido)alkyl directly. Addition of C-H across a Zr-amido bond followed by rapid H transfer is another possibility. The calculated barrier for H$_2$ addition across Zr-N is half that for addition across N-H$_2$ in (NH$_2$)$_2$Zr=N.$^{13}$H.

The direct pathway is also supported by the reactivity of related Cp$_2$Zr=NR complexes. Waleh et al.$^{16}$ have demonstrated that zirconocene-imido complexes Cp$_2$Zr=NR are capable of activating the C-H bond of benzene. In the zirconocene case, the presence of the imido intermediate was confirmed by X-ray crystallography.

Reported Ti-imido complexes display reactivity somewhat different from that of their Zr counterparts.$^{14,15}$ When (Ar$^+$O)$_2$(pp)$_2$Ti=NPH$_2$ (Ar$^+$ = 2,6-diisopropylphenyl; pp$^+$ = 4-pyridinopyridine) is refluxed in benzene for several hours, no C-H activation products are observed. Wolczanski et al.$^{16}$ have studied the reactivity of Ti-imido complexes; activation of benzene was effected by these complexes, but not methane. However, a related Ti-imido complex ((Si'O)$_2$Ti=NNSi') with different ancillary ligands has recently been shown to activate methane.$^{16}$

Increased steric congestion in five-coordinate versus three-coordinate Ti complexes cannot be ruled out as a factor in the observed reactivity differences.

For the present work, the transition metals of interest are from group IVB, with model complexes of the form (X),M=NH, (M = Ti, Hf, X = H; and M = Zr, X = H, NH$_2$, Cl). As part of a continuing focus on the electronic structure and reactivity of multiply bonded transition-metal complexes,$^{11}$ a computational study of the activation of methane by group IVB imido complexes was initiated. The results for the activation of methane (and the reverse reaction, methane extrusion) by group IVB imido complexes are reported here. Group IVB imido complexes were chosen because previous work$^{13,17}$ has shown that the chemistry of these systems can be accurately described and remain computationally tractable.

Computational Methods

The calculations described herein employ the ab initio quantum chemistry program GAMESS.$^{18}$ Effective core potentials (ECPs) replace the chemically less important core electrons and thus make calculations feasible for all rows of the transition series. The ECPs and valence basis sets of Stevens, Basch, Krauss, and Jasien (SBKJ)$^{19}$ are used for heavy atoms, while HSs are described with the -3 1G basis set. The ECP replaces the innermost core orbitals for the metals and all core orbitals for the main-group elements. Thus, the ns, np, nd, (n + 1)s, and (n + 1)p are treated explicitly for the d block; for the main-groups, ns and np are treated explicitly. For the transition-metal imido complexes, the SBKJ valence basis sets are quadruple- and triple-$\zeta$ for the s, p, and d shells, respectively; main-group elements have a double-$\zeta$ valence basis. The transition-metal potentials are generated from all-electron Dirac-Fock calculations and thus include Darwin and mass velocity relativistic corrections, while spin–orbital coupling is averaged out in potential generation as described in the original paper.$^{19}$ Basis sets for heavy main-group elements are augmented with a d polarization function.$^{20}$ Geometry optimizations employ the method of Baker$^{21}$ and, unless stated otherwise, are done at the restricted Hartree–Fock (RHF) level for closed-shell singlets. The present combination of ECPs, basis sets, and level of theory has been thoroughly validated for a large number of transition-metal complexes, particularly with respect to geometric prediction.$^{13,17}$ Bond lengths and bond angles involving transition metals are typically predicted within 1–3% of experimental models. Further comparisons between theory and experiment are presented below, with more extensive results elsewhere.$^{13,17}$ Frequencies are calculated at stationary points to identify them as minima (zero imaginary frequencies) or transition states (one imaginary frequency). Plotting the imaginary frequency and perturbing the geometry along this mode (in the forward or reverse direction), followed by reoptimization, is used to assess which transition state (TS) connects which minima. In some cases, the intrinsic reaction coordinate (IRC)$^{22}$ is followed from the TS toward reactants and products. Intrinsic stretching frequencies are calculated using the method of Boatz and Gordon.$^{23}$

Previous work has clearly demonstrated that d$^6$ transition-metal complexes like those studied here are well described at the single determinant level; thus, geometric parameters are accurately predicted since a flexible basis set is used.$^{13,17}$ Calculated energetics will usually be poor if electron correlation is ignored. For species described well at the single determinant level, the correlation contribution can be treated as a perturbation to the HF energy. For this reason, the correlation energy is calculated using second-order Møller–Plesset (MP2) theory.$^{24}$ Core electrons are not included in the MP2 active space. All quoted energies are determined at the MP2 level using the RHF geometries with zero-point energy (ZPE) correction included.

A series of test calculations were carried out on the TS for (H)$_2$Zr=NH + CH$_4$ $\rightarrow$ (H)$_2$Zr(CH$_3$)(NH$_2$) to assess the importance of correlation for the prediction of TS geometry. MCSCF$^{24}$ and RHF transition states are very similar. The bonds which break and form in the TS differ by an average of only 0.03 Å, while exocyclic bonds change even less. The coefficient of the Hartree–Fock configuration at the MCSCF-optimized TS is $\approx$92%. As a comparison, the TS for an organic analogue (the formaldehyde$^{25}$) has C-H 0.09 Å at a comparable level of theory. Thus, the MCSCF results support the present approach, MP2 energies at RHF geometries, used with great success in studies of main-group reactivity.$^{26,27}$

Result and Discussion

1. Reactants and Products. The geometry and electronic structure of transition-metal imido complexes, including the reactants in the present study, have been discussed elsewhere.$^{13,17}$ Only the highlights are sketched here. The geometry of methane is, of course, well known. As expected for imido complexes,
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Figure 1. Geometries of the methyl(amido) products calculated as described in the text. Other geometric parameters are similar to values reported elsewhere.\(^{(11)}\)

the imido (NH) ligand is coordinated in a linear fashion.\(^{(28)}\)

Metal-imido bond lengths are 1.65 (Ti), 1.80 (Zr), and 1.79 Å (Hf) for the parent complexes (H),M=NH. Calculated Zr-imido bond lengths in the dichloroimido and bis(amido)imido are 1.79 and 1.83 Å, respectively. Note the shortening of the Zr-imido bond upon replacement of hydrido with amido ligands.\(^{(29)}\)

Metal-ligand and N-H bond lengths and bond angles are typical.\(^{(11,17)}\)

The Ti-imido and Zr-dichloroimido minima are C\(_2\), i.e., trigonal planar about the metal, while (H),M=NH (M = Zr, Hf) and (NH\(_2\)),Zr=NH possess slightly pyramidal (C\(_1\)) geometries.

The calculated Zr=N stretching frequency for (NH\(_2\)),Zr=NH is 955 cm\(^{-1}\); it is important to note that this is not an intrinsically stretching frequency and that other normal modes are mixed in. Vibrational frequencies calculated at the RHF level are usually scaled by 0.9 to account for electron correlation effects (\(v_{Zr=NH}\) (scaled) = 860 cm\(^{-1}\)).\(^{(10)}\)

An IR band at 865 cm\(^{-1}\) was assigned to the Zr=NH stretch of the THF adduct of (NH\(_2\)Si(TMS)\(_2\)),Zr=NH.\(^{(29)}\)

The coordination of a Lewis base to the metal is another factor to be considered. The coordination of Lewis bases (e.g., NH\(_3\) and H\(_2\)O) to metal-imidos lowers the calculated metal-imido stretching frequency by only 10–20 cm\(^{-1}\) depending on the Lewis base and the metal.\(^{(11,7a)}\)

Thus, our calculated \(v_{Zr-NH}\) is in agreement with the assignment of the IR band at 865 cm\(^{-1}\) to a Zr=NH stretch.\(^{(8a)}\)

The products of methane activation are pseudotetrahedral-methyl(amido) complexes, Figure 1. While the dihydridoamido complexes prefer an eclipsed conformation of NH relative to the M-C bond, (Cl),Zr(CH\(_3\))(NH\(_2\)) is staggered and Zr(NH\(_2\))\(_3\)(CH\(_3\)) is \(C\(_3\)\), Figure 1. The various rotamers are within 1–2 kcal mol\(^{-1}\).


\((29)\) A preliminary X-ray structural analysis (Wolczanski, P. T., unpublished results) of (NH\(_2\)Si(TMS))\(_2\),Zr=NH\(_2\) reveals a geometry unlike any other reported Zr-imido complex, i.e., Zr=NH = 1.97 Å, Zr=NSi = 1.57 å. Further studies of this interesting complex will be forthcoming.


Table I. Calculated Enthalpic Data\(^{(a)}\)

<table>
<thead>
<tr>
<th>imido reactants</th>
<th>(\Delta H^\text{EN}_{\text{react}}) kcal mol(^{-1})</th>
<th>(\Delta H^\text{EN}_{\text{elim}}) kcal mol(^{-1})</th>
<th>(\Delta H_{\text{fus}}) kcal mol(^{-1})</th>
<th>(\Delta H_{\text{sol}}) kcal mol(^{-1})</th>
</tr>
</thead>
<tbody>
<tr>
<td>(H),Ti=NH</td>
<td>14.5</td>
<td>23.4</td>
<td>8.9</td>
<td>8.9</td>
</tr>
<tr>
<td>(H),Zr=NH</td>
<td>14.1</td>
<td>34.1</td>
<td>20.0</td>
<td>8.7</td>
</tr>
<tr>
<td>(H),Hf=NH</td>
<td>7.3</td>
<td>38.0</td>
<td>30.7</td>
<td>9.3</td>
</tr>
<tr>
<td>(NH(_2)),Zr=NH</td>
<td>12.0</td>
<td>33.4</td>
<td>21.4</td>
<td>0.8</td>
</tr>
<tr>
<td>(Cl),Zr=NH</td>
<td>12.9</td>
<td>34.9</td>
<td>22.0</td>
<td>0.8</td>
</tr>
</tbody>
</table>

\(^{(a)}\) Calculated enthalpic data (kcal mol\(^{-1}\)) for (X),M=NH + CH\(_4\rightleftharpoons (X),M(CH\(_3\))(NH\(_2\)). Using methods described in the text. \(\Delta H^\text{EN}_{\text{react}}\) is the barrier for C–H activation by imido reactant; \(\Delta H^\text{EN}_{\text{elim}}\) is the barrier for the reverse reaction, i.e., methane elimination from methyl(amido)-metal-imido product; \(\Delta H_{\text{fus}}\) is the reaction enthalpy; \(\Delta H_{\text{sol}}\) is the energy of the CT complex below separated products. All quantities are measured versus separated reactants. \*nc = not calculated.


\(^{(32)}\) (a) This assumption is based on the order in which bond dissociation energies for the homoleptic tetraneopentyl complexes of the group IVB metals.\(^{(14)}\)


\(^{(34)}\) Morse, J. J.; Parker, G. M.; Burke, T. J. Organometallics 1989, 8, 2471.
energies for the Ti-, Zr-, and Hf-imido complexes are 8.9, 8.7, and 9.3 kcal mol\(^{-1}\), respectively. These charge-transfer (CT) complexes were obtained by taking the TS for \(\text{C-H}^*\) activation (vide infra), distorting the geometry along the imaginary mode toward reactants, and then reoptimizing the geometry.\(^{16}\) Upon reoptimization, stationary points are obtained (2) in which the geometries of the CH4 and (H)M=NH fragments are nearly identical to those calculated for the individually optimized reactants, with metal-carbon distances of 2.63 (Ti), 2.80 (Zr), and 2.79 Å (Hf). A Mulliken population analysis reveals that 0.15 (Ti), 0.23 (Zr), and 0.17 (Hf) electrons are transferred from methane to the metal; calculated charges do not change on the other atoms of the imido complex. The charges on C and H in isolated methane are \(-0.24\) and \(+0.06\), respectively. In the CT complexes, the calculated atomic charges assume average values of \(-0.48 \pm 0.02\) for C, 0.15 \pm 0.01 for H, and 0.22 \pm 0.01 for H, 2. Thus, the C-H bonds which were polarized (C-H) in isolated methane become more so in 2. Wolczanski et al.\(^{16}\) have proposed that the extremely polarized metal-imido bond induces greater C-H\(^*\) polarization, aiding in subsequent C-H scission.

The importance of the polarity of the metal-activating ligand bond has been pointed out by various researchers.\(^{30,36,37}\) The present results support the "polarization hypothesis." The calculations further suggest that the formation of a methane complex is the agency through which greater polarization of the H(CH)\(^*\) bond is brought about, facilitating what has been envisaged\(^{16,35}\) as a concerted heterolytic C-H cleavage.

3. Transition States. The most plausible TS geometry for methane activation by a high-valent complex is a four-center one, since no change in the formal oxidation state of the metal occurs.\(^{35}\) An initial search for a TS connecting CH4 and (H)Zr=NH with (H)Zr=CH3(NH)\(^*\) was done in C, symmetry. A C, stationary point was found which possesses two imaginary frequencies, 1940/ (H-transfer from C to N) and 1721 (rotation about the methyl group). The latter mode destroys C, symmetry and leads to a C, transition state which is 2 kcal mol\(^{-1}\) lower (at the RHF level with ZPE correction). The geometry of the C, and C, stationary points are essentially identical, other than in torsional angles involving the methyl group. Despite the reduction in symmetry, the atoms which comprise the four-center TS are nearly planar. Figure 2: M-N-H....C torsional angles are 0.02° (Ti), −0.03° (Zr), and 0.004° (Hf).

The enthalpies of C-H activation (\(\Delta H^\text{calc}\)) for methane by (H)M=NH are calculated to be 14.5, 14.1, and 7.3 kcal mol\(^{-1}\) versus separated reactants for the Ti, Zr, and Hf systems, respectively. Table I. Replacemnt of the hydrido ligands in the Zr complex lowers the barriers to activation: 12.9 and 12.0 kcal mol\(^{-1}\) for the bis(amido)- and dichloro-Zr complexes, respectively, Table I. As mentioned, direct experimental data for the methane activation step by group IVB imidos is limited.\(^{14}\) Intramolecular C-H activation by the Ta=C double bond in TaO(2,6-(t-Bu)-4-X-phenyl)(CH3)\(^{23}\)=CH2 has \(\Delta H^\text{act} = 14.3 \pm 1.0\) kcal mol\(^{-1}\) (X = H) and \(\Delta H^\text{act} = 15.4 \pm 1.0\) kcal mol\(^{-1}\) (X = OMe).\(^{3}\) Based on the calculated \(\Delta H^\text{act} = 23.0 \pm 0.7\) kcal mol\(^{-1}\) (M = Ti); \(\Delta H^\text{act} = 21.6 \pm 1.0\) kcal mol\(^{-1}\) (M = Zr).\(^{33}\) Thus, the present calculations suggest that the inactivity found by Rothwell et al.\(^{15}\) is due to steric reasons, e.g., the unavailability of an open coordination site for the hydrocarbon or the inability of the hydrocarbon to displace py.

Combining the calculated \(\Delta H^\text{act}\) with the exothermicities quoted above shows that the reverse process, methane elimination, has a barrier (\(\Delta H^\text{elim}\)) of 23.4 (Ti), 34.1 (Zr), and 38.0 kcal mol\(^{-1}\) (Hf), depending on the metal for the dihydrido complexes, Table I. Published and unpublished data from Wolczanski et al.\(^{30,36}\) are consistent with the trends as well as the magnitudes, in the barriers to methane elimination as a function of the metal. Activation parameters for the elimination of methane from (Si'O)\(^2\)Ti-(NH'Si')(CH3)\(^{23}\) are \(\Delta H^\text{elim} = 21\) kcal mol\(^{-1}\) and \(\Delta S^\text{elim} = -11\) eu.\(^{16}\) Replacement of the hydrido ligands in the Zr complex with Cl and NH; ligands lowers and raises, respectively, the barrier to methane elimination by 13.8 kcal mol\(^{-1}\) and the reaction entropies by 24.9 eu.\(^{16}\) \(\Delta H^\text{elim}\) for (NH3)\(^2\)ZrCH3, measured over the temperature range 87.1-127.1 °C, is 25.9 (4) kcal mol\(^{-1}\), in reasonable agreement with the calculated values, given the approximations made. One last set of calculations with an augmented basis set (adding p polarization functions to H atoms) and correction of the calculated \(\Delta H^\text{elim}\) to 100 °C was carried out for (NH3)\(^2\)ZrCH3. The basis set improvement and temperature correction lowers the calculated barrier to methane elimination to 31.4 kcal mol\(^{-1}\), in improved agreement with experiment.

The amount of bond stretching and compression in the TS is approximately the same for each system. Figure 2, as determined by the percentage change from equilibrium values in reactants and products. Metal-nitrogen bonds are stretched by 3-4% versus those in metal-imido reactant; metal-carbon bonds are stretched by 8-10% compared to those in methyl(amido) product. C-H and N-H bonds are further from equilibrium values, being 28-31% and 41-46% larger than normal (r(CH) = 1.0 Å, r(NH) = 1.01 Å), respectively. The effect of modifying the ancillary ligands (X = H, Cl, NH3) on the Zr transition-state geometry is small. It seems plausible to assume that the ligands exert the majority of their influence through steric effects. The activation parameters for methane elimination from (NH3)\(^2\)ZrCH3 point to "substantial Zr-C bond breaking (\(\Delta H^* = 25.9\) kcal mol\(^{-1}\)) in a relatively constrained transition state (\(\Delta S^\text{act} = -7(1)\) eu) in the transition state.\(^{38}\) The calculated structural...
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Figure 2. Calculated transition-state geometries for the methane activation/elimination processes. The circles with diagonal, horizontal, and crossed lines are metal, chlorine, and nitrogen atoms, respectively. The open circle is hydrogen; the filled circle is carbon. TSs through c are C; d and e are C.

Figure 3. Plot of the changes in the important bond lengths and bond angles along the IRC for the activation of methane by (H)Zr=NH at MCSCF level.

As mentioned in previous work on H2 and CH4 activation by high-valent metal systems, the geometry of the transition state leads to a short metal-transannular hydrogen distance. On the average, the MH distance is only 1–2% shorter than the corresponding terminal hydride bond length, Figure 2. Despite the short diastereity in metal-hydrogen interatomic distances, the interaction (as estimated by the Mulliken bond overlap population (BOP)) is substantially less with H2 than with the terminal hydride ligands. The BOP(MH) for each TS is 0.18 (Ti), 0.13 (Zr), and 0.22 (Hf), while the average BOP for the metal-terminal hydride bonds in the TS is 0.74 (Ti), 0.68 (Zr), and 0.75 (Hf). If we assume that the metal-terminal hydride BOP corresponds to a single bond value, then the fractional bond order for the MH interaction is 0.24 (Ti), 0.19 (Zr), and 0.29 (Hf) (3).

Therefore, the short MH distance does not signify a formal single bond. However, there is a significant stabilizing interaction, as indicated by the positive sign of BOP(MH). As the C–H bond is stretched and activated, it seems reasonable to assume that H2 will have available bonding capacity, some of which is used in forming the N–H bond and some of which is made available to the metal to stabilize H2 as it migrates between C and N.

4. Intrinsic Reaction Coordinate for Methane by the Zr–Imido Complex. The intrinsic reaction coordinate (IRC) (22) is the lowest-energy path connecting the reactants and products and is thus of great interest to the chemist. The IRCs for methane activation by various imidiz at differing levels of theory are qualitatively similar. The primary points of interest can be gathered from an exposition of one example. A plot of the changes in pertinent bond lengths and bond angles along the IRC of (H)Zr=NH + CH4 → (H)Zr(NH)(CH3)(NMe2), calculated at the MCSCF level, is shown in Figure 3. The abscissa in Figure 3 is the mass weighted distance along the reaction coordinate (S_ab in bohr-amu1/2). The TS is set to S_ab = 0 bohr-amu1/2. For the four atoms which make up the four-center TS (M, N, H2, C) there are six internal coordinates (N–H, M–H2, C–H2, M–N–H, N–H–C, M–N–H–C), only five of which are of interest since the four atoms remain essentially coplanar (i.e., M–N–H–C = 0°). In addition to these five internal coordinates, the M–N–H angle is a good indicator of the progress of the nitrogen ligand in its transformation from an imido to an amido.

Proceeding from the extreme right (imido plus methane reactants), close inspection of the IRC shows that until S_ab ≈ 1, little structural change is taking place other than canting of the C–H bond to be activated toward the Zr=N bond upon which it will be broken. The C–H bond changes very little until S_ab = 1, after which it begins to lengthen considerably. From Figure 3, there are several motions along the IRC which seem to be correlated. The “break” point at S_ab = 1 for the C–H bond coincides with a change in direction of the N–H–C angle, which was unchanged up to this point on the reaction coordinate, begins to increase sharply. Plotting the geometries along the IRC reveals that these motions (from S_ab = 1 to S_ab = -1) correspond to the transfer of H2 from C to N in a fashion which is reminiscent of the motion of the central hydrogen in a A–H=B hydrogen bond. At S_ab = -1, a “break” point occurs for the N–H bond of the incipient amido, i.e., it has been formed and no further change is required. Note that the “break” point at S_ab = -1 is correlated with a sharp decline in the Zr=N–H angle, Figure 3. Since the Zr=N–H and Zr–N–H change at the same rate after S_ab = -1, Figure 3, the amido ligand is pivoting in place as it is rehybridized from sp (imido) to sp2 (amido). It is interesting that the imido ligand does not pivot until the NH bond is almost completely formed. After S_ab = -1, all important changes have taken place and all that is left is for the angles about Zr to relax to the tetrahedral geometry of the methyl(amido) product.

The two most interesting points to emerge from the IRC are the timing of N–H bond formation with the pivoting of the amido

(38) The calculated enthalpies of activation for methane elimination from the Zr–amido complexes are -9.4 (X = H), -5.2 (X = NH2), and -5.7 eu (X = CI) for an average of ~7 ± 2 eu, in reasonable agreement with the experimental value (~7 ± 1 eu), given the approximations.
ligand and the parabolic shape of the \( \text{MCH} \) curve along the reaction coordinate. If one considers the reverse reaction, extrusion of methane from Zr(\( \text{H}_2\))\( \text{(NH}_2\))\( \text{(CH}_3\))\( \text{H} \), the reaction occurs in several phases: the C–Zr–N bond angle compresses (step i), the amido ligand pivots (step ii), and then the N–H bond breaks and is transferred to the methyl group to eliminate methane (step iii in eq 1). At \( S_{\text{TS}} \approx -1 \), the Zr–N–H angle is \( \approx 87^\circ \) and Zr–N–H

\[ \approx 168^\circ \] with a H–N–H angle of 105°. Thus, at \( S_{\text{TS}} \approx -1 \), the Zr complex bears a striking structural resemblance to the iso-electronic (in terms of valence electrons) Ta=CH moiety in electron-deficient alkyldienes \( L\text{M}=(\text{C}=(\text{R})^3 \text{F}) \). These complexes have Ta=C–H angles close to 90°, as revealed by neutron diffraction. The distorted Ta=C–H angles and longer than normal \( \text{C–H} \) bonds suggest agostic C–H interactions with the metal and incipient C–H bond breaking. Calculated BOPs in the region of \( S_{\text{TS}} \approx -1 \) are one-third less for agostic N–H, compared to N–H\( \text{H} \), although the former is only 0.05–0.07 Å longer than the latter (N–H = 1.00 Å). The importance of agostic interactions in \( \alpha \)-elimination and \( \text{CH} \) activation processes involving metal–carbon bonds has been recognized since the phenomenon was first characterized. Calculations indicate that agostic interactions are also important in methane activation/elimination by their nitrogen analogues.

The second interesting point in the IRC concerns the behavior of \( \text{MH} \), along the reaction coordinate. The M–H distance decreases along the IRC from reactants toward TS, reaches a minimum in the vicinity of the TS, and then increases as methyldialkane products are formed, Figure 3. The minimum in the M–H distance coincides with the crossing of the NH and CH\( \text{H} \), curves, i.e., these latter distances are equal. The correspondence of this crossing point with the minimum in the MH\( \text{H} \) curve suggests that the M–H distance is strongest (i.e., the distance is shortest) when H is roughly halfway between N and C and, presumably, H\( \text{H} \), has its maximum available bonding power. This correspondence has been seen elsewhere and can be interpreted as additional evidence for the importance of the NH interaction in modulating the methane activation TS.

**Summary and Conclusions**

Several interesting points have resulted from this study of methane activation by group IVB imido complexes. Some of the more pertinent conclusions are summarized here.

1. The exothermicity of methane activation \( (\text{H}_2\text{M}=\text{NH}+\text{CH}_4 \leftrightarrow (\text{H}_2\text{M}(\text{NH}_2))\text{(CH}_3)) \) increases as the metal becomes heavier (H > Zr > Ti), while the activation barriers decrease in the opposite order (Ti > Zr > Hf) relative to their nitrogen analogues.

2. A growing body of evidence, gas-phase and solution-based, supports the presence of alkane complexes of the metal-imido "active site" for a longer period of time, weakening C–H by the transfer of electron density out of the C–H bonding region, and taking part in the aforementioned polarization phenomenon. A Mulliken population analysis indicates that \( \approx 0.2 \) electrons are transferred from \( \text{CH}_2 \) to the metal, inducing greater C\( \text{H}^\alpha \) polarization. It has been proposed that the metal-imido bond induces a polarization of the C–H bond, which aids in the C–H activation process. The present results support such a proposal and suggest that the alkane complex is the vehicle through which increased C\( \text{H}^\alpha \) polarization is set in motion, facilitating the subsequent rupture in a concerted, heterolytic C–H cleavage.

3. The enthalpies of C–H activation \( (\Delta H^\alpha) \) for the addition of a methane C–H bond across the M=NH bond of \( (\text{H}_2\text{M}=\text{NH}) \) are calculated to be 14.5 (Ti), 14.1 (Zr), and 7.3 kcal mol\(^{-1} \) (Hf) above those of separated reactants, Table I. Although experimental evidence for this process is limited, the calculated \( \Delta H^\alpha \) values for the metal-imido complexes are in reasonable agreement with estimates available from metal–carbon complexes. As noted above, \( \Delta H^\alpha \) shows the expected correlation with \( \Delta H_{\text{TS}} \), i.e., \( \Delta H^\alpha \) decreases as \( \Delta H_{\text{TS}} \) becomes more exothermic.

4. Calculated \( \Delta H^\alpha_{\text{TS}} \) are 22.4 (Ti), 24.1 (Zr), and 38.0 kcal mol\(^{-1} \) for the diphosphido complexes. Data from Wolczanski et al.\(^{8a,16} \) are supportive of the trends, as well as the magnitudes, in \( \Delta H^\alpha_{\text{TS}} \) as a function of the metal. \( \Delta H^\alpha_{\text{TS}} = 25.9(4) \) kcal mol\(^{-1} \) for methane elimination by (\( \text{NH}_3\))\( \text{ZrCH}_3 \). Methane elimination from (\( \text{NH}_3\))\( \text{ZrCH}_3 \) is \( \Delta H^\alpha_{\text{TS}} = 100(3) \) kcal mol\(^{-1} \) at the highest level of theory. Higher levels of correlation will probably lower \( \Delta H^\alpha_{\text{TS}} \) since the TS is more sensitive to improvements in correlation and basis set due to the presence of bonds which are far removed from equilibrium values.

5. The amount of bond stretching and compression in the TS versus equilibrium values is approximately the same for all systems studied. The calculated geometries of the activation/elimination TSs indicate a late transition state for methane elimination (and thus an early TS for the reverse process), in agreement with experimentally determined activation parameters.\(^{8a,16} \)

6. The IRC analysis indicates that NH bond activation occurs after pivoting of the amido ligand if the reaction is reviewed as the microscopic reverse of methane activation, i.e., \( \alpha \)-elimination of methane. After pivoting of the amido ligand occurs but before H-transfer from N to C is significant, the Zr-amido geometry at this point on the reaction coordinate resembles that of the Ta=CH fragment in Schrock's electron-deficient Ta=alkylidene. An analysis of the Mulliken bond overlap population indicates that the agostic N–H is significantly weaker than the H-transfer from N to C is significant, the Zr-amido geometry at this point on the reaction coordinate resembles that of the Ta=CH fragment in Schrock's electron-deficient Ta=alkylidene. An analysis of the Mulliken bond overlap population indicates that the agostic N–H is significantly weaker than the

7. The Mulliken bond overlap populations (BOPs) indicate a stabilizing interaction between the metal and the transannular H (H\( \text{H} \)) in the TS for methane activation/elimination by group
Incorporation of Hydration Effects within the Semiempirical Molecular Orbital Framework. AM1 and MNDO Results for Neutral Molecules, Cations, Anions, and Reacting Systems

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Abstract: A recently developed continuum model for the incorporation of hydration effects into the AM1 and MNDO semiempirical molecular orbital Hamiltonians is evaluated by comparison with representative experimental data. The new method gives quantitatively correct predictions for the aqueous-phase conformational behavior of H_{3}N^{+}CH_{2}CH_{2}CO_{2}^{-} and the aqueous-phase activation energetics of the SN_{2} reaction of Cl^{-} and CH_{3}Cl. The dissociation energetics of r-BuCl are correctly predicted, while the computed dissociation pathway includes a shallow energetic minimum corresponding to the expected tight ion pair. The absolute hydration energies of neutral molecules and cationic and anionic species are generally predicted satisfactorily. Systematic deviations from the experimental data due to the intervention of hydrophobic effects (not currently allowed for) are ascribed to the general insufficiency of the AM1 and MNDO molecular orbital Hamiltonians in representing these classes of structures. Such deviations can be reduced, or eliminated, by careful empirical parametrization of the model.

Introduction

The reaction field approach has attracted much recent attention as a general basis for the computation of solvation energies.\(^1\)\(^2\) Despite the crudeness of the model, in which the solvent is treated as a structureless dielectric continuum, it appears to offer considerable promise in chemical, biochemical, and pharmacological applications. To date, development of the theory within the quantum mechanical framework has evolved in three distinct directions. In one, the solute is placed in an ellipsoidal or often a structureless dielectric cavity within the solvent with which it interacts via its dipole or in some cases higher-order multipole moments.\(^3\)\(^-\)\(^5\)

Although this method has enjoyed considerable popularity, there are obvious limitations associated with the need to constrain an arbitrarily shaped solute to a cavity of specified shape. For similar reasons, its application to reaction processes is highly limited. In the so-called extended Born approach, the solute is treated as an array of atom-centered spheres, each of which contributes to the overall hydration energy in a manner related to Born's original treatment of monatomic ions.\(^6\)\(^-\)\(^12\) A number of aplici