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Contractor for the U.S. Department of Energy
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Summary

The formation of high sludge strength agglomerates is a key concern to the *Sludge Treatment Project* (STP) to ensure the sludge can be retrieved after planned storage for up to 10 years in *Sludge Transport and Storage Containers* (STSC) at T Plant. This report addresses observations of agglomerate formation, conditions that the data shows lead to agglomeration, the frequency of agglomerate formation and postulated physiochemical mechanisms that may lead to agglomeration. Although the exact underlying chemistry of K Basin sludge agglomerate formation is not known, the factors that lead to agglomeration formation, based on observations, are as follows:

- High Total Uranium Content (i.e., sample homogeneity and influence from other constituents)
- Distribution of Uranium Phases (i.e., extent of conversion from uraninite to uranium oxide hydroxide compounds)
- Sample Dry-out (loss of cover water)
- Elevated temperature
- Solubility of U(IV) phases vs. U(VI) phases
- Long storage times

Agglomerated sludge has occurred infrequently and has only been observed in four laboratory samples, five samples subjected to hydrothermal testing (performed for 7 to 10 hours at $\sim 185^{\circ}\text{C}$ and 225 psig), and indirectly during six sampling events in the KE Basin. In the four laboratory samples where agglomerates were observed, the agglomerates exhibited high shear strength and the sample container typically had to be broken to remove the solids. The total uranium content (dry basis) for the four samples (KE Pit, KC-2/3 SS, KC-2/3 M250 and 96-13) were ~ 8 wt%, ~ 59.0 wt%, 68.3 wt% and 82 wt%. The agglomerates that were present during the six sampling events were undoubtedly disturbed and easily broken apart during sample collection, thus no agglomerates were observed in subsequent laboratory analyses. The highest shear strengths measured for K Basin sludge samples were obtained after hydrothermal treatment (7 to 10 hr at 185°C) of high-uranium-content KE canister sludge. The unconfined compressive strength of samples from this testing, measured by a pocket penetrometer, infers that their shear strength may be between 120 kPa and 170 kPa (PNNL-16496). These short-duration hydrothermal tests were conducted at temperatures much greater than the temperature of the T Plant canyon cells (-7°C to 33°C); however, the strength results provide an initial bounding target for sludge stored for many years, and an upper range for simulants (042910-53451-TP02 Rev 1).

Sampling and characterization activities conducted in 2009 have measured the total uranium content and speciation for sludge stored in Engineered Containers SCS-CON-220, -240, -250, and -260 (PNNL-19035). Based on on-going testing that has measured the shear strength of uranium samples containing varying uranium (IV) to uranium (VI) ratios and the characterization of the Engineered Containers SCS-CON-220, -240, -250, and -260, it is unlikely that agglomerates will form on a large scale in this sludge. The highest measured total uranium concentration in the Engineered Container SCS-CON-220 sludge is 35.2 wt% and only 4 wt% to 6 wt% (dry) in Engineered Containers SCS-CON -240, -250, and -260. The uranium concentrations in Engineered Containers SCS-CON-220, -240, -250, and -260 sludge are below the threshold for agglomerate formation. Settler sludge however is estimated to contain ~ 80 wt% (dry) total uranium, which could lead to the formation of high strength agglomerates depending on the relative concentrations of U(IV) and U(VI) compounds.

One of the chief concerns of the STP is sludge dry-out. Samples archived in PNNL hot cells have been known to dry out and form hard clods of material, which are then difficult to reconstitute (HNF-6705). In 1996, all but one of the samples archived at the 222-S Laboratory dried out. These samples were composed of sludge collected from the KE Basin floor and Weasel Pit. However, in the STP's current design plans for sludge stored in STSCs at T Plant, there are provisions for continual water level observation and periodic water replenishment when needed, which dramatically lowers the risk of a dry-out event.

Recent 2009 KE Engineered Container sludge and KW Engineered Container sludge physical characterization lead to opportunistic data on the behavior of dried sludge. As shown in section 5 of this report, sludge core samples taken from Engineered Container 220 that have inadvertently dried after characterization activities indicate that the sludge is mostly composed of very "weak" and friable solids. Since Engineered Container SCS-CON-220 sludge has the highest total uranium (uranium metal and uranium oxide) concentration of the four Engineered Containers sampled to date, if a dry-out event should occur in the STSC while stored at T Plant there is qualitative evidence to suggest that the solids would be "weak" and friable in nature and thus easily mobilized. Information will need to be obtained during characterization of Settler sludge to determine characteristic of this sludge upon drying.

Acronyms/Abbreviations:

d_p	Particle Diameter
NLOP	North Loadout Pit
PNNL	Pacific Northwest National Laboratory
SEM	Scanning Electron Microscopy
SNF	Spent Nuclear Fuel
STP	Sludge Treatment Project
TEM	Transmission Electron Microscopy
UOH	Uranyl Oxide Hydrate
XRD	X-Ray Diffraction

Mineral Chemical Formulas:

Albite	$\text{NaAlSi}_3\text{O}_8$
Gismondine	$\text{CaAlSi}_2\text{O}_8 \cdot 4\text{H}_2\text{O}$
Becquerelite	$\text{Ca}(\text{UO}_2)_6\text{O}_4(\text{OH})_6 \cdot 8(\text{H}_2\text{O})$
Meta-Schoepite	Formula = $\text{UO}_3 \cdot 2\text{H}_2\text{O}$ Phase = $(\text{UO}_2)_4\text{O}(\text{OH})_6 \cdot 5\text{H}_2\text{O}$
Schoepite	Formula = $\text{UO}_3 \cdot 2.25\text{H}_2\text{O}$ Phase = $(\text{UO}_2)_4\text{O}(\text{OH})_6 \cdot 6\text{H}_2\text{O}$
Uraninite	UO_2 , U_4O_9 , and U_3O_7
Uranophane	$\text{Ca}(\text{UO}_2)_2(\text{SiO}_3\text{OH})_2 \cdot 5\text{H}_2\text{O}$

Table of Contents

1.0 Purpose	1
2.0 Introduction.....	1
3.1 Observations of Agglomerates in Storage and Testing.....	2
3.2 Observations of Hardpan Layers in Sampling Campaigns	7
3.3 Frequency of Agglomerate Formation.....	8
3.4 Conditions that Lead to the Formation of Agglomerates.....	11
3.5 Postulated Physical and Chemical Mechanisms of Agglomerate Formation	14
4.0 Agglomerated Sludge Mobilization.....	22
4.1 Shear Strength of Varying U Oxide Mixtures	23
4.2 Evaluation of Jet Erosion.....	25
5.0 Physical Properties of Dried KE Engineered Container sludge and KW Engineered Container sludge	26
6.0 Conclusion	27
7.0 References.....	31
8.0 Bibliography of Reviewed Papers	33
Appendix A Compaction and Evaporation Test sample characteristics (PNNL-15092).....	34

List of Figures

Figure 1 Agglomerated Sample 96-13	3
Figure 2 Agglomerated KC-2/3 M250 sample.	5
Figure 3 Possible uranium alteration phases.....	7
Figure 4 Solubility curve for U(IV) and U(VI) compounds	13
Figure 5 Electron Micrograph of Agglomeration Process	15
Figure 6 TEM image of Metaschoepite Formation.....	16
Figure 7 Agglomeration of UOH fractal aggregates at the air-water interface.....	16
Figure 8 Shrinking Core Kinetics for uraninite oxidation	18
Figure 9 Illustration of Agglomerates Formation	19
Figure 10 Correlation of total uranium (dry basis)	21
Figure 11 Correlation of Shear Strength with U(VI) concentration (Dry basis).....	25
Figure 12 Post-Test Examination of Physical Property Samples from SCS-CON-220 Core Samples	26

List of Tables

Table 1 Summary of Sludge Samples of K Basins Sludge that have been collected.....	9
Table 2 Sludge Archive Samples for Compositing, RE-Jarring, Blending and Inventory	10
Table 3 KW Engineered Container sludge dry weight % total uranium.....	20
Table 4 KE Engineered Container sludge dry weight % total uranium.....	20
Table 5 Uranium Speciation Data by Engineered Container Core Composite Sieve Fraction.....	20
Table 6 Results of mixed oxide shear strength in a pure uranium system.....	24

1.0 Purpose

The purpose of this document is to consolidate existing data associated with the formation of high strength K Basin sludge agglomerates, as discussed in *Sludge Treatment Project Engineered Container Retrieval and Transfer System Technology Maturation Plan* (STP-PRC-00113, pg 33). A separate report will be issued to address sludge chemistry and rheology information on sludge aging mechanisms. Past observations and frequency of agglomerate formation, conditions that lead to agglomeration and postulated physical and chemical mechanisms by which sludge agglomeration occurs are discussed. These key subjects are investigated using available quantitative data and observations collected during sludge testing and sampling campaigns.

2.0 Introduction

Agglomeration of K Basin sludge particles is of particular interest to the Sludge Treatment Project (STP) due to the expected long-term storage in water filled Sludge Transport and Storage Containers (STSCs) in T-Plant canyon cells for up to 10 years. T Plant will be used as interim storage while the STP considers final treatment methods for disposing of the K Basin sludge as remote-handled transuranic waste. Thus, it is important to understand the effects of long-term storage (e.g., agglomeration/compaction, rate of oxidation, volumetric expansion, etc) will have on the K Basin sludge in order to prepare for sludge retrieval prior to treatment and disposal after the prescribed storage period. The focus of this report is to consolidate information on agglomerate formation in K Basin sludge. For this report agglomerates will be defined as sludge of sufficient strength such that mobilization and transfer of the sludge is significantly impeded (i.e., requires container to be disassembled/fractured to facilitate sample recovery).

The sludge storage conditions in the STSC differ significantly from that of the KW Basin pool. In the basin pool water is chilled (10 °C to 15°C), filtered (sand filter) and passed through ion exchange beds (containing both anion and cation exchange media). T Plant can experience canyon temperatures ranging from -7°C to 32°C (HNF-SD-SNF-TI-015, Vol 2, Rev 14A, Table 4-25). However, the temperature swing in the T Plant storage cells is expected to be less severe due to the cells being partially below ground and the insulating properties of the thick concrete cell walls and ground. As a result, between February 2008 and February 2010, the KW Basin pool pH has been near neutral (4.6 to 6.0) and the conductivity has been low (1.15 $\mu\text{S}/\text{cm}$ to 2.80 $\mu\text{S}/\text{cm}$). Once loaded into the STSCs, the sludge and supernatant will be essentially a closed system, with the STSCs vented to the T Plant cell air atmosphere. The conditions at PNNL, under hot cell storage, are expected to be similar to those anticipated at T Plant and better predictors of sludge behavior in STSCs than the in-basin observations.

3.0 Discussion

Observations of KE Basin sludge samples forming agglomerates during hot cell storage and testing are discussed in section 3.1. Observations of agglomerate and hardpan formation during sludge sampling campaigns are discussed in section 3.2. The frequency of agglomerate formation among sludge samples collected from the K Basins, conditions that lead to agglomeration and postulated chemical and physical mechanisms of agglomerate formation are discussed in section 3.3, 3.4 and 3.5, respectively.

3.1 Observations of Agglomerates in Storage and Testing

The most relevant study of sludge behavior during storage, and a significant instance of agglomeration among samples collected from the KE Basin occurred during a 28-month (May 2002 to September 2004) sludge compaction test performed by PNNL at hot cell conditions (i.e., 25 °C to 35 °C and a slight vacuum). The test population consisted of six KE Basin sludge samples with test identification numbers 96-13, FE-5, KC-2/3 whole, KC-4, KC-5, and SNF Comp (PNNL-15092). The six samples were collected from the KE Basin during sampling campaigns that took place in FY 1996 and FY 1999. The 96-13 sample was collected from a single location; the other samples were composite samples, made up from sludge collected from multiple locations in the KE Basin. Appendix A provides relevant information, such as starting and ending uranium phases, total uranium content and identity (i.e., location of collection) for these samples. The testing was performed to quantify the amount of sludge compaction and loss of supernatant that can be expected during long-term storage at T Plant.

In preparation for the compaction and evaporation test, the six samples were placed in graduated cylinders and excess KE Basin water was added to cover the sludge. The addition of KE Basin water was performed periodically during the 28-month compaction and evaporation test to ensure the sludge was covered at all times. Of the six samples in the study, only sample 96-13 which is composed of sludge collected from a single fuel storage canister (canister constructed of aluminum, containing fuel in poor condition) showed significant formation of high strength agglomerates and actually self-cemented in the graduated cylinder, as seen in Figure 1. The agglomeration and self-cementation was so severe, the glass graduated cylinder had to be broken with the hot cell manipulator arm to recover the sample. The sample was then sheared into pieces with considerable, but not quantified force by the manipulator arm and analyzed. The remaining five samples (FE-5, KC-2/3 whole, KC-4, KC-5, and SNF Comp) were readily re-suspended in their respective supernatant at the conclusion of the 28-month study, and there was no evidence of agglomerate formation.

It must be noted that upon completion of this test, sample 96-13 was sheared into small chunks and stored in water within jars in 2004. The sample was then re-examined in July 2009, where the well defined agglomerates that were present ~ 6 years earlier had collapsed to form a puck at the container bottom, with the material having undergone oxidation [from blue/black, indicative of U(IV) phases, to green/yellow, indicative of U(VI) phases] (July 2009 Monthly Report, PNNL).

In addition, prior to the *28-month Compaction and Evaporation* test, sample 96-13 experienced a complete loss of supernatant despite efforts to keep sludge samples covered with water at all times. Six months prior to the initiation of the 28-month test, sample 96-13 was reconstituted by addition of water and mixing. In addition to the dryout and agglomeration experienced by sample 96-13, gas bubbles

were observed once in the supernatant of this sample during the compaction and evaporation study indicating that oxidation was ongoing and uranium metal was present (PNNL-15092, pg 3.1.2). The presence of hydrogen gas evolving from the reaction of uranium metal and water was speculated to be the cause for discontinuities (i.e., visually interpreted as voids) seen in the sludge sample in Figure 1. Uranium metal oxidation generates heat as a reaction by-product, which may have contributed to the previous loss of supernatant from the sample.

Significant gas generation was observed in the SNF Comp sample (KE floor and canister sludge + fuel fragments), however agglomerates were not observed during this testing. No gas generation was observed in samples FE-5, KC-2/3 whole, KC-4, or KC-5, during the testing, even though sample KC-2/3 whole was previously determined to have a uranium metal content of ~1.9 wt% (PNNL-13320).

Sample 96-13 was found to have uncharacteristically high total uranium content (82 wt%-dry) compared to the sludge that will be stored in the STSCs at T Plant (i.e., Engineered Container sludge). Two other samples, KC-2/3 whole and SNF Comp, also had relatively high total uranium concentrations of 59.0 wt% (dry) and 60.2 wt% (dry), respectively, but they did not experience the formation of high strength agglomerates or undergo self-cementation.

Figure 1 Agglomerated Sample 96-13

Left photo is sample in a 3.5 cm inner diameter graduated cylinder after settling test and right photo is sample after being removed from the graduated cylinder¹.

¹ Figure courtesy of PNNL-15092, pg 3.10

X-ray diffraction (XRD) analyses (HNF-SP-1201, pg 29, Table 3.4) performed about 5 years prior to the *Compaction and Evaporation* testing indicate that the primary uranium compounds present in sample 96-13 in 1997 were tetravalent uraninites of the formula UO_2 , U_4O_9 , and U_3O_7 . XRD analysis on KC-2/3 whole indicated the presence of tetravalent uraninites (U_4O_9) with some schoepite (PNNL-13320). XRD analyses were not performed immediately before the start of the compaction study. Therefore, the uranium compounds present at the time of the *Compaction and Evaporation* test are not precisely known. Subsequent XRD analyses performed in FY 2007, 3 years after the *Compaction and Evaporation* study, indicate that the uraninite compounds in samples 96-13 and KC-2/3 whole had oxidized to metaschoepite and schoepite, both of which are UOH compounds (PNNL-17078). No uranium phases were identified in sample KC-5. Sample SNF-Comp contained a combination of KE floor and canister sludge, plus uranium metal, and likely contained some tetravalent uraninites. Samples FE-5, KC-4, were not analyzed by XRD at the time of collection, so the phases present at that time are not available for comparison. The results from XRD scans for the samples analyzed after long-term storage of up to 8 years in the PNNL hot cells demonstrate that the tetravalent uraninite compounds have undergone oxidation to form hexavalent UOH compounds (i.e., schoepite and metaschoepite and in some cases becquerelite, and uranophane) as seen in Appendix A.

Similar to sample 96-13, KC-2/3 M250 also exhibited agglomeration and self-cemented at PNNL hot cell temperatures (25°C to 35°C). KC-2/3 M250 is a subsample of KC-2/3 whole (KE Canister sludge composite used in the *Compaction and Evaporation* testing) and was prepared by wet-sieving sample KC-2/3 whole to provide a sludge fraction that is $< 250\ \mu\text{m}$, prior to being stored in the PNNL hot cells from 1999 to 2007, a period of ~ 8 years. This sample was discovered to have agglomerated and self-cemented in its storage jar in 2007 when it was planned for use in *Hydrothermal* testing as a surrogate for Settler sludge. Sample KC-2/3 M250 was also found to have a relatively high total uranium content of 68.3 wt% (dry). Similar to sample 96-13, sample KC-2/3 M250 exhibited significant strength and could not be retrieved from the storage jar without breaking it (PNNL-16496, pg 2.3). The sample was removed from the broken jar in one cemented rigid column and subsequently sheared into pieces before being blended (homogenized). As seen in Figure 2 (a), sample KC-2/3 M250 looks very similar to sample 96-13 (Figure 1). This sample was analyzed by XRD and the main phases present were identified as hexavalent uranium compounds schoepite and metaschoepite.

In an attempt to understand the effectiveness of the homogenization and the size of remaining agglomerates, a 6 g sub-sample of the agglomerated KC-2/3 M250 sludge (after the parent sample was homogenized with a blender) was wet-sieved with $500\ \mu\text{m}$ and $1000\ \mu\text{m}$ screens, as seen in Figure 2 (b). Several large agglomerates, $1000\ \mu\text{m}$ to $3000\ \mu\text{m}$ in diameter were retained on the $1000\ \mu\text{m}$ screen as seen in Figure 2 (c). These agglomerates however disintegrated easily with slight pressure from a flexible plastic pipette and therefore would be expected to have negligible shear strength (PNNL-16496, pg 2.3).

(a)

(b)

(c)

Figure 2 Agglomerated KC-2/3 M250 sample.

(a) Agglomerated KC-2/3 M250 sample after being removed from the graduated cylinder. (b) 6 g sub-sample of KC-2/3 M250 after homogenization prior to wet sieving. (c) After sieving the 6 g sample and washing with water, agglomerates 1000 µm to 3000 µm can be seen on the 1000 µm screen².

² Figure courtesy of PNNL-16496, pg 2.4

KC-2/3 SS (the settled solids fraction of KC-2/3 whole consisting of 59.0 wt% U total on a dry basis) experienced agglomeration while in PNNL hot cell storage. Like samples 96-13 and KC-2/3 M250, the sample jar had to be broken to extract the sludge. Unfortunately, no XRD analysis could be located for sample KC-2/3 SS.

KE Pit (Weasel Pit sludge) also experienced agglomeration while in storage in the PNNL hot cells. Like the three previous samples (96-13, KC-2/3 M250 and KC-2/3 SS), the sample jar had to be broken to remove the sludge. The KE Pit sample is a composite of sludge samples obtained from the KE Basin Weasel Pit and is estimated to contain 8 wt% total uranium (dry). To prepare the sample individual KE Basin Weasel Pit samples were dried and mixed with water to form this composite.

XRD analyses performed in 1995 found no crystalline uranium phases present in KE Pit, however other crystalline phases (quartz, albite, lepidocrocite, goethite, calcite, talc, rutile, and gismondine) were observed. KE Basin historical knowledge and XRD results of KE Weasel Pit sample KES-H-08 (WHC-EP-0877) support the presence of mordenite (chemical name for Zeolon-900 ion exchange material). Even though mordenite was not detected in the KE Pit sample, the zeolite compounds albite ($\text{NaAlSi}_3\text{O}_8$) and gismondine ($\text{CaAl}_2\text{Si}_2\text{O}_8 \cdot 4\text{H}_2\text{O}$) were detected. These two zeolite phases could have originated from the ion exchange material (mordenite), as it is known to dissolve and re-precipitate in different chemical forms under alkaline conditions (PNNL-12108). This is significant because poor water quality in the KE Basin during the years prior to the collection of the KE pit sample could have caused the basin water to become alkaline (WHC-EP-0877). Therefore, the presence of these phases may explain why KE Pit formed agglomerates when no uranium crystalline phases (e.g., metaschoepite and schoepite) were detected. It should be noted that amorphous uranium phases could have been present at the time of testing.

Other instances of high strength agglomerate formation have been observed during hydrothermal testing performed in 2007. As a part of the hydrothermal testing, five sludge samples were subjected to hydrothermal corrosion via step-wise heating and pressure increases over a period of 7 hours to 72 hours until target conditions of 185°C and 225 psig were achieved. The five tests consisted of sludge samples composed of the following constituents:

- Test 1: Fine KE Canister Sludge and Fine Uranium Metal Fuel Particles
- Test 2: Fine KE Canister Sludge Composite and Uranium Metal Coupons
- Test 3: KE Container Sludge + Flocculant + Fuel Particles
- Test 4: KE Container Sludge + Flocculant
- Test 5: KE Container Sludge + Organic Ion Exchange Resin Beads

Even though the hydrothermal testing conditions (temperature and pressure) far exceeded what the KE Engineered Container sludge, KW Engineered Container sludge, or Settler sludge will be subjected to while stored at T Plant, the testing does provide some valuable information about the uranium alteration phases that can be produced at elevated temperatures. The diagram below (Figure 3) shows the possible uranium phases that can be produced when treated at elevated temperatures (~185 °C) in addition to when stored at hot cell conditions (~ 25 °C to 35 °C). The conditions at T Plant are similar to those in the PNNL hot cells therefore the same phases can be expected at T Plant as those indicated by the solid reaction pathways. The dashed reaction pathways show products formed under elevated temperatures.

Figure 3 Possible uranium alteration phases³

3.2 Observations of Hardpan Layers in Sampling Campaigns

Agglomerated sludge has been observed six times in the KE Basin during the FY 1995 and FY 1999 sampling campaigns in both the Weasel Pit and NLOP locations. The agglomerated sludge was reported as “hardpan” or “hard packed” material. Three cases of “hardpan” or “hard packed” material of sludge were observed in the 1995 sampling campaign and discussed in WHC-SP-1182 and three additional cases of “hardpan” material were observed in 1999 and are discussed in HNF-4746.

In 1995, “hard packed” material was observed while collecting sludge samples KES-R-18, KES-S-19 and KES-T-20 from the Weasel Pit. The “hard packed” material however was broken up easily in all three cases and the samples were collected without difficulty. During the collection of samples from the Weasel Pit in 1999, one of the sampling isolation tubes could not penetrate a layer of material at a sludge depth of 9 (in), while collecting sample KE-9. The basin floor was known to be another 2 (ft) down (HNF-4746, pg 5). It is not clear whether this impenetrable layer was a result of hardpan formation or simply the case of equipment (pumps, piping, etc.) buried in the sludge. In another instance, layers of hardpan material were found at a level of ½ (in) from the Weasel pit floor while collecting sample KE-10 in 1999 (HNF-4746, pg 5). Another observation of hardpan occurred when sampling with isolation tubes in the center of the NLOP in 1999. It was observed that the last 10 (in) of sludge was “crusty”, while collecting sample KE-5 (HNF-4746, pg 5).

While evidence suggests the presence of agglomerates in these six samples, upon collection the agglomerates were undoubtedly disintegrated and thus were not observed during subsequent laboratory analyses. Furthermore, it is unclear whether the formation of hard layers or hardpans observed when sampling in the KE Basin have any relation to the agglomerates seen in archived samples or sub-samples

³ Figure courtesy of PNNL-16496, pg 3.23.

used for testing. No data suggest that the formation of hardpans in the Basins is directly linked to the formation of agglomerates of hexavalent UOH particles that has occurred in the PNNL hot cells. These two phenomena (i.e., “hardpan” and “hard packed” layer formation and agglomeration of UO_2 oxidation products) may be two entirely separate instances of anomalous KE Basin sludge behavior.

Recent 2009 KW Engineered Container sampling shows no evidence of significant agglomerate or “hard pan” formation. Only slight resistance was encountered when inserting the isolation tube in Engineered Container 220 sludge.

3.3 Frequency of Agglomerate Formation

A total of 88 core samples have been collected from the K Basins since 1993, 77 of which were collected before sampling activities in 2009. The remaining 11 core samples were collected from the KW Engineered Containers in 2009. Many of these samples were sub-sampled and/or combined into composite samples for various tests and analyses.

A summary of the samples collected from the K Basins between 1993 and 2009 is provided in Table 1. Sample material collected from the KE NLOP in 1993 was not archived and provides little insight into sludge agglomeration. The samples collected between 1995 and 2009 were sub-divided in the analytical laboratories (either 222-S Laboratory or the Radiological Processing Laboratory (RPL)), and a significant number of these sample splits were placed into archive storage. In some cases, multiple sample splits were made from parent sample materials (e.g., parent sample KC-2/3 from 1999 has been divided into more than 10 sub-samples).

A systematic assessment of all K Basin sludge samples for evidence of agglomeration has not been performed. However, in 2007, a large number of archive samples were examined, disturbed (i.e., moved during examination), and repackaged as part of an effort to create larger quantities of test material. Table 2 was compiled from the results of the 2007 sludge sample and sub-sample inspection (PNNL-17078). Table 2 shows the samples that have formed high strength agglomerates (highlighted in red) while stored in PNNL hot cells along with their respective origins. Of the 77 samples (88 total samples – 11 2009 KW samples) considered, only 4 samples (96-13, KC-2/3 M250, KC-2/3 SS and KE Pit) or composited sub-samples (See Table 2) exhibited agglomeration as defined in Section 2.0.

Table 1 Summary of Sludge Samples of K Basins Sludge that have been collected

Sampling Campaign	Year Samples Taken	Number of Samples	Approximate Volume of Sludge Samples (liters)
1995- KE Floor and Weasel Pit	1995	20	4
1996 – KE Basin fuel canister	1996	9	2.4
1996 KW Basin fuel canister	1996	8	0.1
1999 Core samples Consolidated samples	1999	11 ^(a) 6 ^(a)	22
2003 One core sample KE NLOP	2003	1	1.4
1993 Core samples from KE NLOP ^(b) USQ sampling for criticality	1993	13 ^(b)	N/A
1998 Wash sludge components (coatings & internal sludge) examined from fuel elements ^(c)	1998	9 ^(c)	N/A
2009 Engineered Container Sludge (KE Originating Sludge, SCS-CON-240, -250 and -260)	2009	6	5
2009 Engineered Container Sludge (KW Originating Sludge, SCS-CON-220)	2009	5	1.4
^(a) These samples were prepared by combining sludge from multiple sampling locations ^(b) 1993 samples from KE NLOP pit were consumed/disposed, and campaign is not considered useful with respect to evaluation of agglomeration. ^(c) Samples from this campaign were small and received at PNNL as dry material.			

Table 2 Sludge Archive Samples for Compositing, RE-Jarring, Blending and Inventory⁴

Sludge Sample	Sludge Sub-sample	Condition	Approx. Vol., ml
KC-2/3 Comp	KC-2/3 Settling Study	Mushy, thick mud, black, no hardpan, with water	100
	KC-2/3 Whole	Stiff saturated mud, green, with water, easily dispersed with water	70
	KC-2/3 SS	Very stiff, glass broken to collect sample, with water	100
	KC-2 SSOL B	Very little solids with abundant water	5
	KC-3 + KC-2 Rec 2	Two jars with similar labels; solids wet but stuck to bottom	40
	KC-2/3 M250	Hard chunks, very black, with water, easily dispersed with water	40
	KC-2/3 P250	Can be scooped out	25
	KC-2/3 M250	Very good with abundant water; source material for STP Tests 1 and 2	40-100
KC-4 Whole	KC-4	Abundant water	50
	KC-4-L P6	Sludge plug recovered, abundant water	70-100
	KC-4 Dup LPG	Well settled solids, with water	30-50
	KC-4	Dry, but black, some retrieved as puck and some as powder	30-40
	KC-4 Rec	Soft and readily suspendible	40
KC-6	KC-6	Fully saturated, beads are clumped but readily break up	100
	KC-6 SSOL A	Some clumping but fully wet	125
KE Floc Comp	KE Floc Sludge (floc'd.)	Saturated, original plug fell apart with shaking	200
	KE Floc Comp (not floc'd.)	Saturated, original plug fell apart with shaking but not as easily as flocculated sludge	70
KC-1 M500	KC-1 M500	Stiff but readily sluiced, suspended, and re-settled	50-70
FE-5 Comp 1	FE-5 Comp 1	Reddish brown, can be penetrated and scooped	200
KC-4-2	KC-4*	Dry chunks, water added and chunks broke up	300
KC Floor Comp	KC Floor Comp	Wet, saturated, readily suspendible	30-50
KE Pit	KE Pit	Wet chunks, jar broken to recover	150-200
96-01	96-01	Very black solids	15
96-05	96-05	Big chunks	80
96-13 KE Comp A	96-13 KE Comp A	May contain some glass	125
96-08 SSOL		Watered, dry before	50
96-15 SSOL		Watered, dry before	100
96-09 SSOL		None	50
96-11 SSOL		Very little sample	<50
96-13 Settler Study		Not watered, wet	125
96-13 Solid Grad		Solids not watered, not wet	100
96-13 SSOL		Not watered	50
KE NLOP Jar #1		Good, with water	165
KE NLOP Jar #2		Good, with water	240
KE NLOP Jar #3		Good, with water	240
KC-6 carboy		Good condition; 6-liter	6400
KE Container Comp		Moist, 70 ml supernatant	25
KE Container Comp Floc		Moist, 30 ml supernatant	10

⁴ Table 1 courtesy of PNNL-17078 (53451-RPT01, 2007)

Due to the relatively short time the 2009 KW Engineered Container samples have spent in PNNL hot cell storage (i.e., conditions similar to those expected at T Plant), these samples were not considered in the determination of the frequency of sludge agglomeration.

It should be noted that the > 100 sludge subsamples stored in the hot cells are visually inspected routinely at least twice a year and water is added as needed to prevent sample dry-out. However, unless required for testing, the samples are not disturbed or mixed. Therefore, it is highly likely that many of these undisturbed samples have consolidated into partially agglomerated forms over time. In addition, many samples collected from the K Basins never had enough time in storage to form agglomerates because they were either consumed in sample characterization testing or inadvertently dried out in the laboratory. With only 4 sludge samples and sub-samples forming agglomerates that impede the removal of sludge, these data provide a high level of optimism that most sludge materials stored for many years at conditions similar to those expected at T Plant can be readily re-suspended and mobilized.

3.4 Conditions that Lead to the Formation of Agglomerates

The data suggest that the formation of high strength agglomerates may have several key contributing factors. These factors include, but may not be limited to, total uranium content (i.e., sample homogeneity and influence of non-uranium compounds), uranium oxide composition (i.e., extent of conversion of tetravalent uraninite to hexavalent UOH compounds), sample dry-out, temperature, solubility, and storage time. Upon review of the available reports on K Basin sludge sampling and characterization, only samples 96-13, KC-2/3 M250, KC-2/3 SS, and KE Pit formed high strength agglomerates at conditions similar to those expected in T Plant. For the agglomerates of these four samples, only sample 96-13 was assessed for shear strength (estimated 15 kPa to 65 kPa), as discussed later in Section 4.2 of this report.

Sample 96-13 has been estimated to have ~ 2 wt% (dry) uranium metal and ~ 82 wt % (dry) total uranium content. It was postulated that uranium metal content might be a common factor among samples that have formed high strength agglomerates. The uranium metal concentration is unlikely to be the underlying cause of the agglomeration observed in sample 96-13 since sample SNF Comp containing crushed fuel, floor sludge, and canister sludge did not show agglomeration and/or self-cementation. The uranium metal concentration of the SNF Comp sample (determined from two prior gas generation tests) was found to be 60% to 70% of the original 48.2 g of uranium metal present. Assuming no uranium metal reacted between the end of gas generation testing and the start of the compaction testing, there was a maximum of 28.9 g to 33.7 g of uranium metal present at the beginning of the compaction testing. This equates to ~ 22.2 wt% (dry) uranium metal.

Another possible explanation for the presence of agglomeration in sample 96-13 and the absence of agglomeration in the SNF Comp sample is total uranium oxide content and the presence of hexavalent UOH compounds. As previously discussed, XRD analyses indicate the concentration of hexavalent UOH compounds metaschoepite and schoepite were relatively high in the agglomerated 96-13 sample when analyzed in 2007. In 1996, sample 96-13 was composed mainly of tetravalent uraninite compounds. The hexavalent UOH content in KC-2/3 M250 was also relatively high and no tetravalent uraninite was observed in 2007 (PNNL-16496, Table 3.3), indicating that there is a link between the formation of agglomerates and the extent of conversion of tetravalent uraninite compounds to hexavalent UOH compounds when samples are stored under PNNL hot cell conditions. This conjecture

is supported by XRD analysis performed in 1999, which showed sample 96-13 was high in tetravalent uraninite compounds and moderate in hexavalent UOH compounds (PNNL-13320, Table 2.7).

The SNF Comp sample, although relatively high in total uranium content (~60.2 wt% dry), was not analyzed after the *Compaction and Evaporation* testing and therefore cannot be compared to samples KC-2/3 M250 and 96-13 for the uranium compounds present. A probable reason sample SNF Comp did not undergo agglomeration is likely related to the extent of oxidation of the sample and the presence of non-uranium compounds. A significant portion of the sample is composed of non-hexavalent UOH compounds (~ 39.8 wt%-dry non-uranium constituents + 22.2 wt%-dry U metal) that may be interfering with the process of agglomeration. The presence of a significant concentration of uranium metal (~ 22 wt%-dry U metal) also indicates that a more reducing environment existed within the SNF Comp sample, which may have inhibited agglomeration further.

Homogeneity of the samples discussed may explain why some samples formed high strength agglomerates or self-cemented, and some did not. Agglomerates are more likely to form in sludge samples in which the majority of the uranium present is in the form of hexavalent UOH compounds. The presence of predominantly hexavalent UOH compounds in sludge may be predicted by the length of aging (storage), lack of gas generation, or shown through XRD results. For the KC-2/3 M250 sample, in which agglomeration occurred, the XRD results from 2007 identified the predominant uranium compound as hexavalent metaschoepite (PNNL-16496, Table 3.3). Samples SNF Comp and KC-2/3 whole, which had relatively high total uranium content, may not have agglomerated due to the slightly lower total uranium content, lower level of homogeneity and less conversion of tetravalent uraninite to hexavalent UOH compounds. Therefore, samples with a higher degree of homogeneity in the uranium concentration present may form agglomerates of higher strength than samples of a more heterogeneous nature.

Temperature and solubility may also play a large role in the formation of agglomerates in K Basin sludge. The higher temperatures of T Plant storage cells and the relatively oxidic environment, compared to the chilled (10°C to 15°C) anoxic K Basins, are expected to increase the rate of oxidation and formation of hexavalent UOH particles that are involved in the formation of agglomerates. Samples 96-13 and KC-2/3 M250 both experienced agglomeration at the hot cell ambient temperatures that typically range from 25°C to 35°C. Sample 96-13 showed evidence of oxidation via observed gas generation, which may also indicate that the sample was self-heating.

The solubility of the two main uranium phases believed to be involved in the oxidation reaction is shown in Figure 4. As the figure clearly demonstrates, hexavalent metaschoepite, found in the agglomerated samples 96-13 and KC-2/3 M250, is several orders of magnitudes more soluble in water at 25 °C and pH range of ~ 6 to 8 than tetravalent uraninite. This pH range is very similar to the pH range observed during the *Compaction and Evaporation* testing conducted with K Basin sludge samples (PNNL-15092, pg's 2.2 – 2.3, Table 2.1 and Table 2.2). This range of pH is likely similar to what can be expected at T Plant during long-term storage in the STSCs and as supernatant water is added periodically.

Figure 4 Solubility curve for U(IV) and U(VI) compounds⁵

It has been postulated that sample dry-out or loss of sludge cover water may also play a key role in agglomerate formation. This postulation mainly stems from the fact the sample 96-13 was discovered dry in the PNNL hot cell where it was stored ~ 6 months prior to the initiation of a 28-month *Compaction and Evaporation* test. The sample was then wetted, mixed, and allowed to soak for 6 months prior to the test. It is unclear to what extent the physiochemical properties of the sludge were changed by this unintentional dry-out, but the sample subsequently formed high strength agglomerates and self cemented in its storage container sometime during the test even though a constant layer of cover water was kept over the sludge during the study (PNNL-15092). The KE Pit sample was prepared from dried KE Basin Weasel Pit samples that were re-wetted and exhibited agglomeration. Sample KC-2/3 M250 exhibited the formation of high strength agglomerates even while being stored under a constant layer of cover water in the PNNL hot cells. The sludge was discovered to have agglomerated when taken out of the hot cells for use in hydrothermal testing (PNNL-16496). Two additional samples, KC-

⁵ Figure 4 courtesy of D.W. Shoesmith, 2000. Review, "Fuel corrosion processes under waste disposal conditions". Journal of Nuclear Materials 282 (2000) 1 – 31.

2/3 SS and KE Pit, both exhibited agglomerate formation while in PNNL hot cells and while being kept under a continuous layer of cover water.

In 1996, all but one of the samples collected from the KE Basin Weasel Pit and Basin Floor experienced a complete loss of supernatant water. Once the samples lost their cover of water it was very difficult to reconstitute (wet/mix) back to their original condition (HNF-6705). It is also difficult to determine whether the chemical and physical properties were permanently altered due to the drying out. This event of all the KE Basin Weasel Pit and Basin Floor samples experiencing a loss of cover water may indicate that self-heating caused by radiolytic decay and/or oxidation of uranium metal present in the samples was ongoing, or simply evaporation due to sample storage conditions. Therefore, it is likely that sample dry-out contributes to the agglomeration of sludge, but is not necessarily required for agglomeration to occur.

Of the above-mentioned conditions that may lead to agglomerate formation, storage time may be the most influential. Four sludge samples archived in the PNNL hot cells have experienced agglomeration as of 2007. These samples are 96-13, KC-2/3 M250, KC-2/3 SS and KE Pit. All four samples except sample 96-13 were collected in 1999 and had been in PNNL hot cell storage for ~ 8 years when discovered to have formed high strength agglomerates. Sample 96-13 was in storage from 1996 to 2002 (~ 6 years) before it was found to have dried out and then formed high strength agglomerates in subsequent compaction and evaporation testing. Storage time may be a major factor contributing to agglomerate formation for KE Engineered Container sludge, KW Engineered Container sludge, and Settler sludge because these sludges may be stored in the STSCs for up to 10 years at T Plant.

3.5 Postulated Physical and Chemical Mechanisms of Agglomerate Formation

Testing performed by PNNL demonstrates that certain sludge types can form high strength agglomerates (i.e., "hardpan"). In some cases, hardpan layers formed at the bottom of test vessels and in other instances they formed at the air-water interface as a floating crust. The formation of K Basin sludge agglomerates is likely related to the oxidation of uranium metal to higher oxidation states in water-saturated environments. Uranium metal is known to oxidize to uranium (IV) oxidation state via the following chemical reaction (equation 1). The rate of metal oxidation (as a linear penetration) is given by equation 2 and is valid for anoxic conditions over a temperature range of 24°C to 350°C (HNF-SD-SNF-TI-015, Vol 2, Rev 14A, Table 4.9).

$$\log_{10} R_c = 9.975 - \frac{3365}{T(K)} \left[\frac{\mu m}{h} \right] \quad (2)$$

Further oxidation of uranium (IV) oxides to uranium (VI) oxides (i.e., schoepite and metaschoepite) can be expected if the dissolved oxygen concentration is sufficiently high, which is likely to be the case when sludge is stored in STSCs at T Plant due to the vented container design. These hexavalent uranyl oxide hydrate (UOH) compounds will likely form at the air-water interface, or depending on the crystal structure and level of hydrophobicity, at the bottom of the vessel at the sludge-water interface. The

depth of oxygen penetration into the sludge bed will likely have a strong influence on the formation of UOH compounds in the STSCs. The diffusion rate of oxygen into a stationary sludge bed has been previously investigated (PNNL-14228, Appendix A) and the oxidation of uranium will likely be diffusion limited.

Previous works for uranium oxide fuel disposal in a geologic repository demonstrate that metaschoepite, an oxidation product of uraninite (UO_2), preferentially forms fractal aggregates at air-water interfaces. These fractal aggregates surprisingly float even though their densities are much greater than that of water. This anomalous behavior of heavy metal oxides floating at the surface of water-filled containers has been observed in several different colloid compounds including uranyl oxide hydrates, tungsten oxide, thorium oxide, plutonium oxide. It is postulated that this partitioning phenomena is due to the level of hydrophobicity and the orientation of surface charge in the crystal structure of metaschoepite (PNNL-14306). Similar instances of floating heavy metal oxide solids have occurred during laboratory testing of uraninite oxidation at PNNL⁶. These floating layers of agglomerates however have no perceptible strength. The shear strength of the UOH compounds formed during uraninite oxidation is discussed further in Section 4.1.

Figure 5, Figure 6 and Figure 7 (below) pictorially demonstrate these phenomena in progress. Figure 5 shows three particles of uraninite roughly 40 μm in diameter undergoing oxidation and forming UOH fractal aggregates of small particles consisting of primarily metaschoepite and small pieces of uraninite. Figure 6 is magnified image of the uraninite oxidation process, and shows particles of uraninite (dark) roughly 0.1 μm to 0.150 μm in diameter producing particles of metaschoepite (light) roughly 0.05 μm , as seen in detail in Figure 7.

Figure 5 Electron Micrograph of Agglomeration Process
Inverted Contrast Backscattered Electron Scanning Electron Micrograph of Micro- Crystals of Meta-Schoepite with Smaller Fragments of UO_2 Formed During Corrosion of Uranium Oxide.⁷

⁶ Delegard, C. H. , 2009. *Preliminary Status: Uraninite Aqueous Oxidation by Oxygen Gas at 50 °C*

⁷ Figure courtesy of PNNL-14306, pg 5.3

Figure 6 TEM image of Metaschoepite Formation

Figure shows the process of colloidal uranyl oxide hydrates forming from corrosion of uraninite particles.⁸

The images in Figure 7 are a magnification of the fractal aggregates formed at the air-water interface. The small roughly $0.05\ \mu\text{m}$ particles of metaschoepite coalesce and form sheets with fractal geometry as seen in both images below. The image on the right also indicates the presence of partially corroded uraninite pieces roughly $5\ \mu\text{m}$ in diameter and black in color.

Figure 7 Agglomeration of UOH fractal aggregates at the air-water interface.⁹

⁸ Figure courtesy of PNNL-14306, pg 5.4

⁹ Figure courtesy of PNNL-14306, pg 5.2

The exact mechanism of agglomerate formation among K Basin sludge is not known. The mechanism is likely a combination of many chemical and physical events that result in the formation of high strength agglomerates and the self-cementation of sludge as seen in samples 96-13, KC-2/3 M250, KC-2/3 SS and KE Pit. Of the possible mechanisms of agglomerate formation, Ostwald ripening could be a possible underlying mechanism of agglomerate formation. Ostwald ripening is defined as the growth of smaller precipitating particles to form larger agglomerates with the same chemical composition. Rhythmic precipitation mechanisms are also likely to be involved in agglomerate formation, as alternating bands of yellow and green UOH solids have been seen during uraninite oxidation testing¹⁰. This layering effect is sometimes referred to as “Liesegang Banding” or “Liesegang Rings”.

Even though multiple mechanisms may be involved in the formation of large high strength agglomerates, several similarities between the KE Basin sludge samples 96-13, KC-2/3 M250, KC-2/3 SS, and KE Pit are immediately apparent. All four samples formed agglomerates at relatively moderate temperatures, comparable to temperatures that can be expected in STSCs during storage at T Plant. XRD analyses show medium to high concentrations of metaschoepite and schoepite for samples 96-13 and KC-2/3 M250, which indicate a high degree of uraninite conversion to hexavalent UOH compounds. Unfortunately, XRD analyses could not be located for sample KC-2/3 SS and no crystalline uranium phases were detected in sample KE Pit. Samples 96-13, KC-2/3 M250 and KC-2/3 SS were all relatively high in total uranium content with a high degree of sample homogeneity. Lastly, all four of the samples were placed in PNNL hot cell storage for a period of 6 years to 8 years before being inspected indicating a correlation between storage time and agglomeration.

The extent of oxidation of the tetravalent uranium compounds to hexavalent UOH compounds is likely related to another important physical parameter, particle size. Since particles of relatively small diameter have a higher surface area to volume ratio, it is likely that the smaller diameter fraction of sludge is oxidizing more quickly than the larger diameter fraction. Although no particle size distribution data could be located for samples 96-13, KC-2/3 SS and KE Pit, the PSD for sample KC-2/3 M250 is roughly 250µm or less.

It is also possible that the removal of sludge from the anoxic environment of the basin is playing a key role in the extent of hexavalent UOH formation. The oxidation of sludge during storage in PNNL hot cells may proceed faster toward equilibrium products (UOH compounds) even if the samples are kept wet, due to the higher temperature (25°C to 35°C) as compared to the water of the K Basin (10°C to 15°C), and a more oxic environment. The rate of conversion of uraninite to metaschoepite at 21°C and 50° C via the reaction with dissolved oxygen in water has been established in PNNL-17678. Ultra pure oxygen gas was bubbled through the uraninite samples to convert the uraninite to metaschoepite. The mean starting diameter for the uraninite particles was 6 nm with agglomerates up to 600 nm (PNNL-17678, page 4.19). At 21°C, over 95% of the uraninite was oxidized to pure metaschoepite after 25 days. Since, the study described in PNNL-17678 used a stirred reaction vessel the conversion of uraninite occurred much faster than would be expected for K Basin sludge stored in the static environment of the STSC. This study however does provide the STP with a bounding case for the amount of uraninite converted (i.e., maximum extent of reaction possible in the STSC). Figure 8 below illustrates the increase in the rate of reaction with temperature for shrinking core kinetics in a stirred reaction vessel in terms of reaction progress.

¹⁰ Delegard, et al., 2009. *Preliminary Status: Uraninite Aqueous Oxidation by Oxygen Gas at 50 C*, Pacific Northwest National Laboratory, Richland WA.

Figure 8 Shrinking Core Kinetics for uraninite oxidation

Plot of the rate of uraninite oxidation at 21 °C and 50 °C in a stirred reaction vessel ($f=0$ indicates no oxidation has occurred or pure uraninite, and $f=1$ indicates oxidation of uraninite has gone to completion or pure metaschoepite).¹¹

The combination of this change in environment (elevated temperature and increased dissolved oxygen content), the relatively high total uranium content (i.e., high sample homogeneity) among three of the four samples examined, the high prevalence of hexavalent metaschoepite in samples 96-13 and KC-2/3 M250, and the relatively long storage periods in PNNL hot cells for all samples examined are all likely contributing factors in the formation of agglomerates. The agglomeration of KE Engineered Container sludge and KW Engineered Container sludge is less likely to occur due to the relatively low total uranium content compared to the sludge samples that have formed high strength agglomerates.

Recent KW Engineered Container (EC) sludge sampling and characterization indicates that the presence of constituents other than uranium oxides in K Basin sludge may actually inhibit agglomerate or “hardpan” formation (PNNL-19035). Samples were obtained from the KW Basin floor and canister sludge stored in KW Engineered Container SCS-CON-220 and KE Basin floor and canister sludge stored in KW Engineered Containers SCS-CON-240, -250, and -260. The following figure is a proposed illustration of this agglomeration inhibition effect. As Figure 9 illustrates, the formation of large agglomerates that could be detrimental to the removal of sludge after the prescribed interim storage at T Plant are physically inhibited by the presence of other particles in the sludge such as hydroxides of iron and aluminum.

¹¹ Figure courtesy of PNNL-17678, pg 4.21

Figure 9 Illustration of Agglomerates Formation

- (a) Illustration of agglomeration formation via coalescence of uranium (VI) oxide particles
(b) Illustrates the inhibition effect of agglomeration when other constituents are present

Sampling and characterization show that sludge from KW Engineered Containers SCS-CON-220, -240, -250, and -260 have much lower total uranium content than sample 96-13, KC-2/3 M250 and KC-2/3 SS (PNNL-19035 and PRC-STP-00212).

Table 3 and

Table 4 shows the total uranium content of the sludge currently stored in Engineered Containers SCS-CON-220, -240, -250, and -260. The data was collected from core sampling campaigns in the summer of 2009, and yield spatial information on the distribution of uranium compounds in the Engineered Containers. The highest total uranium concentration, 35.2 wt% (dry) was measured in core sample B1 from KW Engineered Container SCS-CON-220. This concentration is well below the level seen in the two samples (i.e. 68.3 wt% and 82 wt% dry) that formed agglomerates and indicates that the sludge is heterogeneous in nature.

Table 3 KW Engineered Container sludge dry weight % total uranium

	SCS-CON-220			
Core	A1	B3	A3	B1
U wt% (total-dry) ¹²	22.7	19.3	24.9	35.2
	23.2	20.2	17.6	34.0

Table 4 KE Engineered Container sludge dry weight % total uranium

	SCS-CON-240		SCS-CON-250	SCS-CON-260	
Core	A2	B1	B2	A1	B3
U wt% (total-dry) ¹³	5.37	4.65	5.39	6.00	6.10
					6.23

Further validity is given to the idea that the KE Engineered Container sludge and KW Engineered Container sludge will not form agglomerates when stored in static water filled STSCs in T Plant when the contents of Engineered Container SCS-CON-220 are studied further. As discussed in Section 4.1, mixtures of uranium oxides that have roughly 50% UO_2 and 50% UO_3 in a pure uranium oxide system have demonstrated the formation of uncharacteristically high shear strength agglomerates or “hardpans”. The approximately 50:50 ratios of UO_2 to UO_3 observed in recent samples taken from KW Engineered Container SCS-CON-220 is seen in Table 5, however no agglomeration has been observed. This may be because the SCS-CON-220 sample is not a pure uranium oxide system as the sample discussed in Section 4.1. The highest total uranium content observed in a SCS-SON-220 sample is 35.2 wt% (dry). The remaining constituents of the sludge may hinder the formation of large, high shear strength agglomerates as discussed previously and shown in Figure 9.

Table 5 Uranium Speciation Data by Engineered Container Core Composite Sieve Fraction

Engineered Container	SCS-CON-220			SCS-CON-240		SCS-CON-250		SCS-CON-260	
Sieve Fraction ¹⁴	LCC	MCC	SCC	MCC	SCC	MCC	SCC	MCC	SCC
U (IV) ¹³	>56.1	50.3 53.7	47.9 49.0	36.9	>92.9	>72.0	>88.1	>80.8 >82.4	67.8 63.8
U (VI) ¹³	>43.9	49.7 46.3	52.1 51.0	63.1	<7.1	<28.0	<11.9	<19.2 <17.6	32.2 36.2

¹³ Some of the core samples had multiple readings for total uranium content and speciation data as reported in PNNL-19035.

¹⁴ LCC is the sieve fraction with particle size greater than 2,000 μm . MCC is the sieve fraction with particle size greater than 500 μm but less than 2,000 μm . SCC is the sieve fraction with particle size less than 500 μm .

The following figure shows the correlation of total uranium content (dry wt% basis) with agglomerate formation among 14 samples collected from 1996 to 2009. Since the samples collected in 2009 have only had a limited period of time to age in the PNNL hot cells, the data associated with agglomerate or “hard pan” formation is somewhat limited. The four samples (96-13, KC-2/3 M250, KC-2/3 SS, and KE Pit) that exhibited agglomeration while stored in the PNNL hot cells are depicted as orange triangles in Figure 10. With the exception of the KE Pit sample, sludge samples with less than 59.0 wt% total uranium were not observed to form agglomerates. The KE Pit sample was prepared from other KE Basin Weasel Pit samples that were dried and re-wetted, which could explain why this sample formed agglomerates even though the total uranium content was only 8 wt% (dry).

Figure 10 Correlation of total uranium (dry basis)

Figure 10 shows the correlation of agglomerate formation with total uranium for 16 samples of KE and KW Basin sludge.

While the formation of agglomerates is undoubtedly a complex mechanism that is yet to be fully understood, the extent of conversion of U(IV) compounds to U(VI) compounds, the total uranium content (i.e., sample homogeneity) of the sludge and storage time are believed to be the main contributing factors. Since only 4 sludge samples and sub-samples taken from the K Basins have been

identified to form agglomerates, the available quantitative data is insufficient to determine the weight of each of these postulated factors.

To summarize, the mechanism of agglomerate formation involves several steps and prerequisite conditions. First, the sludge must contain a high level of total uranium oxides that can react with dissolved oxygen in water to yield UOH compounds (e.g., metaschoepite and schoepite). These small oxidation products then spall off the larger uranium oxide pieces and coalesce to form fractal aggregates of UOH particles. This process is expedited in environments with higher temperatures and a ready supply of oxygen, which is similar to the conditions in the PNNL hot cells or T Plant. These networks of UOH particles subsequently congregate and form larger agglomerates. In order for large high strength agglomerates to form, the sample homogeneity must be relatively high, as in samples 96-13 (82 dry wt% total U), KC-2/3 M250 (68.3 dry wt% total U) and KC-2/3 SS (~ 59.0 dry wt% total U). Due to the relatively low total uranium content in the core samples collected from the KW Engineered Containers, the formation of rigid high shear strength agglomerates is not expected to occur.

There is however always the possibility of localized pockets of sludge high in uranium concentrations that could possibly meet all the prerequisites for the formation of high shear strength agglomerates inside the STSCs. It is likely, however, that jet retrieval will break up any high uranium content pockets of sludge during Phase 2 jet retrieval. Additionally, the Settler Tank sludge, which is being retrieved into KW Engineered Container SCS-CON-230, is estimated to contain 77 wt% to 82 wt% (dry) total uranium. The Settler Tank sludge is likely to be more susceptible to agglomerate formation than the KE Engineered Container sludge and KW Engineered Container sludge types due to the higher total uranium content and smaller particle size ($< 600 \mu\text{m}$), which is similar to the total uranium content of the 96-13 sample.

Ongoing testing at PNNL, by Delegard et al. may provide additional insight on the mechanisms responsible for agglomerate formation. The testing is utilizing an apparatus consisting of six test tubes containing settled uraninite solids that are subjected to oxidation via a steady flow of ultra-pure (99.99%) oxygen at 50°C .

4.0 Agglomerated Sludge Mobilization

If agglomeration is suspected to occur during storage at T Plant, the STP needs to know the relative shear strengths of such material for the design of a mobilization and retrieval tool. The following sections discuss the effect of the conversion of tetravalent uranium (IV) compounds (i.e., uraninites) to hexavalent uranium (VI) compounds (i.e., UOHs) has on the shear strength of a 100% uranium oxide system. In addition to the pure uranium oxide system discussed in section 4.1, real K Basin agglomerated sludge sample 96-13 was evaluated for shear strength and the results are discussed in Section 4.2.

4.1 Shear Strength of Varying U Oxide Mixtures

Testing performed by PNNL in February 2010¹⁵ show that mixtures of uranium oxides with varying tetravalent uranium and hexavalent uranium (i.e. UO_2 : $\text{UO}_3 \cdot 2\text{H}_2\text{O}$) ratios differ significantly in the shear strength or yield stress. Samples of tetravalent uraninite in water were oxidized by bubbling ultra pure oxygen into the water at room temperature. Three samples with uranium oxide compositions 100% UO_2 (sample 1), 49% UO_2 :51% $\text{UO}_3 \cdot 2\text{H}_2\text{O}$ (sample 2), and 35% UO_2 :65% $\text{UO}_3 \cdot 2\text{H}_2\text{O}$ (sample 3) were transferred to 120 ml jars and allowed to settle undisturbed for a period of 3 ½ months. After the settling periods, the undisturbed shear strength were determined to be 350 Pa to 430 Pa, greater than 27,000 Pa (rheometer maximum torque exceeded), and 80 Pa to 100 Pa for samples 1, 2, and 3, respectively.

In the process of removing the shear vane from sample 2 (49% UO_2 :51% $\text{UO}_3 \cdot 2\text{H}_2\text{O}$), the material was partially mixed. This provided PNNL the opportunity for a partially mixed shear measurement. In order to take an intermediate measurement on sample 2, the sample was allowed to settle for 72+ hours after the partial mixing. After the 72+ hour settling time, sample 2 was determined to have shear strength of 3,800 Pa to 16,000 Pa, which is significant. All three samples were then mixed aggressively and allowed to settle for 72+ hours. Samples 1, 2, and 3 were determined to have shear strengths of 140 Pa to 190 Pa, 980 Pa to 1,700 Pa, and 45 Pa to 53 Pa, respectively.

To determine if the sample shear strength rebuilds after aggressive mixing, an additional shear strength measurement was collected in March 2010 following ~5 weeks of undisturbed settling¹⁶. Prior to the 5 week settling period, all three samples were manually mixed with a metal spatula for ~5 minutes. The 49% UO_2 : 51% $\text{UO}_3 \cdot 2\text{H}_2\text{O}$ mixture could not be completely homogenize with the spatula, and after mixing, the sample was lumpy with well-mixed portions. A layer of water was added to the top of the container to minimize water loss during the settling period inside the fume hood. Unfortunately, the container was not airtight and evaporation of the overlying water occurred. The top of the sample formed a very thin 'skin' and a small fracture down the center of the sample was observed. The shear strength observed for the 49% UO_2 : 51% $\text{UO}_3 \cdot 2\text{H}_2\text{O}$ mixture was slightly higher than the values obtained at 72 hr settling time. Water was observed seeping out of the sample when the vane was inserted further down (~32 mm) suggesting that pockets of water were trapped within the sample. The 100% UO_2 sample had a well-defined settled layer and exhibited slightly higher shear strength values than at the 72 hour settling time. The shear strength of the sample did not reached the "as received" values initially measured. The 35% UO_2 : 65% $\text{UO}_3 \cdot 2\text{H}_2\text{O}$ also had a well defined settled layer and the shear strength values obtained were similar to the 72 hour settling data. Again the "as received" shear strength was not reproduced after ~5 weeks of settling.

The shear strength of the less than 500 micron small container composite (SCC) fraction of sludge samples collected from KE Engineered Containers SCS-CON-220, -240, -250, and -260 have also been measured (PNNL-19035). The samples identified as KW 240 SCC, KW 250 SCC and KW 260 SCC have estimated shear strengths of 300 Pa, 300 Pa and 79 Pa, respectively. The results of the uranium oxide shear strength testing and the 2009 KE Engineered Container shear strength testing and uranium speciation testing are summarized below in Table 6.

¹⁵ February 2010 Monthly Status Report For PNNL Contracts Providing Technical Support to K Basin Closure and Sludge Treatment Projects, Pacific Northwest National Laboratory, Richland WA.

¹⁶ March 2010 Monthly Status Report For PNNL Contracts Providing Technical Support to K Basin Closure and Sludge Treatment Projects, Pacific Northwest National Laboratory, Richland WA.

Table 6 Results of mixed oxide shear strength in a pure uranium system

Test/Sample ID	U Total (Dry wt%)	Uranium Speciation	Shear Strength (Pa)		
			3 ½ months of settling and oxidation	Aggressively Mixed then Settled for 72+ hrs	Aggressively Mixed / Settled for 1 month
Sample 1	100	100% UO ₂	350 to 430	140 to 190	187 to 250
Sample 2	100	49% UO ₂ 51% UO ₃	>27,000	980 to 1,700	1600 to 2000
Sample 3	100	35% UO ₂ 65% UO ₃	80 to 100	45 to 53	39 to 55
"As Settled" (20+ days)					
KW 240 SCC	5.37	>92.9% UO ₂ <7.1% UO ₃	NA	300 ± 200	NA
KW 250 SCC	5.39	>88.1% UO ₂ <11.9% UO ₃	NA	300 ± 100	NA
KW 260 SCC	6.23	67.8% UO ₂ 63.8% UO ₂ 32.2% UO ₃ 36.2% UO ₃	NA	79 ± 3	NA

The results from Table 6 suggest that mixtures of uranium (IV) and uranium (VI) oxides produce maximum shear strength when the composition is roughly 50:50 (i.e., 50% UO₂ and 50% UO₃). PNNL staff state that the behavior of sample 2 (~ 50% UO₂:50% UO₃·2H₂O) has been observed before and was not unexpected. These results listed in Table 6 are plotted below in Figure 11 and shows that at roughly 50% U(VI) conversion, the shear strength of the sample was at a maximum (~ 1,700 Pa).

To date, the highest shear strength measurements for "actual" K Basin sludge samples were obtained after hydrothermal treatment for 7 to 10 hr at 185°C of high-uranium-content KE canister sludge. The unconfined compressive strength of samples from the hydrothermal testing, measured by a pocket penetrometer, infers that their shear strength may be between 120 and 170 kPa (PNNL-16496). These short-duration hydrothermal tests were conducted at temperatures much greater than the projected maximum temperature of the T Plant canyon cells (-7 °C to 32°C); however, the strength results provide an initial bounding target for sludge stored for many years, and an upper range for simulants (042910-53451-TP02 Rev 0).

Figure 11 Correlation of Shear Strength with U(VI) concentration (Dry basis)
Plot of measured shear strength as a function of U(VI) wt% on a dry basis for three mixed uranium oxides samples and three KE and KW Engineered Container samples (see Table 6).

4.2 Evaluation of Jet Erosion

In a report by PNNL, *Assessment of Jet Erosion for Potential Post-Retrieval K-Basin Settled Sludge* (PNNL-18831), the bounding shear strength of the KE Basin sludge sample 96-13 was estimated. The KE Basin Canister sludge sample 96-13 was characterized as a heterogeneous material with non-cohesive “chunks” being held together by a cohesive “paste”. The estimated shear strength of the cohesive “paste” material is 3 kPa to 5 kPa, and that of the non-cohesive “chunks” is 380 kPa to 770 kPa. The estimated bulk shear strength of sample 96-13 was determined to be 15 kPa to 65 kPa. These shear strength measurements however are based on review of video taken when the 96-13 sample was removed from the graduated cylinder, used in the 28-month compaction and dry-out test. The shear strength was estimated based on video and “estimated applied forces, approximated contact areas, and the application of an estimate of the relation of compressive and shear stresses”. Therefore, the results stated above should be considered qualitative, at best.

Recommendations in PNNL-18831 for future work include using a pocket penetrometer to quantify the shear strength of actual K Basin sludge samples in a hot cell environment. It is recommended that this method of determining shear strength be applied to homogeneous sludge samples, such that the results will be directly comparable to that obtained using a shear vane.

Erosion mechanisms for both cohesive and non-cohesive (i.e., “paste” and “chunks”, respectively) were also presented in PNNL-18831. Recommendations were made for future experimental hot cell work to determine the erodibility of cohesive and non-cohesive fractions of actual K Basin sludge. By determining the bounding conditions by which erosion occurs for both the larger non-cohesive “chunks” and the smaller cohesive “paste” fraction, the STP can design, procure and implement an acceptable device for post interim storage sludge retrieval from the STSCs.

5.0 Physical Properties of Dried KE Engineered Container sludge and KW Engineered Container sludge

Physical property characterization testing with sludge core samples from KW Engineered Container 220 in 2009 provides data on the behavior of agglomerates once dried out. Figure 13 and the associated text have been borrowed from PNNL-19035 and demonstrate that the majority of solids self-fractured when the samples dried out. This resulted in free-flowing granular material or friable solids that could be broken up with a small stainless steel spatula with little force. These images of KE Engineered Container sludge and KW Engineered Container sludge that have dried out during testing may indicate that if the sludge in the STSCs were to dry-out the resulting solids would be friable and self-fracture making them easily mobilized when re-wetted.

Sample TI004-1G (KW 220 Core A3)
(Free flowing powder)

Sample TI004-2G (KW 220 Core B1)
(Agglomerates fractured with little force using a stainless steel spatula)

Sample TI004-1G (KW 220 Core A3)
(Large agglomerate in left corner of picture was the top fraction of the sample. The remainder of sample appears to have a high concentration of organic ion exchange resin).

Sample TI004-1G (KW 220 Core A3)
(Large agglomerate, seen in middle left picture, fractured with little force using 1-cm wide stainless steel spatula)

Figure 12 Post-Test Examination of Physical Property Samples from SCS-CON-220 Core Samples

6.0 Conclusion

The STP plans to store K Basin sludge in water filled STSCs at T Plant for up to 10 years prior to retrieval from the STSCs for final processing and disposal. The formation of high strength agglomerates, also known as “hard pan” or “hard packed” materials, while K Basin sludge is stored in the STSCs is of particular concern to the STP and is an issue that needs to be addressed prior to post interim storage activities.

Four instances of high strength agglomerate formation have been observed among samples collected from the KE Basin and stored in the PNNL hot cells. These four samples are designated 96-13, KC-2/3 M250, KC-2/3 SS and KE Pit. Six additional instances of “hardpan” or “hard packed” material have been observed while collecting samples from the KE Basin during the FY 1995 and FY 1999 sampling campaigns.

Sample 96-13 (82 wt% total uranium) was collected in FY 1996 and is composed of sludge collected from a single fuel storage canister (canister constructed of aluminum, containing fuel in poor condition). Prior to the initiation of a 28 month *Compaction and Evaporation* test, the sample was found dry in the PNNL hot cell where it was stored and had to be crushed and reconstituted for 6 months before the study commenced (PNNL-15092). At an unspecified time during the study sample 96-13 formed high strength agglomerates and self-cemented in the graduated cylinder where it was being stored. The sample could not be removed using conventional means (i.e., stainless steel spatula, or inverting the graduated cylinder) and the graduated cylinder had to be broken to extract the sludge. The sample was subsequently sheared with the hot cell manipulator arm and stored in water-filled jars in 2004. Sample 96-13 was re-examined in July 2009. It was observed that the sample had dried out in storage and the large high strength agglomerates present in 2004 had self-fractured into small particles (July 2009 Monthly Status Report).

A similar situation to sample 96-13 occurred with another sample collected from the KE Basin and stored in PNNL hot cells. Sample, KC-2/3 M250 (68.3 wt% total uranium) was prepared in FY 1999 by wet-sieving sample KC-2/3 whole to obtain a “minus” 250 μm fraction. The sample was stored in the PNNL hot cells until 2007, when it was retrieved for utilization in *Hydrothermal* testing (PNNL-16496). During the ~ 8 years of storage at hot cell temperatures of ~ 25°C to 35°C, the sample formed high strength agglomerates and self-cemented in the storage jar, despite a water layer being kept over the sludge in the storage jar. Similar to sample 96-13, this sample could not be removed by conventional means and the storage jar was broken to retrieve it.

Two additional samples, KC-2/3 SS (estimated 59 wt% total uranium) and KE Pit (~8 wt% total uranium-dry basis) also exhibited agglomerate formation while being stored in the PNNL hot cells (PNNL-17078, pg 29, Table 1). Like samples 96-13 and KC-2/3 M250, these samples could not be removed from their storage jars by conventional means and the jars had to be broken to extract the sludge.

In addition to the four laboratory samples, six instances of “hardpan” or “hard packed” material were observed during FY 1995 and FY 1999 sampling campaigns. When these samples were collected, the agglomerates that were likely present were disrupted by the sampling method and not observed during subsequent laboratory testing. It is likely that more samples collected during these campaigns would have formed agglomerates. However many samples were fully consumed in laboratory testing or inadvertently allowed to dry, which may have altered their physical and chemical properties.

Agglomeration was observed in sample KE Pit while being stored in the PNNL hot cells, however further analysis related to agglomeration of this sample could not be located and likely was not performed.

XRD analyses of samples 96-13 in 1997 (HNF-SP-1201, pg 29, Table 3.4) and KC-2/3 M250 in 2000 (PNNL-13320) indicate that the major uranium compounds in the samples were uraninites. Subsequent analyses performed in 2007 (PNNL-17078) show that the uranium compounds in samples 96-13 and KC-2/3 M250 have since oxidized to hexavalent uranium phases (e.g., metaschoepite and schoepite). The formation of metaschoepite and schoepite are believed to be a primary factor in the formation of high strength agglomerates.

XRD analyses performed on sample KE Pit in 1995 indicate no uranium phases present, however there could have been amorphous uranium phases present. XRD did however detect two possible alteration phases of mordenite (Zeolon-900 ion exchange material), albite ($\text{NaAlSi}_3\text{O}_8$) and gismondine ($\text{CaAlSi}_2\text{O}_8 \cdot 4\text{H}_2\text{O}$). The detection of two possible alteration phases of mordenite is significant due to its tendency to dissolve and re-precipitate as different chemical forms under alkaline conditions. Poor water quality in the KE basin during the years prior to the collection of KE Pit, may have resulted in agglomerate formation. This may explain why KE Pit formed agglomerates without testing positive for the uranium phases commonly seen among agglomerated sludge samples, namely metaschoepite and schoepite. Unfortunately, no XRD analysis was located for sample KC-2/3 SS for comparison.

Additional analyses point toward a possible correlation between the level of homogeneity of the sludge and the formation of high strength agglomerates. Samples 96-13 and KC-2/3 M250 were found to have a relatively high total uranium content of 82 wt% (dry) and 68.3 wt% (dry), respectively, indicating a high level of sample homogeneity with respect to uranium content. Sample KE Pit was determined to have ~ 8% total uranium content (dry basis), however the iron content of this sample was relatively high (36.4 wt% dry basis) compared to the other three samples which showed agglomerate formation. No specific uranium speciation data could be found on sample KC-2/3 SS. The total uranium content of KC-2/3 SS is ~59.0 wt% (dry), based on KC-2/3 whole.

Sample homogeneity with respect to uranium content is thought to effect the formation of large high strength agglomerates. Sludge that is of a more heterogeneous nature and contains hexavalent uranium compounds like that observed in the Engineered Containers in KW Basin has not been shown to form agglomerates (PNNL-19035). It should be noted that the samples collected from the KW Engineered Containers have not had sufficient time in storage in the PNNL hot cells to determine if agglomerates will form over the up-to 10 year storage period at T Plant.

In addition to total uranium content and sample homogeneity, the conditions that the sludge is stored in may also be a factor in the formation of agglomerates. Samples 96-13 and KC-2/3 M250, which formed agglomerates, were stored in PNNL hot cells for up to 8 years prior to being removed for use in testing. The temperature in the PNNL hot cells (25°C to 35°C) is higher than the temperature of the K Basin water (10°C to 15°C). The concentration of oxygen available to the sample is also much higher than in the anoxic water in the K Basin. Since the oxidation of tetravalent uraninite compounds has a rate that is first order with respect to oxygen and a rate constant that increases with temperature (i.e., Arrhenius relationship), the increase in temperature and available oxygen are likely to accelerate the formation of hexavalent uranium compounds and the formation of agglomerates. The exothermic nature of oxidation and radiolysis reactions occurring simultaneously generates heat, which may contribute further to the formation of agglomerates.

The relative solubility of metaschoepite and schoepite versus uraninite may be a factor in the formation of high strength agglomerates. As seen in Figure 4, the solubility of uranium (VI) compound is several orders of magnitude higher than that of uranium (IV) in water at 25 °C and a pH range of 6 to 8. When the sludge is removed from the anoxic and chilled basin and placed into a STSC at elevated temperatures and relatively much more oxic environment, the rate of uranium (IV) oxidation is expected to increase. The more soluble uranium (VI) will likely react with dissolved oxygen available through the STSC vent or from the radiolysis of water to form uranium (VI) compounds (metaschoepite and possibly schoepite). This reaction will then result in the precipitation of metaschoepite when concentrations reach solubility limits.

It has been suggested that sample dry-out may also play a key role in agglomerate formation. This postulation mainly stems from the fact the agglomerated sample 96-13 was discovered dry in the PNNL hot cell in 2002 (PNNL-15092). It is unclear to what extent the physiochemical properties of the sludge were changed by this accidental loss of cover water, but the sample subsequently formed high strength agglomerates and self cemented in its storage container sometime during the testing even after being reconstituted (PNNL-15092). Sample KC-2/3 M250, however, exhibited the formation of high strength agglomerates while being kept under a constant layer of cover water in the PNNL hot cells. The sludge was discovered to have agglomerated when taken out of the PNNL hot cells in 2007 (PNNL-16496). Two additional samples, KC-2/3 SS and KE Pit, both exhibited agglomerate formation while in PNNL hot cells after being kept under a continuous layer of cover water. Therefore, it is likely that sample dry-out does contribute to the agglomeration of sludge, but is not necessarily required for agglomeration to occur.

Of the conditions that may lead to agglomerate formation mentioned above, storage time may be the most influential. As of 2007 only samples 96-13, KC-2/3 M250, KC-2/3 SS and KE Pit, have experienced agglomerate formation while in PNNL hot cell storage (PNNL-17078). All four samples except sample 96-13 (collected in 1996) were collected in 1999 and had been in PNNL hot cell storage for ~ 8 years when observed to have formed high strength agglomerates. Sample 96-13 was in storage from 1996 to 2002 (~ 6 years) before it was found to have dried out and then formed high strength agglomerates during subsequent testing. Storage time may have a significant impact on agglomerate formation for KE Engineered Container sludge, KW Engineered Container sludge, and Settler sludge placed in the STSCs for up to 10 years due to the observed correlation between storage time and the formation of high strength agglomerates in sludge samples at PNNL.

The actual physiochemical mechanism(s) of high strength K Basin agglomerate formation is not known. However several commonalities and proposed mechanisms that likely participate in this phenomenon are the formation of floating uranium (VI) aggregates or loose flocs, and ultimately agglomerates at the air-water interface, rhythmic precipitation of metaschoepite and other uranium (VI) compounds forming "Liesegang Bands" at the sludge-water interface and Ostwald ripening seen in many crystal formation mechanisms.

Sampling and characterization activities conducted in 2009 have measured the uranium content and uranium speciation for sludge stored in four of the six Engineered Containers located in the K West Basin. Ratios of U(IV) to U(VI) that are similar to samples that have shown significant shear strength in recent testing¹⁷ are also present in KW Engineered Container 220, but agglomerate formation has not

¹⁷ February 2010 Monthly Status Report For PNNL Contracts Providing Technical Support to K Basin Closure and Sludge Treatment Projects, Pacific Northwest National Laboratory, Richland WA.

been observed in core samples collected from this container. This can be explained by the degree of homogeneity of the sludge in container 220. The maximum total uranium content in Engineered Container 220 sludge is 35.2 wt% (dry) as measured in core B1, which is much lower than the total uranium content in samples that have formed agglomerates. In other words, the sludge stored in Engineered Container 220 has a significant concentration of non-uranium particles that may be inhibiting the formation of large high strength agglomerates. Settler sludge however is estimated to contain ~ 80 wt% (dry) total uranium and thus may be more susceptible to agglomerate formation depending on the relative concentrations of U(IV) and U(VI) compounds that are yet to be determined.

There is a possibility of high strength agglomerates self-cementing inside storage vessels, and the shear strength and erodibility properties of these materials need to be studied and quantified. One study performed by PNNL provides the STP with a rough estimate of the shear strength of sample 96-13 and assessed the potential for using a jet erosion method to erode and mobilize agglomerates (PNNL-18831). The shear strength of this sample was estimated to be 15 kPa to 65 kPa for the bulk material. This value was much higher for the non-cohesive "chunks" (380 kPa to 770 kPa) and much lower for the cohesive "paste" (3 kPa to 5 kPa) materials, as discussed in Section 4.2.

Ongoing uraninite oxidation testing by PNNL will provide the STP with more accurate readings for the shear strength of agglomerated sludge with mixed uranium oxide composition and possibly lead to further insight on agglomerate formation mechanisms.

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Appendix A Compaction and Evaporation Test sample characteristics (PNNL-15092).

Sample ID	Sample Type	Date Collected/ prepared	Storage /Test Temperature °C	Duration	Agglomeratio n Strength	Uranium Total, wt% dry	Starting Uranium Phases ^{18,12}	Ending Uranium Phases ^{19,11}	Misc. Comments
FE-5	KE Weasel Pit	1999	31 to 38	28 month	Insignificant	5.32	None detected	Uranophane (L)	
KC-2/3 whole	KE Canister sludge composite from all three bays	1999	31 to 38	28 month	Insignificant	59.0	Sc, Ur	Metaschoepit e (H), schoepite (H)	
KC-2/3 M250	"Minus" 250 µm fraction of KC-2/3 whole (KE Canister sludge composite)	1999	~34	~8 years in PNNL Hot Cells	Not measured	68.3	Ur (H), MSc (M)	Metaschoepit e (H)	Subsample prepared shortly after collection in 1999 by wet sieving to obtain a "Plus" 250 µm fraction and a "Minus" 250 µm fraction".
KC-2/3 SS	Settled Solids Fraction of KC-2/3 whole	1999	~34	~8 years in PNNL Hot Cells	Not Measured	59.0	Not Measured		
KE-Pit	Weasel Pit Sludge Sample	1999	~34	~8 years in PNNL Hot Cells	Not Measured	7.99	Not Measured		
KC-4 (whole)	KE Floor sludge between canisters (Composite sample)	1999	31 to 38	28 month	Insignificant	16.6	Not Measured	Schoepite (M), sodium uranyl oxide hydroxide (M), Be (H)	
KC-5	KE Floor sludge from main basin (Composite sample)	1999	31 to 38	28 month	Insignificant	6.36	Be, Sc (KC-5 P250 only)	Be (H) from KC Floor Comp	KC-5 (Sub-sample of KC-4 whole)
96-13 (solid grad) ²⁰	KE Canister sludge	1996	31 to 38	28 month	monolithic 15 to 65 kPa PNNL-18831	82.0	Ur	Metaschoepit e, schoepite	Sample experienced a complete loss of supernatant while in storage, prior to testing.
SNF- Comp	Floor, canister, and crushed fuel mix (Composite sample)	1999	31 to 38	28 month	Insignificant	60.2	Be, Sc, Ur	Not Measured	Original sample name: SNF-Comp 60S-PG
KE NLOP	North Loadout Pit	2003	31 to 38	28 month	Insignificant	2.51	None Detected	Metaschoepit e	

¹⁸ Starting U Phases data taken from Table 2.2, PNNL-15092 for all samples excluding KC-2/3 M250 (see foot note 12)

¹⁹ Ending U Phases data taken from Table 6 of PNNL-17078, 2007 (a.k.a, Letter Report 53451-RPT01, 2007)

²⁰ XRD was performed on sample 96-13 in 2007, 3 years after conducting the long-term settling test (PNNL-17078)

¹² KC-2/3 M250 ending uranium phases identified as metaschoepite (PNNL-16496, Table 3.3)